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HINDERED ROTOR PROGRAMME AND SOME CRYSTALLOGRAPHIC
STRUCTURES

by



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ABSTRACTS

Chapter I:

Lipscomb and King's hindered rotor expression was applied to the rigid body routine of M.J. Bennett and B.M. Foxman. The programme SFLS5 was modified to calculate the contribution of the hindered rotor to the structure factors and to do a full matrix least squares refinement of the input parameters. The programme was tested on several sets of published X-ray data. The results showed that the model gives correct bond lengths without recourse to "riding corrections". The barrier parameter B_d is related to the RMS angular displacement of the ring but is not necessarily a measure of activation energy.

Chapter II:

The compound ferrocinium picrate $[\text{Fe}(\text{C}_5\text{H}_5)_2]^+$ $[\text{H}_2\text{C}_6\text{N}_3\text{O}_7]^-$ crystallizes in space group Cmcm ($Z = 4$) with cell dimensions $a = 12.513(7)\text{\AA}$, $b = 20.267(9)\text{\AA}$, $c = 6.903(7)\text{\AA}$. The structure was refined, using SFLS5 as modified for the hindered rotor model, to a conventional R factor of 6.0%. The compound consists of two ionic species, the ferrocinium cation, and the picrate ion. Both ions sit on sites of mm symmetry and show some disorder.

Chapter III:

Hydridotriphenylsilyl (π -cyclopentadienyl) dicarbonyl manganese, $(\pi\text{-C}_5\text{H}_5)(\text{OC})_2\text{Mn}(\text{H})\text{Si}(\text{C}_6\text{H}_5)_3$, crystallizes in the space group $P2_1/n$ with unit cell dimensions $a = 13.200(3)\text{\AA}$, $b = 17.328(4)\text{\AA}$, $c = 9.438(2)\text{\AA}$, $\beta = 92.47(4)^\circ$, $Z = 4$. The structure was refined to a conventional R factor of 4.2%. The geometry around the manganese atom is that of a distorted square pyramid, with that of the silicon atom a distorted tetrahedron. The hydridic hydrogen was located on full data and low angle difference maps and was subject to refinement. It is located in a bridging position, lying $1.54(4)\text{\AA}$ from the manganese and $1.76(4)\text{\AA}$ from the silicon.

Chapter IV:

Dihydrooctacarbonyl dirhenium, $\text{H}_2\text{Re}_2(\text{CO})_8$, crystallizes in space group $P2_1/c$, $a = 8.96(2)\text{\AA}$, $b = 11.62(2)\text{\AA}$, $c = 12.85(2)\text{\AA}$, $\beta = 109.2(1)^\circ$, with four molecules per unit cell. The structure was refined to a conventional R factor of 10.5%, all atoms isotropic. There is octahedral geometry about each rhenium, the hydrogens being assumed to occupy the vacant co-ordinated positions. The doubly hydrogen bridged Re-Re bond of $2.896(3)\text{\AA}$ is shorter than normal Re-Re single bonds.

The topics covered in this thesis involve X-ray diffraction. Undefined symbols have their standard crystallographic meaning and are defined in Appendix C. The standard theoretical and experimental procedures can be found in detail elsewhere^{1,2,3}.

Chapter I

THE HINDERED ROTOR

INTRODUCTION

In X-ray diffraction the structure factor F_{hkl} can be generally expressed as²

$$F_{hkl} = \int \int \int \rho_{xyz} \exp[2\pi i \underline{h} \cdot \underline{r}] dx dy dz . \quad [1]$$

If the assumptions are made that atoms are at rest and the electron density is zero elsewhere in the cell, the expression simplifies to

$$F_{hkl} = \sum_i f_i \exp[2\pi i (hx_i + ky_i + lz_i)] \quad [2]$$

where f_i is the scattering amplitude of atom i at an angle θ and allows for the finite size of the atom.

In general, atoms are not at rest, but are in continual motion. To allow for this, Debye³ derived a temperature parameter B which was related to the mean square displacement of an isotropic harmonic oscillator. Then F_{hkl} became

$$F_{hkl} = \sum_i f_i \exp[-B(\sin\theta/\lambda)^2] \exp[2\pi i (hx_i + ky_i + lz_i)] . \quad [3]$$

In many cases a better description of the motion of the atoms can be obtained by assuming a general rectilinear harmonic oscillator⁴. This model is conveniently handled in tensor form and the six unique thermal parameters of the general case are related to the root mean square

amplitudes and the direction cosines of the three principal axes of the ellipsoid of vibration. The isotropic and anisotropic models are usually considered practical to describe most crystal structures.

The major disadvantage of these models lies in the large number of parameters required to describe the crystal structure, especially in large molecules such as $\text{RhH(CO)(P(C}_6\text{H}_5)_3)_3}$ ⁵. Since the isotropic model requires four parameters per atom (x, y, z, B) in the general position and the anisotropic model requires nine parameters per atom (x, y, z and for example $\beta_{11}, \beta_{22}, \beta_{33}, \beta_{12}, \beta_{13}, \beta_{23}$) a full matrix refinement of a large molecule would require excessive computer time and storage. There are two obvious solutions to this problem. The first is to use block diagonal refinement⁶ rather than full matrix refinement. This cuts down greatly on both time and storage requirements, but has the disadvantage that some large parameter correlations can be ignored, even with an apparently sensible choice of blocks.

The second method retains the use of full matrix refinement and utilizes known geometries for certain rigid groups^{7,8}, the disadvantage of the method being that significant deviations from ideal geometry will be missed. In this method, certain groups of atoms are

treated as a unit. It is assumed that the groups are internally rigid and the temperature factors of the atoms are dependent only on the motion of the rigid body as a whole. Such a group may be defined by setting up an internal Cartesian co-ordinate system within the rigid body. The positions of individual atoms within the group can then be specified with reference to the internal co-ordinate system. This co-ordinate system can be related to the crystallographic co-ordinate system by means of a transformation matrix. Since the atoms within the group are assumed to have fixed positions relative to the internal co-ordinate system, it is only necessary to refine the origin and orientation of the co-ordinate system with respect to the crystallographic axes, as well as refining temperature factors for the individual atoms and/or the rigid body as a whole. Use of such a model greatly reduces the number of parameters to be refined.

Use of isotropic temperature parameters for the atoms is based on the assumption that the rigid body is undergoing only relatively small oscillations from its mean position⁹. The opposite extreme occurs if the rigid body is rotating freely. Expressions for the scattering of X-rays from such groups have been calculated by Bijvoet and Ketellar¹⁰ and Zachariasen¹¹.

Intermediate between the two extremes of free rotation and no rotation is the case where the rotation of the group is hindered. Cruickshank¹² found this to be the case in benzene at -3°C . After anisotropic refinement of the carbon atoms, he found that the tangential motion of the atoms was much greater than the radial motion indicating an oscillation of the benzene ring about the hexad axis. In order to describe this motion better, he assumed that the benzene ring was a rigid body. Then its motion could be described in terms of two tensors, T describing the translational motion, and ω associated with oscillations¹³. Using this method of analysis, he showed that the major ring oscillation was about the hexad axis with a RMS (root mean square) angular oscillation of 7.9° . From this it was also possible to calculate the apparent shortening of the C-C bond due to the motion of the ring¹⁴. The method was later extended by Schomaker and Trueblood¹⁵ to the case where there was no pre-ordained centre of rotation and allows for a coupling of rotation and translation via the introduction of a third tensor S . The Cruickshank T and ω tensor elements have been used as model parameters within a structure factor and least squares programme¹⁶ although strictly speaking the site symmetry of the group is limited. The general case of Schomaker and Trueblood

has not been incorporated into a structure factor and least squares programme and its application is currently limited to the analysis of the results of an anisotropic independent atom refinement.

A different approach to the problem of hindered rotation was taken by Lipscomb and King¹⁷ in 1949, at which time the anisotropic tensor description was not known. They derived an expression for the calculation of the intensity of X-ray scattering from rigid groups undergoing hindered rotation. The assumption was made that the rigid group rotates about a fixed axis with a sinusoidal hindering potential. The method was used by Prewitt¹⁸ in 1965 when he treated the cyclopentadienyl rings in $[\text{C}_5\text{H}_5\text{FeS}]_4$ as hindered rotors. He assumed that each cyclopentadienyl ring formed a regular, planar pentagon with D_{5h} symmetry and that the motion of the atoms is predominantly in the plane of the ring. The other motions of the group were described by a single isotropic temperature factor B , this being a modification of the original King and Lipscomb model. This model was refined by means of two matrices in a block diagonal refinement; one 4×4 matrix for x, y, z and B , and a 2×2 matrix for the radius of the ring, and the relative barrier. Use of this model gave cyclopentadienyl C-C bond lengths in good agreement with the average C-C bond

length obtained when corrections for thermal motion had been made for the individual atom models. It also required far fewer parameters to describe the system although this resulted in a higher R factor; 6.9% for the hindered rotor as compared to 5.8% for the individual atom. Also, the hindered rotor model had the advantage that it placed the electron density on the arc of a circle giving a curved distribution rather than attempting to describe the distribution in a rectilinear manner as in the conventional isotropic and anisotropic models. However, there were disadvantages. The particular blocking of the least squares matrix was unfortunate in that the correlation between the two parameters associated with a description of the thermal motion was ignored. This caused severe over-shifting and necessitated the introduction of damping factors. Because the ring orientations were well defined from the previous individual atom refinement, these were not refined and the programme was developed with a limited scope. However, the utility of the model had been demonstrated for a routine organometallic structure.

In order to refine n parameters by means of full matrix least squares refinement, a symmetrical matrix containing $(n/2)(n+1)$ unique elements is required. Both the computing time required and the storage demands for

a least squares cycle are proportional² to $(n/2)(n+1)$. Thus, any decrease in the number of parameters to be refined will result in a proportionately greater decrease in time and storage requirements and in the computing cost. It was therefore deemed more economical to develop a programme that allowed the ring orientations to be refined rather than to obtain these parameters from a complete individual atom refinement.

In the following work, the rigid body expressions available in the rigid body modification of the basic full matrix least squares programme SFLS5 by C.T. Prewitt (1966) were used in conjunction with Lipscomb and King's¹⁷ hindered rotor expressions. SFLS5 was then further modified to calculate the contribution to the structure factors by a group undergoing hindered rotation and to do a full matrix least squares refinement on the input parameters.

THEORETICAL

Notation

g_i	complex scattering power of the i^{th} atom
\bar{g}	average complex scattering power of atoms in a ring
$P(\theta')$	probability function of the ring at an angle θ'
\underline{h}	reciprocal lattice vector corresponding to the Miller indices h, k, l
\underline{r}	instantaneous position vector of the atom which may be resolved into two components:
\underline{k}	vector from the real lattice origin to the centre of rotation, and
\underline{v}	vector from the centre of rotation to the instantaneous position of the atom
V_0	potential barrier to rotation
V	potential at rotation angle θ'
n	number of minima in the potential
θ'	instantaneous rotation angle measured from the x axis of the internal Cartesian co-ordinate system in the rigid body

$x_{c_i}, y_{c_i}, z_{c_i}$

fractional co-ordinates of the
centre of gravity of the i^{th} rigid
body

ξ

rotation angle at which the potential
is a minimum

D, E

angles (in radians) defining the
orientation of the plane of the
ring. The direction cosines (with
respect to the chosen internal
Cartesian co-ordinate system) of
the plane are:

$$m_1 = \cos D \sin E$$

$$m_2 = -\sin D$$

$$\text{and } m_3 = \cos D \cos E$$

R

radius of the rigid body in \AA

$\left. \begin{array}{l} X(i) \\ Y(i) \\ Z(i) \end{array} \right\}$

Cartesian co-ordinates in \AA of i^{th} atom

[L]

transformation matrix from the internal
co-ordinate system of the rigid body
to conventional crystallographic
co-ordinates

q_1 to q_6

non-zero elements of the matrix [L]

$$q_1 = 1. / (a \sin \beta \sin \gamma^*)$$

$$q_2 = \cot \gamma^* / b \sin \alpha$$

$$q_3 = - \frac{1}{c} \left[\cos \alpha \cos \gamma^* - \cot \beta \frac{1}{\sin \gamma^*} \right]$$

$$q_4 = 1./ (b \sin \alpha)$$

$$q_5 = -c \cot \alpha$$

$$q_6 = 1./c$$

$g p_i$	population factor of the i^{th} ring
B_i	Debye temperature factor for the i^{th} ring
θ	conventional Bragg angle
λ	wavelength of X-radiation used
f_i	form factor for atoms of the i^{th} ring
T	absolute temperature
k	Boltzman's constant
i	imaginary number = $\sqrt{-1}$
B_d	relative barrier $B_d = V_o/2kT$

The rigid body expression of M.J. Bennett and B.M. Foxman assumes a rigid ring system based on a regular planar polygon of order n possessing D_{nh} symmetry. A right handed Cartesian co-ordinate system is defined such that the origin lies at the centre of the polygon, with the x and z axes being parallel to the a^* and c axes. If the ring lies in the XY plane, then the co-ordinates

of the first atom can be defined as $(R, 0, 0)$. The position of the i^{th} atom is then

$$\begin{aligned} X(i) &= R \cos 2\pi(i-1)/n \\ Y(i) &= R \sin 2\pi(i-1)/n \end{aligned} \quad [4]$$

$$\text{and } Z(i) = 0$$

with respect to the Cartesian co-ordinate system. The direction cosine of the ring normal in its standard orientation is $(0, 0, 1)$. Rotating this polygon about the z axis, the i^{th} vertex becomes

$$\begin{aligned} X(i) &= R \cos(\xi + 2\pi(i-1)/n) \\ Y(i) &= R \sin(\xi + 2\pi(i-1)/n) \\ Z(i) &= 0 \end{aligned} \quad [5]$$

The direction cosines of the plane are still $(0, 0, 1)$. If this is followed by a rotation of D radians about the x axis and E radians about the y axis, the direction cosines of the plane become $(\cos D \sin E, -\sin D, \cos D \cos E)$. The co-ordinates of the i^{th} atom are then

$$\begin{aligned} X(i) &= R[\cos(\xi + 2\pi(i-1)/n) \cos E \\ &\quad + \sin(\xi + 2\pi(i-1)/n) \sin D \sin E] \\ Y(i) &= R[\sin(\xi + 2\pi(i-1)/n) \cos D] \\ Z(i) &= R[-\cos(\xi + 2\pi(i-1)/n) \sin E \\ &\quad + \sin(\xi + 2\pi(i-1)/n) \cos E \sin D] \end{aligned} \quad [6]$$

Transformation to crystallographic co-ordinates using

matrix [L] plus translational correction gives the fractional co-ordinates of the i^{th} atom as

$$\begin{aligned}
 x(i) &= q_1 R[\cos(\xi + 2\pi(i-1)/n) \cos E \\
 &\quad + \sin(\xi + 2\pi(i-1)/n) \sin D \sin E] + x_c \\
 y(i) &= q_2 R[\cos(\xi + 2\pi(i-1)/n) \cos E \\
 &\quad + \sin(\xi + 2\pi(i-1)/n) \sin D \sin E] \\
 &\quad + q_4 R[\sin(\xi + 2\pi(i-1)/n) \cos D] + y_c \\
 \text{and } z(i) &= q_3 R[\cos(\xi + 2\pi(i-1)/n) \cos E \\
 &\quad + \sin(\xi + 2\pi(i-1)/n) \sin D \sin E] \\
 &\quad + q_5 R[\sin(\xi + 2\pi(i-1)/n) \cos D] \\
 &\quad + q_6 R[-\cos(\xi + 2\pi(i-1)/n) \sin E \\
 &\quad + \sin(\xi + 2\pi(i-1)/n) \cos E \sin D] + z_c
 \end{aligned} \tag{7}$$

where q_i are the elements of the transformation matrix [L]. Replacing the fixed angle ξ by the continuous variable θ' , the dot product, $2\pi \underline{h} \cdot \underline{v}$, used in later expressions, becomes

$$2\pi \underline{h} \cdot \underline{v} = c \cos \theta' + d \sin \theta' \tag{8}$$

$$\begin{aligned}
 \text{where} \quad c &= 2\pi R[(h q_1 + k q_2 + l q_3) \cos E - \\
 &\quad l q_6 \sin E]
 \end{aligned} \tag{9}$$

$$\begin{aligned}
 \text{and } d &= 2\pi R[(h q_1 + k q_2 + l q_3) \sin D \sin E \\
 &\quad + (k q_4 + l q_5) \cos D + l q_6 \cos E \sin D].
 \end{aligned}$$

$$\text{If } a' \text{ is defined as } a' = \sqrt{c^2 + d^2} \quad [10]$$

$$\begin{aligned} \text{and } \eta &= \arccos c/a' \\ &= \arcsin d/a' \end{aligned} \quad [11]$$

then the dot product $2\pi \underline{h} \cdot \underline{v}$ becomes

$$2\pi \underline{h} \cdot \underline{v} = a' \cos(\theta' - \eta). \quad [12]$$

These rigid body expressions can now be used in calculating the contribution of groups in hindered rotation to the structure factor. The derivation parallels that of Lipscomb and King¹⁷.

The contribution of the hindered rotor groups to the structure factor can be expressed as

$$F_{hkl} = \sum_i g_i \quad [13]$$

the summation being over all atoms in the unit cell which are members of a group undergoing hindered rotation. The instantaneous complex scattering power of an atom in a ring is given by

$$\begin{aligned} g_i &= f_i \exp[2\pi i \underline{h} \cdot \underline{r}] \\ &= f_i \exp[2\pi i \underline{h} \cdot \underline{k}] \exp[2\pi i \underline{h} \cdot \underline{v}]. \end{aligned} \quad [14]$$

If the probability distribution is given by $P(\theta')$, then

$$\begin{aligned} \bar{g} &= f_i \exp[2\pi i \underline{h} \cdot \underline{k}] \cdot \\ &\quad \int_0^{2\pi} P(\theta') \exp[2\pi i \underline{h} \cdot \underline{v}] d\theta'. \end{aligned} \quad [15]$$

Using a potential function

$$V = -\frac{1}{2} V_0 \cos n(\theta' - \xi) \quad [16]$$

and assuming a Maxwell-Boltzman distribution for the probability function, $P(\theta')$, then

$$P(\theta') = \frac{\exp[V_o \cos n(\theta' - \xi)/2kT]}{\int_0^{2\pi} \exp[V_o \cos n(\theta' - \xi)/2kT] d\theta'} \quad [17]$$

Substituting this expression for $P(\theta')$ into equation [15] and using the dot product expression of [12], one obtains

$$\bar{g} = f_i \exp[2\pi i \underline{h} \cdot \underline{k}] \quad [18]$$

$$\frac{\int_0^{2\pi} \exp[Bd \cos n(\theta' - \xi) + ia' \cos(\theta' - \eta)] d\theta'}{\int_0^{2\pi} \exp[Bd \cos n(\theta' - \xi)] d\theta'}$$

where $Bd = V_o/2kT$.

Defining $M_n^\xi(a', Bd) =$

$$\frac{\int_0^{2\pi} \exp[Bd \cos n(\theta' - \xi) + ia' \cos(\theta' - \xi)] d\theta'}{\int_0^{2\pi} \exp[Bd \cos n(\theta' - \xi)] d\theta'} \quad [19]$$

then \bar{g} becomes

$$\bar{g} = f_i \exp[2\pi i \underline{h} \cdot \underline{k}] M_n^\xi(a', Bd). \quad [20]$$

$M_n^\xi(a', Bd)$ can be expanded into an infinite series.

Watson²⁰ gives the following expression:

$$\begin{aligned} \exp[s \cos \tau] &= I_0(s) + 2 \sum_{m=1}^{\infty} I_m(s) \cos m\tau \\ &= \sum_{m=0}^{\infty} \epsilon_m I_m(s) \cos m\tau \end{aligned} \quad [21]$$

where $\epsilon_m = 1$ when $m = 0$, and $\epsilon_m = 2$ when $m \neq 0$. $J_m(s)$ is a Bessel function of order m and $I_m(s)$ is a hyperbolic Bessel function of order m . Using these expressions, the numerator of $M_n^\xi(a', Bd)$ becomes

$$\begin{aligned} &\int_0^{2\pi} \left[\sum_p \epsilon_p I_p(Bd) \cos p(n\theta' - \xi) \right] \left[\sum_m \epsilon_m I_m(ia') \cos m(\theta' - \xi) \right] d\theta' \\ &= \sum_p \sum_m \epsilon_p \epsilon_m I_p(Bd) I_m(ia') \int_0^{2\pi} \cos p(n\theta' - \xi) \cos m(\theta' - \xi) d\theta'. \end{aligned} \quad [22]$$

By expanding the cosine term the numerator can be written as

$$\begin{aligned} &\sum_p \sum_m \epsilon_p \epsilon_m I_p(Bd) I_m(ia') \cdot \\ &\int_0^{2\pi} (\cos pn\theta' \cos p\xi \cos m\theta' \cos m\xi \\ &+ \sin pn\theta' \sin p\xi \cos m\theta' \cos m\xi \\ &+ \cos pn\theta' \cos p\xi \sin m\theta' \sin m\xi \\ &+ \sin pn\theta' \sin p\xi \sin m\theta' \sin m\xi) d\theta'. \end{aligned} \quad [23]$$

Making use of (1) summation of integral equalling the integral of summation, and (2) the orthogonality of sines and cosines, this expression simplifies to:

$$2\pi \sum_p \epsilon_p I_p(Bd) I_{pn}(ia') (\cos p\xi \cos pn\eta + \sin p\xi \sin pn\eta). \quad [24]$$

Using the relationship

$$J_m(-is) = (-i)^m I_m(s) \quad [25]$$

given by Watson²⁰ the numerator becomes

$$2\pi \sum_p \epsilon_p i^{pn} I_p(Bd) J_{pn}(a') \cos(pn(\xi - \eta)) \quad [26]$$

and by the same method the denominator is $2\pi I_0(Bd)$.

$M_n^\xi(a', Bd)$ then becomes

$$\frac{\sum_p \epsilon_p i^{pn} I_p(Bd) J_{pn}(a') \cos(pn(\xi - \eta))}{I_0(Bd)}. \quad [27]$$

Using this expression, the contribution to the structure factor from the hindered rotors can be written as

$$F_{hkl} = \sum_i f_i g_{p_i} \exp[-B_i \frac{\sin^2 \theta}{\lambda^2}] \exp[2\pi i \underline{h} \cdot \underline{k}] M_{n_i}^{\xi_i}(a'_i, Bd_i) \quad [28]$$

the summation being over all the rings undergoing hindered rotation.

Programming

The programme SFLS5 was modified to calculate the contribution of a hindered rotor to the structure factor. These modifications were mainly in the addition of two new subroutines ROTOR and BESSEL, as well as modifications to the input and output. It can be seen that there are two

possible routes by which to calculate the M function:

(1) by means of the integral expression, equation [19], or

(2) by use of the infinite series summation form (Bessel functions), equation [27].

The first of these, the integral form, can be calculated by means of Weddel's Rule¹⁹. If one divides the range of integration into 7 intervals, then

$$\int_{x_0}^{x_0+6h} \phi(x) dx = \frac{3h}{10} [y_0 + 5y_1 + y_2 + 6y_3 + y_4 + 5y_5 + y_6] \quad [29]$$

This can be expanded for use with $(6n+1)$ intervals, where n is an integer. An integral of the form $\int_a^b e^{f(x)} g(x) dx$ can be expanded into a series of integrals where $f(x)$ and $g(x)$ are two functions of x .

$$\begin{aligned} \int_a^b e^{f(x)} g(x) dx &= \int_a^b [1 + f(x) + \frac{f^2}{2!}(x) + \dots] g(x) dx \\ &= \int_a^b g(x) dx + \int_a^b f(x) g(x) dx \\ &\quad + \frac{1}{2!} \int_a^b f^2(x) g(x) dx + \dots \end{aligned} \quad [30]$$

which hopefully are easier to integrate. Thus the $M_n^\xi(a', Bd)$ function can be broken up into a series of integrals which can be calculated by means of Weddel's Rule, or integrated analytically. However, it was found that even integrals

of the form $\frac{1}{2} \int g(x) f^2(x) dx$ became extremely complicated (Figure 1) and required a minimum of 67 intervals. Since this is an extremely slow method of calculation, it was used only as a check on the results of the infinite Bessel function series calculations.

The infinite series form of $M_n^\xi(a', Bd)$ is

$$M_n^\xi(a', Bd) = \sum_p \epsilon_p i^{pn} J_{pn}(a') \frac{I_p(Bd)}{I_0(Bd)} \cos(pn(\xi - \eta)) \quad [27]$$

$$\text{where } I_p(Bd) = \sum_m \frac{(Bd/2)^{2m+p}}{m! (m+p)!} \quad [31]$$

$$\text{and } J_{pn}(a') = \sum_m \frac{(-1)^m (a'/2)^{2m+pn}}{m! (m+pn)!} \quad [32]$$

$J_{pn}(a')$ being a Bessel summation²⁰ and $I_p(Bd)$ a hyperbolic Bessel summation. It can be seen that whether any term in the summation is real or imaginary depends completely on the value of i^{pn} ; all other functions are real. Thus, if n , the order of the ring, is even, pn is always even and i^{pn} is ± 1 ; $M_n^\xi(a', Bd)$ is then a real summation. If n is odd, then pn alternates even and odd as p increases causing i^{pn} to alternate among ± 1 and $\pm i$. $M_n^\xi(a', Bd)$ is a complex function which can be broken up into real and imaginary contributions:

Figure 1

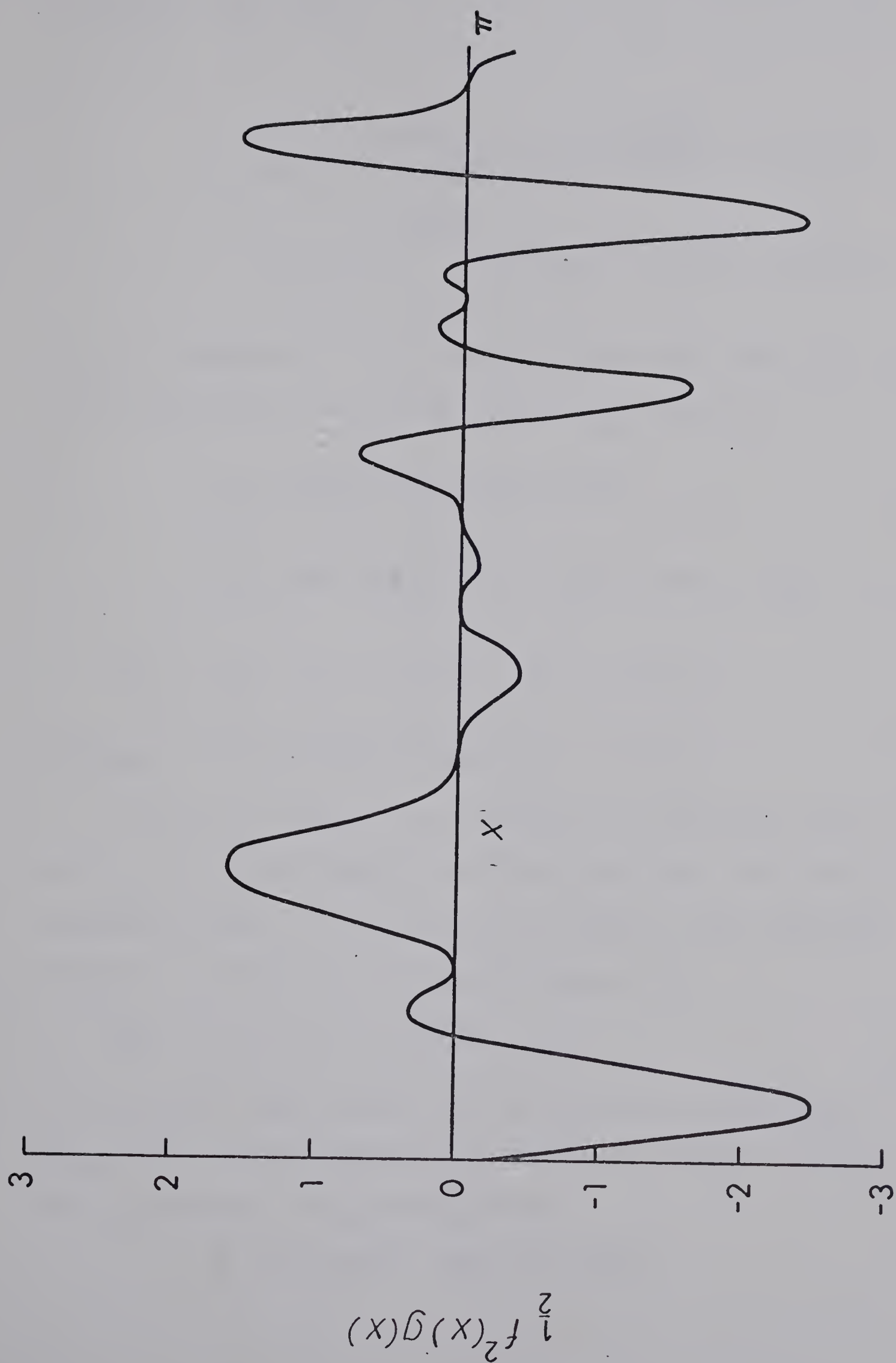
$$\frac{1}{2} g(x) f^2(x) \text{ versus } x$$

$$\text{where } g(x) = \cos[a' \cos(x-\eta)]$$

$$f(x) = B d \cos n(x-\xi)$$

The following values were used:

$$\begin{array}{lll} h = 1 & k = 2 & \ell = 1 \\ \xi = 0.23 & D = 0.60 \text{ radians} & \\ n = 5 & E = 0.028 \text{ radians} & \\ Bd = 2.5 & R = 1.20 \text{ \AA} & \\ q_1 = 0.10 & q_2 = 0.20 & q_3 = 0.10 \\ q_4 = 0.10 & q_5 = 0.27 & q_6 = 0.10 \end{array}$$



$$\begin{aligned}
 M_n^\xi(a', Bd) &= M_R + iM_I \\
 &= \sum_{p \text{ even}} (-1)^{pn/2} \epsilon_p J_{pn}(a') \frac{I_p(Bd)}{I_0(Bd)} \cos pn(\xi - \eta) \quad [33] \\
 &\quad + i \sum_{p \text{ odd}} (-1)^{\frac{pn-1}{2}} \epsilon_p J_{pn}(a') I_p(Bd) \cos pn(\xi - \eta).
 \end{aligned}$$

$F_{hk\ell}$ is complex and can also be broken down into real and imaginary terms, generally called $A_{hk\ell}$ and $B_{hk\ell}$.

$$\begin{aligned}
 &= \sum_i f_i \exp(2\pi i \underline{h} \cdot \underline{k}) M_n^\xi(a', Bd) \\
 &= \sum_i f_i (\cos 2\pi \underline{h} \cdot \underline{k} + i \sin 2\pi \underline{h} \cdot \underline{k}) (M_R + iM_I) \quad [34]
 \end{aligned}$$

$$\therefore A_{hk\ell} = \sum_i f_i (M_R \cos 2\pi \underline{h} \cdot \underline{k} - M_I \sin 2\pi \underline{h} \cdot \underline{k}) \quad [35]$$

$$\text{and } B_{hk\ell} = \sum_i f_i (M_I \cos 2\pi \underline{h} \cdot \underline{k} + M_R \sin 2\pi \underline{h} \cdot \underline{k}) \quad [36]$$

Thus for n odd it is necessary to calculate both M_R and M_I as well as the sin and cos terms even for the centrosymmetric case. For the special case of even membered rings in a centrosymmetric space group then

$$A_{hk\ell} = \sum_i f_i M_R \cos 2\pi \underline{h} \cdot \underline{k} \quad [37]$$

In a centric space group with no anomalous dispersion $B_{hk\ell} = 0$. This property was used as a further check on the calculation of M_R and M_I since

$$M_I \cos 2\pi \underline{h} \cdot \underline{k} = -M_R \sin 2\pi \underline{h} \cdot \underline{k} . \quad [38]$$

In order to do a least squares refinement, it is necessary to calculate the derivatives of the $M_n^\xi(a', Bd)$ function. The ring is described by means of ten parameters, and analytical expressions for the derivatives of $F_{hk\ell}$ with respect to each parameter were obtained. Those of the population factor gp , the co-ordinates of the centre of gravity x_c, y_c, z_c and the overall temperature parameter, B , were straightforward, since they are not involved in the calculation of the M function.

Since

$$F_{hk\ell} = \sum gp_i f_i \exp\left[\frac{-B_i \sin^2 \theta}{\lambda^2}\right] \exp[2\pi i \underline{h} \cdot \underline{k}] M_n^\xi(a', Bd) \quad [28]$$

then the derivative with respect to the population factor becomes

$$\frac{\partial F}{\partial (gp)_i} = f_i \exp\left[\frac{-B_i \sin^2 \theta}{\lambda^2}\right] \exp[2\pi i \underline{h} \cdot \underline{k}] M_n^\xi(a', Bd). \quad [39]$$

Using the expansion of $\underline{h} \cdot \underline{k}$ to $(hx_{c_i} + ky_{c_i} + z_{c_i})$ and taking the derivative of F with respect to x_{c_i}

$$\begin{aligned} \frac{\partial F}{\partial x_{c_i}} &= gp_i f_i \exp\left[\frac{-B_i \sin^2 \theta}{\lambda^2}\right] (2\pi i h) \\ &\quad \exp[2\pi i \underline{h} \cdot \underline{k}] M_n^\xi(a', Bd). \end{aligned} \quad [40]$$

Analogous expressions can be written for $\frac{\partial F}{\partial y_{c_i}}$ and $\frac{\partial F}{\partial z_{c_i}}$

where $2\pi i h$ is replaced by $2\pi i k$ and $2\pi i l$ respectively. Differentiation with respect to the overall temperature factor gives

$$\frac{\partial F}{\partial B} = - \frac{\sin^2 \theta}{\lambda^2} \exp \left[\frac{B_i \sin^2}{\lambda^2} \right] f_i \exp[2\pi i \underline{h} \cdot \underline{k}] M_n^\xi(a', Bd). \quad [41]$$

The remaining five parameters are used in the calculation of $M_n^\xi(a', Bd)$ and the analytical expressions for their derivatives are somewhat more complicated.

The relative barrier, Bd , is found only in the terms $I_o(Bd)$ and $I_p(Bd)$. The derivative of $M_n^\xi(a', Bd)$ with respect to Bd is given by

$$\begin{aligned} \frac{\partial M_n^\xi(a', Bd)}{\partial Bd} &= \frac{1}{I_o^2(Bd)} [I_o^2(Bd) \sum_p i^{pn} J_{pn}(a') \cdot \\ &\quad \cos pn(\xi - \eta) \frac{\partial I_p(Bd)}{\partial Bd} \\ &\quad - \frac{\partial I_o(Bd)}{\partial Bd} \sum_p \epsilon_p i^{pn} J_{pn}(a') \cdot \\ &\quad I_p(Bd) \cos pn(\xi - \eta)]. \end{aligned} \quad [42]$$

Since

$$I_p(Bd) = \sum_m \frac{(Bd/2)^{2m+p}}{m! (m+p)!} \quad [31]$$

then

$$\frac{\partial I_p(Bd)}{\partial Bd} = \sum_m \frac{2m+p}{Bd} \frac{(Bd/2)^{2m+p}}{m! (m+p)!} \quad [43]$$

That is, each term in this summation is $\frac{2m+p}{Bd}$ times the original Bessel term. An exactly similar expression occurs for

$$\frac{\partial I_0(Bd)}{\partial Bd} = \sum_m \frac{2m}{Bd} \frac{(Bd/2)^{2m}}{(m!)^2} \quad [44]$$

It was for the calculation of such modified Bessel terms that the subroutine BESSEL was written rather than using the computer library function.

In the expression for $M_n^\xi(a', Bd)$, only the term $J_{pn}(a')$ is a function of the radius, R .

$$\frac{\partial M_n(a', Bd)}{\partial R} = \sum_p i^{pn} \epsilon_p \frac{I_p(Bd)}{I_0(Bd)} \cos n(\xi - \eta) \frac{\partial J_{pn}(a')}{\partial R} \quad [45]$$

Since

$$J_{pn}(a') = \sum_m \frac{(-1)^m (a'/2)^{2m+pn}}{m! (m+pn)!} \quad [32]$$

then

$$\begin{aligned} \frac{\partial J_{pn}(a')}{\partial R} &= \sum_m \frac{(-1)^m (a'/2)^{2m+pn-1}}{m! (m+pn)!} \frac{\partial (a')}{\partial R} (2m+pn) \\ &= \sum_m \frac{(-1)^m (a''/2)^{2m+p} (2m+pn)}{m! (m+pn)!} \end{aligned} \quad [46]$$

where $a'' = a'/R$.

As in the case of $\frac{\partial M_n^\xi(a', Bd)}{\partial Bd}$, this is a modified Bessel summation.

The derivative $\frac{\partial M_n^\xi(a', Bd)}{\partial \xi}$ is simply

$$\sum_p -pn i^{pn} \epsilon_p J_{pn}(a') \frac{I_p(Bd)}{I_0(Bd)} \sin(pn(\xi - \eta)). \quad [47]$$

The most complex expressions for the derivatives of $F_{hk\ell}$ arises from those with respect to the orientation angles D and E. In each term of the $M_n^\xi(a', Bd)$ summation $J_{pn}(a')$ and $\cos(pn(\xi - \eta))$ are functions of both D and E. Dealing first with the derivative with respect to D

$$\begin{aligned} \frac{\partial M_n^\xi(a', Bd)}{\partial D} = & \sum_p i^{pn} \frac{I_p(Bd)}{I_0(Bd)} \left[J_{pn}(a') \sin pn(\xi - \eta) \frac{\partial \eta}{\partial D} \right. \\ & + \cos pn(\xi - \eta) \sum_m \frac{(-1)^m (a'/2)^{2m+pn}}{m! (m+pn)!} \cdot \\ & \left. \frac{2m+pn}{a'} \cdot \frac{\partial a'}{\partial D} \right]. \end{aligned} \quad [48]$$

where $\frac{\partial a'}{\partial D} = \frac{d}{a'} \cdot \frac{\partial d}{\partial D}$

$$\frac{\partial \eta}{\partial D} = - \frac{c}{a' \sqrt{a'^2 - c^2}} \frac{\partial a'}{\partial D} \quad [49]$$

$$\text{and } \frac{\partial d}{\partial D} = 2\pi R[(h q_1 + k q_2 + \ell q_3) \cos D \sin E \\ - (k q_4 + \ell q_5) \sin D + \ell q_6 \cos D \cos E] .$$

The expressions for d, a' and η are given in equations [9], [10] and [11] respectively.

The derivatives of the $M_n^\xi(a', Bd)$ function with respect to E are calculated in a similar manner with the result:

$$\frac{\partial M_n(a', Bd)}{\partial E} = \sum_p i^{pn} \frac{I_p(Bd)}{I_0(Bd)} \left[pn J_{pn}(a') \sin pn(\xi - \eta) \frac{\partial \eta}{\partial E} \right. \\ \left. + \cos pn(\xi - \eta) \sum \frac{(-1)^m (a'/2)^{2m+pn}}{m! (m+pn)!} \cdot \frac{2m+pn}{a'} \cdot \frac{\partial a'}{\partial E} \right] \quad [50]$$

$$\text{where } \frac{\partial a'}{\partial E} = -2 R[(h q_1 + k q_2 + \ell q_3) \sin E + \ell q_6 \cos E]$$

$$\frac{\partial d}{\partial E} = 2\pi R[(h q_1 + k q_2 + \ell q_3) \sin E \cos E \\ - \ell q_6 \sin D \sin E] \quad [51]$$

$$\text{and } \frac{\partial \eta}{\partial E} = \frac{d}{(a')^2} \cdot \frac{1}{\sqrt{1 - (d/a')^2}} \frac{\partial a'}{\partial E}$$

The derivatives of the F_{hkl} with respect to the last five parameters can be calculated from the derivatives of M by application of the chain rule to equation [28].

Once written, this programme was tested by several different methods. The M function itself was calculated by two different methods:

(1) the infinite series, the test for convergence being that the final term was less than 0.01% of the summation, and

(2) the integral method using Weddel's Rule¹⁹.

The results were also compared with those given by Lipscomb and King¹⁷ and agreed to within the four figures given.

The calculations of derivatives were also checked by several methods:

(1) by calculating $\frac{\partial M}{\partial \phi}$ analytically from the infinite series form of M , ϕ being one of the variables D , E , ξ , R , or Bd ,

(2) by calculating M at various values of ϕ , plotting the results and obtaining the slope, and

(3) by means of the numerical analysis technique of forward differences²¹.

Once these tests indicated that the programming was correct, the hindered rotor calculations were tested on several rings found in published crystal structures. These were (1)

benzene¹², ruthenocene²², (3) dicarbonyl- π -cyclohepta-
trienyl- σ -pentafluorophenyl molybdenum²³, (4) cyclo-
pentadienyl manganese tricarbonyl²⁴, and (5) benzene-
chromium tricarbonyl²⁵.

Diagrams of these molecules produced using the
published parameters of the individual atom refinements
are shown in Figures 2 to 6. With the exception of
ruthenocene, the drawings show each molecule projected
onto the plane of the ring. The diagrams were all
produced by C.K. Johnson's ORTEP programme and give
qualitative indications of librations about the ring axes.

Figure 2

The structure of ruthenocene
G.L. Hardgrove and D.H. Templeton
Acta Cryst., 12, 28 (1959)

Figure 3

The structure of benzene
E.G. Cox, D.W.J. Cruickshank and J.A.S. Smith
Proc. Roy. Soc., A247, 1 (1958)

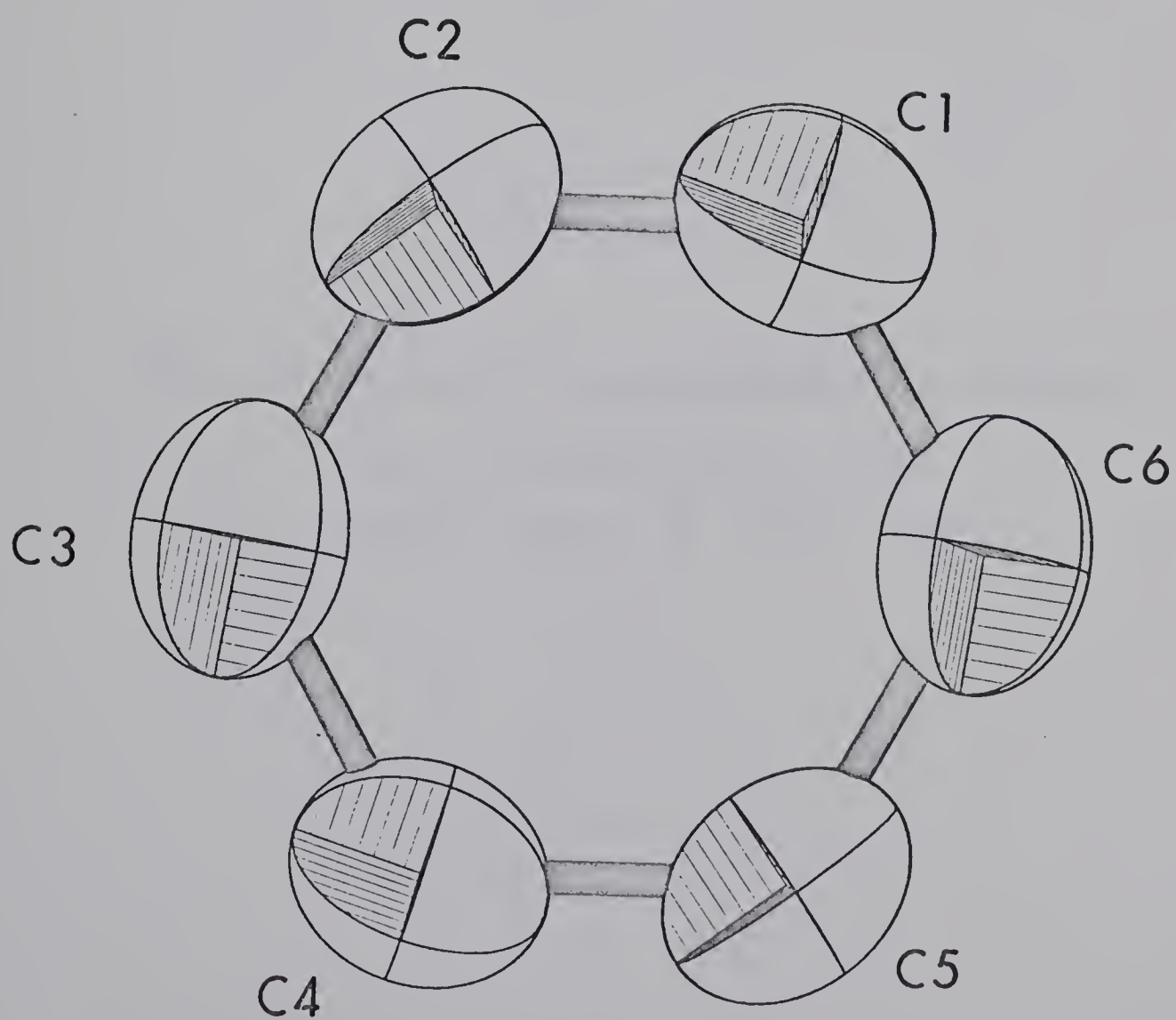
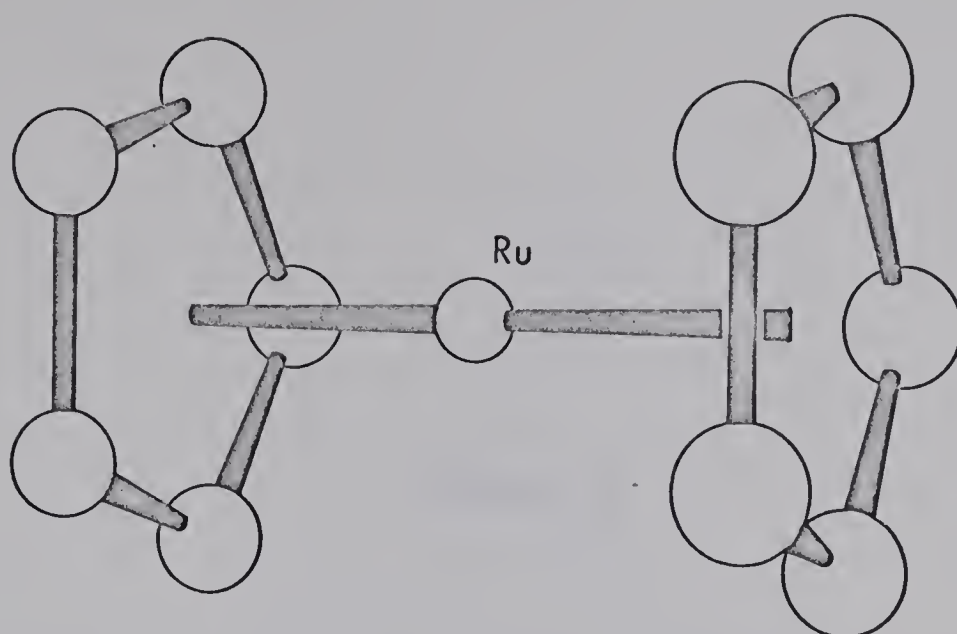


Figure 4

The structure of cyclopentadienyl manganese tricarbonyl

A.F. Berndt and R.E. Marsh

Acta Cryst., 16, 118 (1963)

Figure 5

The structure of benzenechromium tricarbonyl

M.F. Bailey and L.F. Dahl

Inorg. Chem., 4, 1314 (1965)

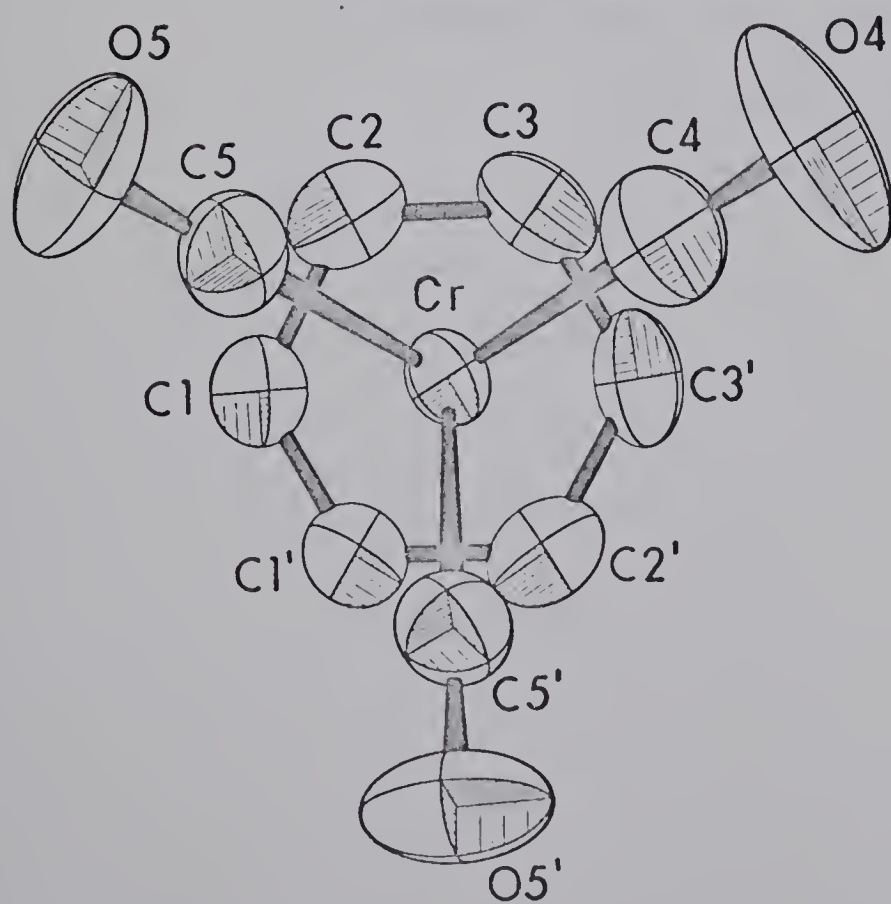
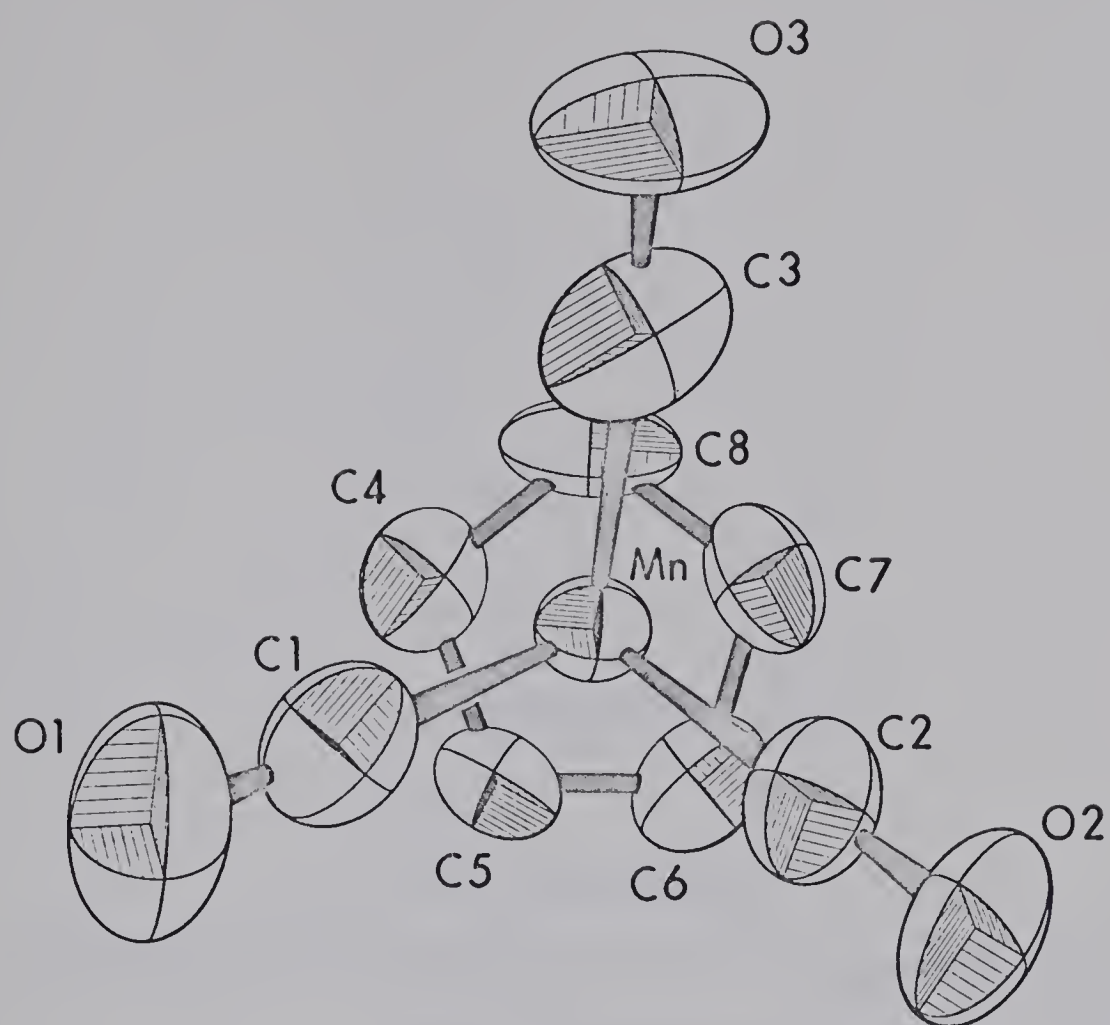
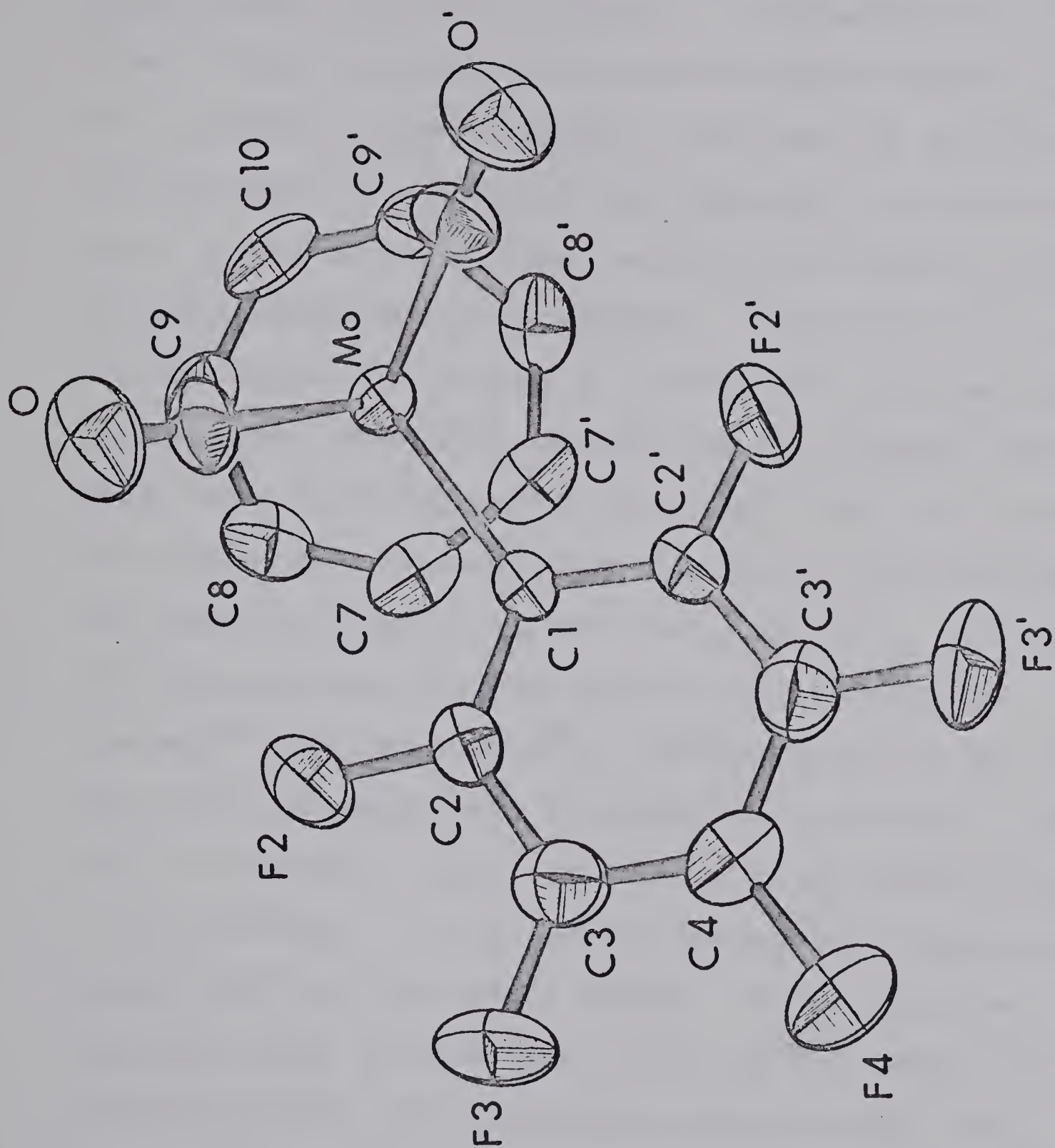


Figure 6

The structure of
dicarbonyl- π -cycloheptatrienyl- σ -pentafluoro-
phenyl molybdenum

M.R. Churchill and T.A. O'Brien

J. Chem. Soc. (A), 1110, (1969)



Results

The refined parameters for the hindered rotor calculations are given in Table I. Parameters for other atoms of the structure did not change significantly from the individual atom refinement values and are not listed. In each case the results of the individual atom refinement given in the published paper were used to obtain the input to the hindered rotor calculation. The centre of gravity and the orientation angles D , E and ξ were obtained from the programme MMR (M.J. Bennett and B.M. Foxman, 1967) which uses the positions of the atoms in the ring as input. The radius was calculated from the average distance between the centre of gravity and the atom positions. The barrier, B_d , was calculated from NMR results where available (benzene²⁷ and ruthenocene²⁸); otherwise set at $B_d = 2$. The isotropic temperature parameter for the ring, B , was set to be slightly greater than that of the metal to which it was attached. In all cases, the hydrogen atoms were placed $1.0\overset{\circ}{\text{\AA}}$ from the carbon atoms. The results are discussed under the headings of (1) refined radius, (2) relative barrier, (3) temperature factor of the ring, (4) the R factor, and (5) time and storage requirements.

Table I

Refined parameters of hindered rotor groups

	$\pi\text{C}_5\text{H}_5\text{Mn}(\text{CO})_3$	benzene	$\pi\text{C}_6\text{H}_6\text{Cr}(\text{CO})_3$	Ruthenocene 1 2		$\pi\text{C}_7\text{H}_7\text{Mo}(\text{CO})_3$ $\sigma\text{-C}_6\text{F}_5$
x	0.3010(7)	0.0*	0.3785(6)	0.3632(18)	0.1060(9)	0.2024(7)
y	-0.1882(10)	0.0*	0.2500	0.2500*	0.2500*	0.2500*
z	0.1669(7)	0.0*	-0.2277(5)	0.6292(9)	0.3797(9)	0.1327(4)
B(\AA^2)	3.67(17)	4.64(11)	3.42(9)	2.14(30)	2.49(33)	3.13(13)
Bd	1.50(12)	1.94(11)	1.97(10)	4.9(20)	5.2(33)	1.94(14)
R($^\circ$)	1.211(6)	1.391(3)	1.415(3)	1.239(13)	1.235(14)	1.626(5)
D radians	4.024(7)	6.053(3)	6.279*	0.0*	0.0*	3.142*
E	2.998(9)	0.821(3)	2.975(3)	0.510(13)	0.531(14)	0.949(4)
ξ	3.018(9)	1.741(3)	0.524*	0.0*	0.0*	0.449(1)
R(%)	8.9	10.1	4.5	8.7	8.7	5.7

* not refined - constrained symmetry element.

Standard deviations in parentheses refer to the last digit listed.

DISCUSSION

Radius

In the refinement of crystal structures by means of the thermal ellipsoid description, the origin of an ellipse tends to refine to the centre of gravity of the local electron density. For the case of an atom moving along a radial track, this centre will not be on the track. If the radial track is the arc AQB of Figure 7, then the centre of gravity would lie at point P. The radius of the ring would show an apparent shortening of the ring by an amount PQ. The ratio of the apparent radius to the true radius would be

$$\frac{\int_0^{\pi} P(\theta') \cos \theta' d\theta'}{\int_0^{\pi} P(\theta') d\theta'} \quad [52]$$

In this case a sinusoidal distribution has been used and $P(\theta')$ can be calculated at various θ' angles as a function of the relative barrier to rotation B_d and the order of the ring. The results are plotted in Figure 8.

For benzene, with a relative barrier from the hindered rotor refinement of 1.95, the ratio of the apparent radius to that of the correct radius would be 0.988. The radius

Figure 7

Radial track of rotation group

where r radius of ring

OP = apparent radius

$P(\theta)$ probability distribution

 along radial track

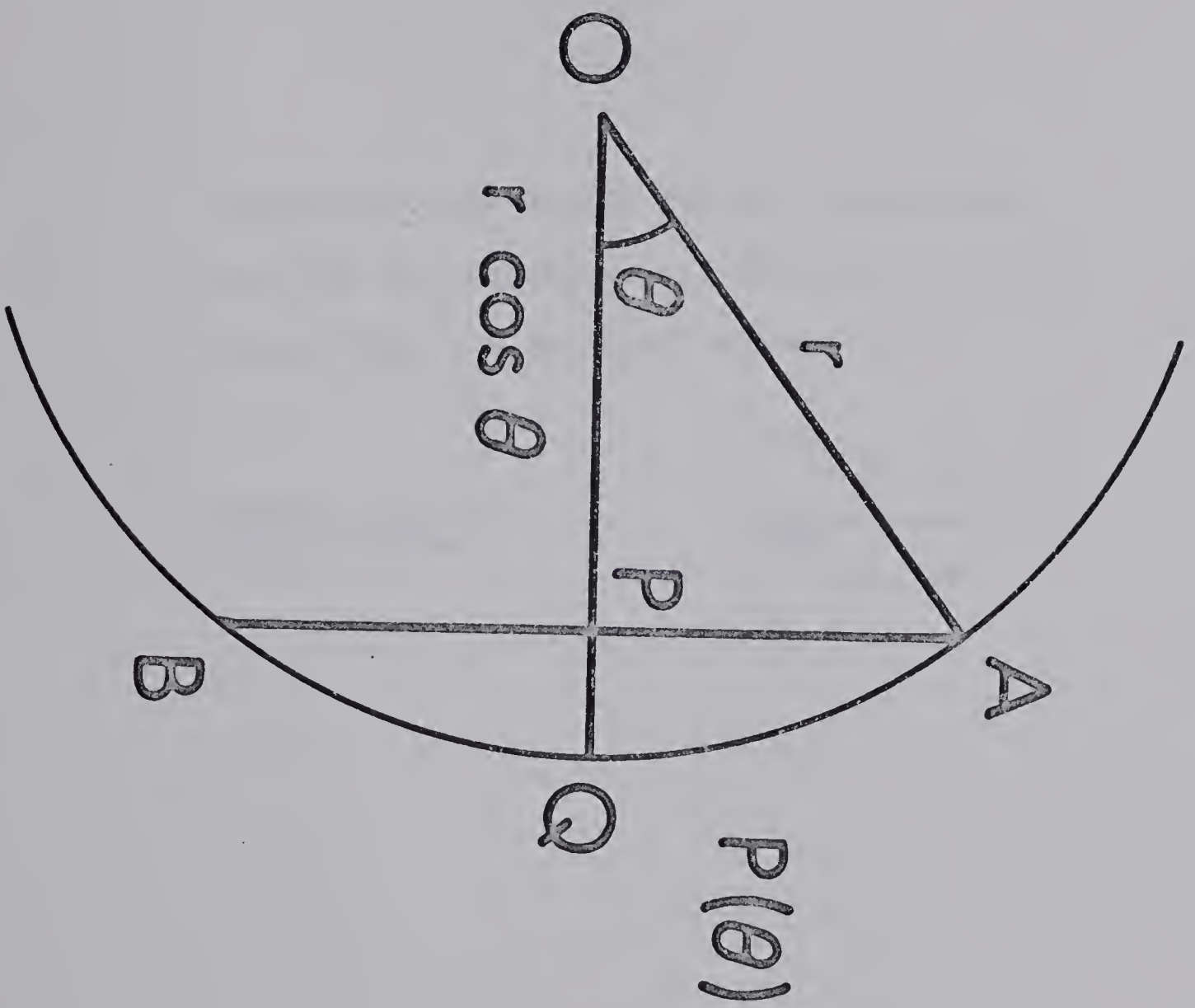
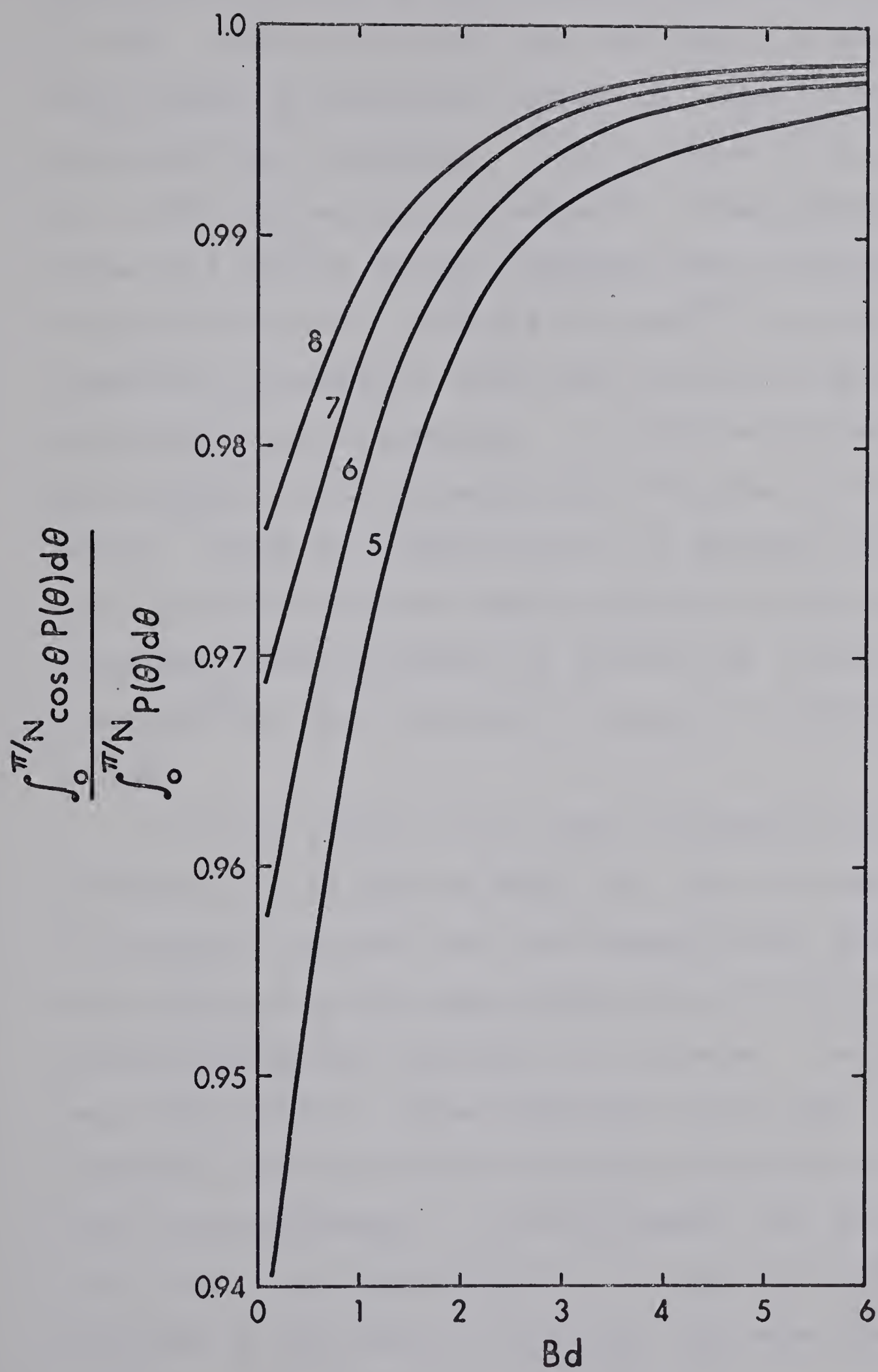


Figure 8

Apparent fractional radius due to the rotation of the ring as a function of the barrier Bd , for various values of n .

$$\text{Apparent fractional radius} = \frac{\int_0^{\pi/N} P(\theta) \cos \theta \, d\theta}{\int_0^{\pi/N} P(\theta) \, d\theta}$$



calculated by means of the hindered rotor method was $1.391\overset{\circ}{\text{\AA}}$. Using this result and the ratio of 0.988, one would expect an uncorrected radius of $1.374\overset{\circ}{\text{\AA}}$ in the individual atom treatment. This is close to the result of $1.377\overset{\circ}{\text{\AA}}$ obtained by Cruickshank¹². The hindered rotor value of $1.391\overset{\circ}{\text{\AA}}$ is in good agreement with Cruickshank's corrected value of $1.392\overset{\circ}{\text{\AA}}$ and the Raman²⁹ value of $1.397\overset{\circ}{\text{\AA}}$. Comparable agreement is found for the case of cyclopentadienyl manganese tricarbonyl. A relative barrier from the hindered rotor refinement of 1.50 gives a ratio of 0.978. Using the hindered rotor C-C distance of $1.424\overset{\circ}{\text{\AA}}$ and a ratio of 0.978 one would predict an uncorrected individual atom refinement C-C distance of $1.394\overset{\circ}{\text{\AA}}$. Berndt and Marsh²⁴ in fact obtained an average C-C distance of $1.394\overset{\circ}{\text{\AA}}$.

A similar pattern of C-C bond lengthening for the other structures can be seen in Table II. In each case, the C-C distance obtained from the hindered rotor refinement was multiplied by the appropriate ratio to obtain a predicted uncorrected individual C-C distance. As can be seen from Table II, these predicted values agree satisfactorily with the results actually obtained from the individual atom refinement. It would appear that the hindered rotor model correctly allows for the apparent bond shortening due to the motion of the ring, and makes the riding

Table II
Comparison of C-C bonds

	Indivigual atom (Å) model	Bd	Ratio from graph	Predicted uncorrected C-C bond Å	Hindered Rotor Result Å	Riding corrected result Å
benzene	1.37	1.95	0.988	1.374	1.391	1.392
$\pi\text{C}_5\text{H}_5\text{Mn}(\text{CO})_3$	1.39	1.50	0.978	1.394	1.424	1.42
$\text{Ru}(\text{C}_5\text{H}_5)_2(1)$	1.446	4.5	0.994	1.442	1.451	-
$\text{Ru}(\text{C}_5\text{H}_5)_2(2)$	1.449	5.2	0.995	1.448	1.455	-
$\pi\text{C}_6\text{H}_5\text{Cr}(\text{CO})_3$	1.401	1.97	0.989	1.393	1.408	-
$\pi\text{C}_7\text{H}_7\text{Mo}(\text{CO})_2\text{C}_6\text{F}_5$	1.407	1.94	0.992	1.400	1.411	-

Table III
Comparison of RMS angular displacement values

Molecule	Barrier Bd	RMS (graph)	RMS (paper)	NMR barrier = $V_O/2kT$
benzene	1.95	8.5°	7.9°	1.8
$(\pi C_6H_6)Cr(CO)_3$	1.97	8.4°	-	-
$(\pi C_7H_7)Mo(CO)_2C_6H_5$	1.94	7.3°	-	-
$\pi C_5H_5Mn(CO)_3$	1.50	12.0°	12°	-
Ruthenocene (1)	4.5	5.8°	-	2.01
Ruthenocene (2)	5.2	5.4°	-	

correction unnecessary, assuming only one dominant libration.

Barrier to Rotation

The potential function for the hindered rotor was

$$V = -\frac{1}{2} V_0 \cos n(\theta' - \xi) \quad [16]$$

The barrier to rotation in kilocalories per mole is V_0 . The refined relative barrier, Bd , is related to the barrier of rotation V_0 by

$$Bd = \frac{V_0 \times 1000}{2 \times 1.987 \times T} \quad [19]$$

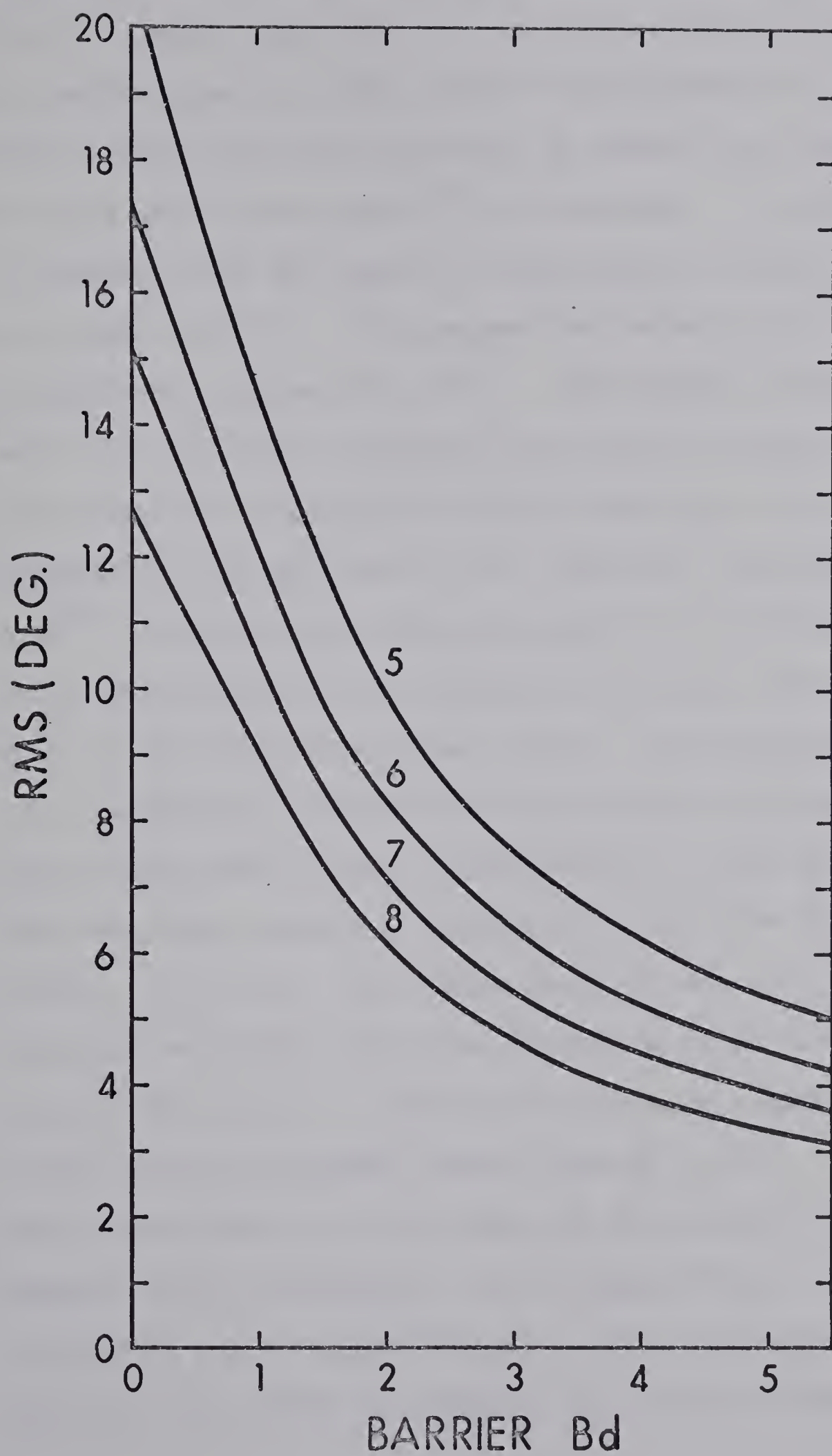
The value of Bd is an indication of the amount of rotation the group undergoes. A barrier of infinity would give no rotation while a barrier of zero would result in free rotation of the rigid group. Intermediate values are directly related to the root mean square angular displacement of the group.

$$RMS = \sqrt{\frac{\int_0^{\pi/N} \theta'^2 P(\theta') d\theta'}{\int_0^{\pi/N} P(\theta') d\theta'}} \quad [53]$$

Since $P(\theta')$ is a known function of Bd (see equation [17]), angular displacement versus relative barrier can be made for various values of n - the order of the ring. Figure 9

Figure 9

Root mean square angular displacement versus relative
barrier for various values of n



is such a plot for five up to eight membered rings. Thus for $\text{Mn(CO)}_3\text{Cp}$, which has a barrier (Bd) of 1.49, one would expect an RMS angular displacement of 12° . This is also the value obtained by Berndt and Marsh²⁴ by means of Cruickshank's¹³ ω -T analysis. In the case of benzene, the RMS angular displacement obtained from the graph is 8.5° . This agrees satisfactorily with Cruickshank's value of 7.9° ¹². RMS angular displacements for the other molecules are given in Table III. The values of the relative barrier obtained can also be compared to the NMR results for benzene²⁷ and ruthenocene²⁸. Andrews and Eades obtained an activation energy for reorientation of the benzene ring about the hexad axis of 3.7 ± 0.2 kcal/mole below $\sim 240^\circ\text{K}$. Unfortunately, above this temperature the barrier height drops rapidly and one can at best make a rough approximation to the barrier from the slope assuming rotation is still the dominant motion. At -3°C , the temperature at which the crystal structure was done, this was estimated to be 1.8 kcal/mole giving a Bd of 1.7. This is in the same range of values as that of the hindered rotor value of 1.9(1). The agreement of the NMR value with that of the hindered rotor appears to be fortuitous. For $\pi\text{C}_6\text{H}_6\text{Cr(CO)}_3$, $\pi\text{C}_5\text{H}_5\text{Mn(CO)}_3$ and $\pi\text{C}_7\text{H}_7\text{Mo(CO)}_2\text{C}_6\text{H}_5$, NMR results were not available with which to compare the hindered rotor results.

The barriers (B_d) obtained for these compounds were of the same order of magnitude as that of benzene which would be expected for oscillations of the same order as benzene. The results obtained are shown in Table III. The value obtained for ruthenocene was greater than those for the other compounds and had a much larger error associated with it: 5.0 ± 2.6 . This did not agree very well with the NMR result obtained by Holm and Ibers²⁸ of $V_O = 2.3 \pm 0.2$ kcal/moles which is equivalent to $B_d = 2.01$. The reason for this discrepancy may be in the nature of the potential function used. In a cosine function the activation energy required for rotation of the ring should equal V_O . If, for example, the potential function were of the form shown in Figure 10 which has a cosine function in the immediate area of the potential minimum and a fixed potential elsewhere, the activation energy would be V_O' while the " B_d " described by a purely cosine function would be $B_d = V_O/2kT$ and not $V_O'/2kT$. A plot of $P(\theta')$ versus θ' (Figure 11) for various values of B_d shows that for $B_d = 4.0$ or greater the atoms are for the most part confined to within a few degrees of the minimum.

In this case, B_d would be insensitive to the shape and height of the potential function outside this region. B_d is then a parameter describing the shape of the function close to the minimum rather than a genuine measure of an

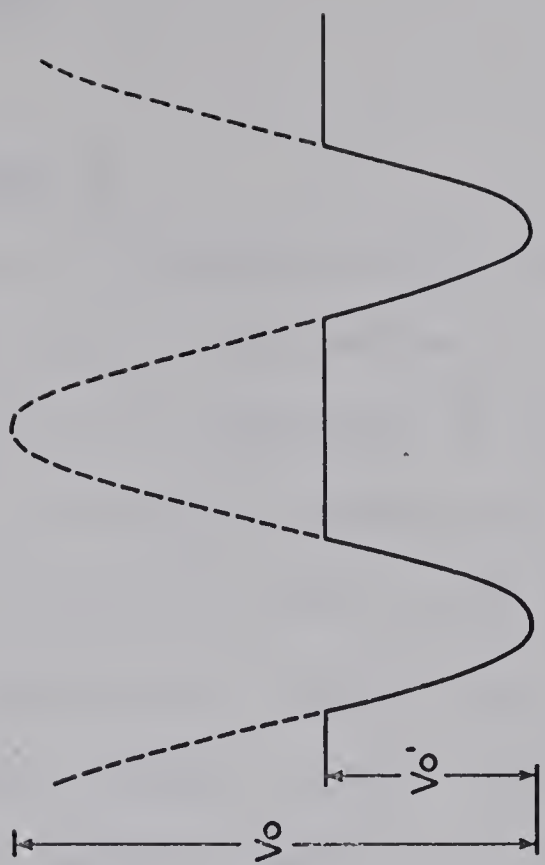
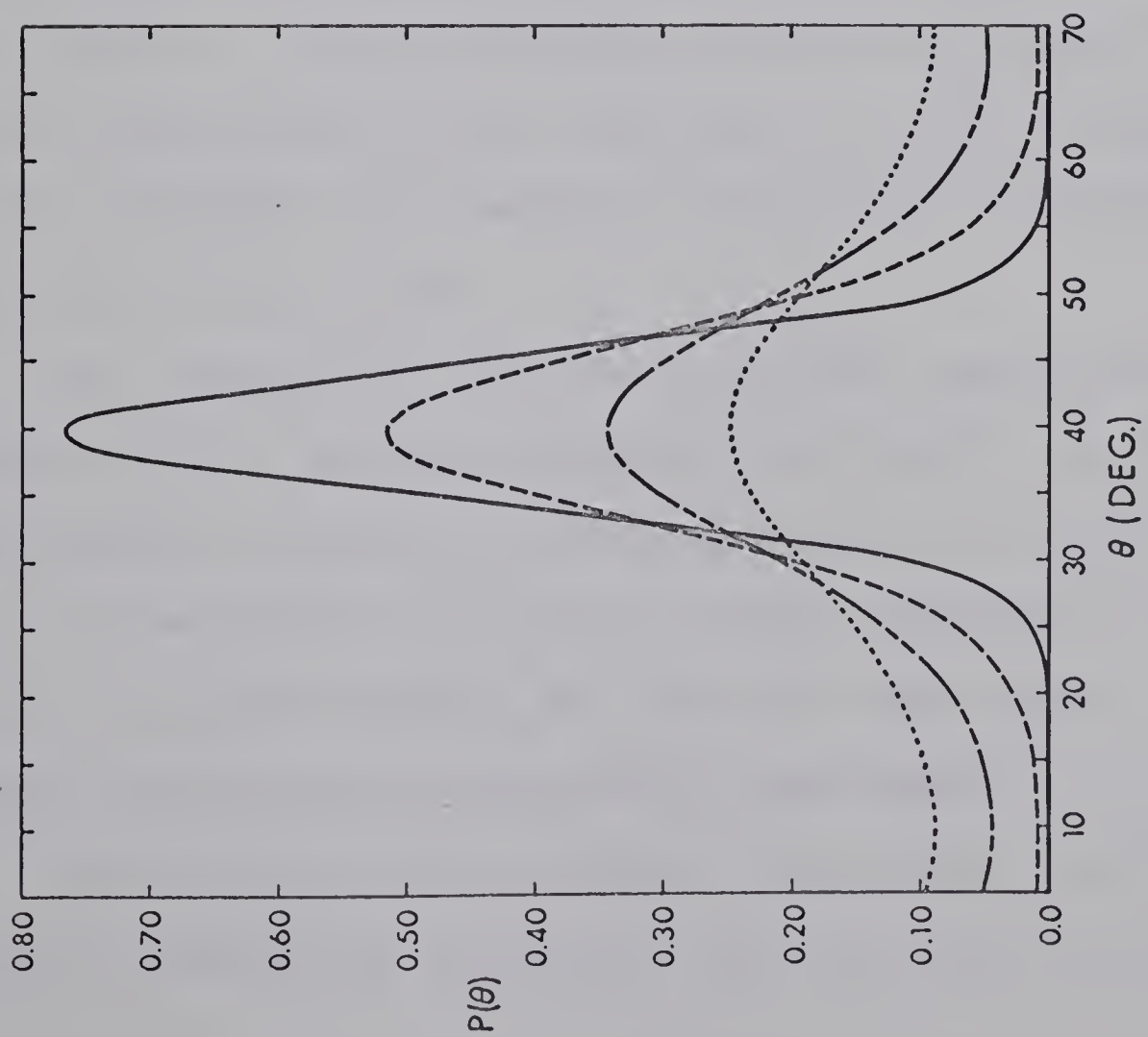
Figure 10

A possible form of the potential function

Figure 11

$P(\theta')$ versus θ' for a cosine distribution function

————	$Bd = 4.0$
-----	$Bd = 2.0$
-----	$Bd = 1.0$
.....	$Bd = 0.5$



activation energy.

Temperature Factor B

In the case of hindered rotors, the assumption is being made that most of the motion of the ring is in the plane of the ring and tangential to it. In the case where the ring is strongly bonded to a metal atom, one might expect the ring to ride on the metal. If this were the case, the temperature factors for the rings as a whole, B , should be very similar to the isotropic B of the metal atom, B_M . This indeed seems to be the case as can be seen in Table IV. In each case, the temperature factor seems to be slightly greater than that of the metal to which it is attached. This difference could result from out of plane oscillations by the ring which are not allowed for in the hindered rotor motion, and also by librations of the molecule as a unit.

The observation that the isotropic temperature factor of the ring is slightly greater than that of the metal can be utilized to obtain a starting value of the barrier. This is desirable as it is not always possible to calculate a relative barrier, B_d , from the NMR result, nor would the results be necessarily meaningful.

Application of the hindered rotor model would generally be preceded by individual atom isotropic refinement

or by rigid body refinement. Both of these methods would give B_M and an average temperature factor for atoms in the ring, B_{iso} . These results can be utilized to obtain starting values for the hindered rotor refinement as follows:

If one takes the case where the major motion of an atom is tangential to the ring and this direction is parallel to the z axis, the x and z axes in an orthogonal co-ordinate system being in the plane of the ring, then the elements of the less common tensor description for thermal motion become

$$\begin{aligned} B_{11} &= (B_M + 0.6) & B_{12} &= 0 \\ B_{22} &= (B_M + 0.6) & B_{13} &= 0 \\ B_{33} &= (B_M + 0.6) + 8\pi^2 \bar{u}_\omega^2 & B_{23} &= 0 \end{aligned} \quad [54]$$

The equivalent isotropic B_{iso} of the atom is

$$B_{iso} = B_M + 0.6 + \frac{8\pi^2}{3} \bar{u}_\omega^2 \quad [55]$$

The root mean square displacement (in Å) is

$$\bar{u}_\omega = R \sqrt{\frac{\int_0^{\pi/N} \theta'^2 P(\theta') d\theta'}{\int_0^{\pi/N} P(\theta') d\theta'}} \quad [56]$$

Substituting for \bar{u}_ω in equation [55], one obtains:

Table IV
Comparison of temperature factors for ring
and metal atom

	B (Å) ²	B _M (Å) ²	ΔB (Å) ²
$\pi\text{-C}_5\text{H}_5\text{Mn(CO)}_3$	3.64 (17)	3.18	0.46
$\pi\text{-C}_6\text{H}_6\text{Cr(CO)}_3$	3.09 (9)	2.63	0.46
ruthenocene (1)	2.11 (30)	1.50	0.61
ruthenocene (2)	2.48 (30)		0.98
$\pi\text{-C}_7\text{H}_7\text{Mo(CO)}_2\text{-C}_6\text{F}_5$	3.13 (13)	2.60	0.53
			Average 0.60

$$B_{iso} = B_M + 0.6 + \frac{8\pi^2}{3} R^2 \frac{\int_0^{\pi/N} \theta'^2 P(\theta') d\theta'}{\int_0^{\pi/N} P(\theta') d\theta'} \quad [57]$$

Since $P(\theta')$ is known as a function of B_d ,

$$\frac{8\pi^2}{3} \frac{\int_0^{\pi/N} \theta'^2 P(\theta') d\theta'}{\int_0^{\pi/N} P(\theta') d\theta'} \quad \text{can be plotted as a function}$$

of the barrier B_d (Figure 12) for various values of n . In Figure 13 a C-C bond distance of 1.392\AA is assumed and

$$\begin{aligned} \Delta B &= B_{iso} - (B_M + 0.6) \\ &= \frac{8\pi^2}{3} R^2 \frac{\int_0^{\pi/N} P(\theta') \theta'^2 d\theta'}{\int_0^{\pi/N} P(\theta') d\theta'} \quad [58] \end{aligned}$$

is plotted against the barrier B_d . Thus, knowing the average B_{iso} of the atoms in the ring and the isotropic temperature factor of the metal, ΔB , can be calculated and a starting value of the barrier obtained from Figure 12 or 13. For example, in $\pi\text{-C}_6\text{H}_6\text{Cr(CO)}_3^{25}$ the individual atom refinement gives B_{iso} of 4.5 for carbon atoms in the

Figure 12

$\Delta B/R^2$ as a function of barrier, Bd

$$\text{where } \Delta B = \frac{8\pi^2}{3} R^2 \frac{\int_0^{\pi/N} \theta'^2 P(\theta') d\theta'}{\int_0^{\pi/N} P(\theta') d\theta'}$$

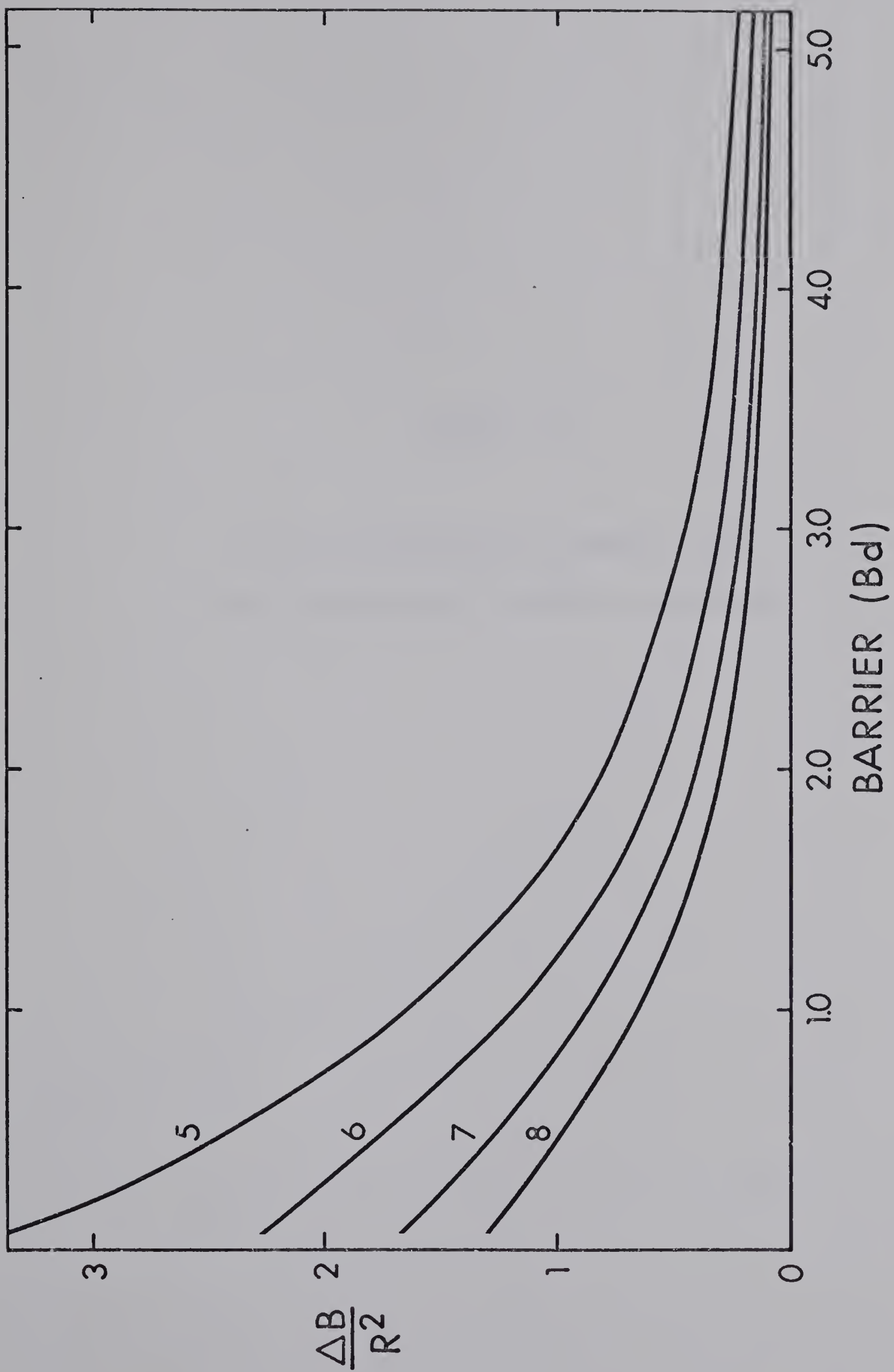
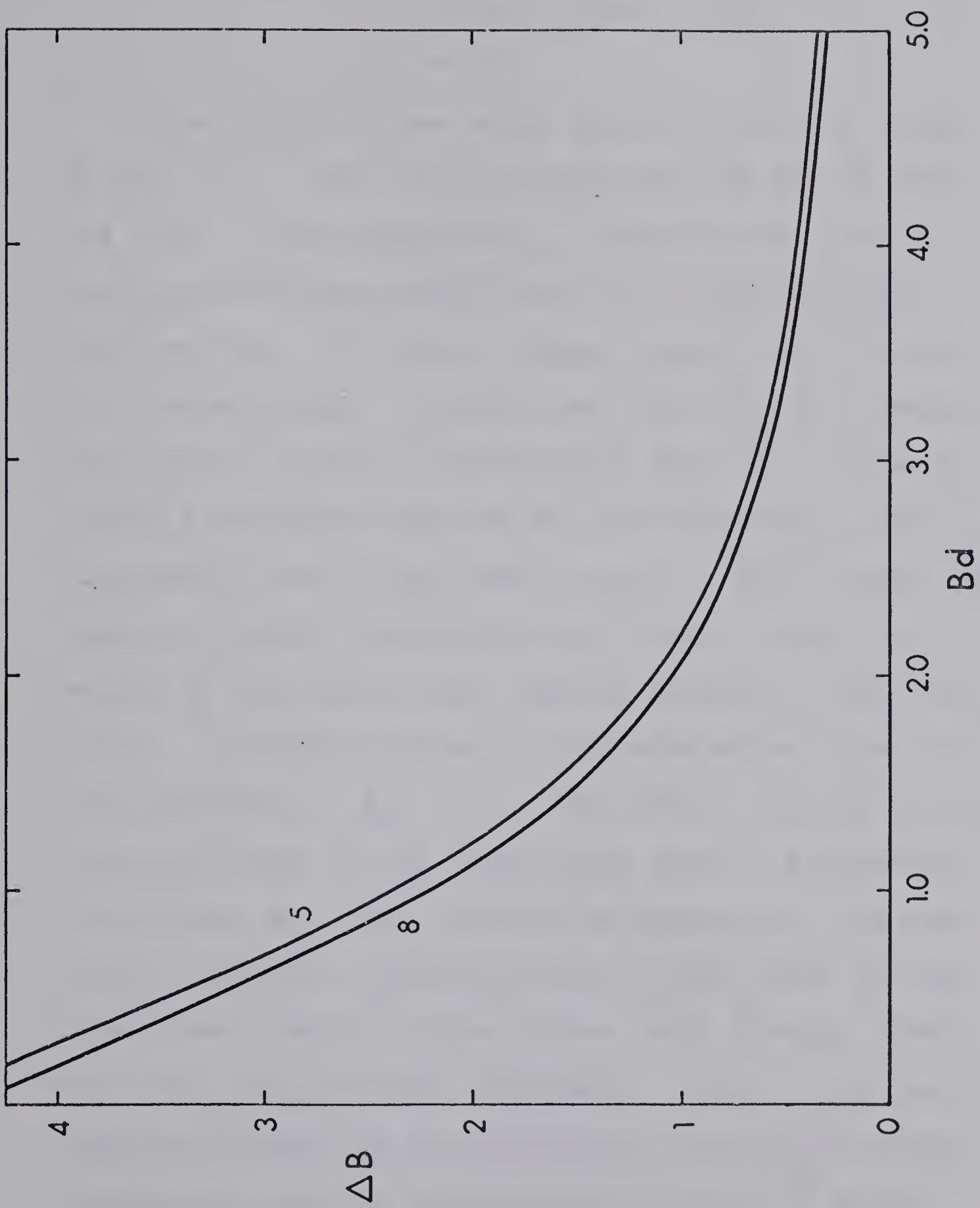


Figure 13

ΔB as a function of barrier, B_d
(C-C distance of $1.392\overset{\circ}{\text{\AA}}$ is assumed)



benzene ring and $B_{iso} = 2.6$

$$\begin{aligned}\therefore \Delta B &\approx B_{iso} - (B_M + 0.6) \\ &= 4.5 - (2.6 + 0.6) \\ &= 1.3\end{aligned}$$

From Figure 13 one would obtain a starting value of $B_d = 1.7$. The final refined value of the barrier was 1.97. For $\pi\text{-C}_5\text{H}_5\text{Mn(CO)}_3$ a starting value of 1.5 would be calculated while that for $\pi\text{-C}_7\text{H}_7\text{Mo(CO)}_2\text{C}_6\text{F}_5$ would be 1.8. The final refined values were 1.50 and 1.94 respectively. Although not too accurate a method, this has the practical application that one is able to obtain a starting value for the relative barrier that is acceptably close to the final result. This reduces the number of cycles required for refinement, since the method of non-linear least squares requires a good starting value. A starting value of the temperature factor of the ring would be $B = B_M + 0.6$. The use of reasonably correct starting values of the temperature factor is important since there is a high correlation between the two parameters, B_d and B , typically being of the order of 0.60. Thus a poor starting value of one could greatly affect the other and slow down refinement. Such a high correlation between the two parameters could explain the oscillation problems experienced by Prewitt¹⁸ et al. as they refined using block diagonal least squares, the

barrier and ring temperature factor being in separate matrices.

Test on $-\text{CF}_3$ Groups

The hindered rotor model was also tested out on the fluorines of the four $-\text{CF}_3$ groups in $\text{Cs}[\text{Y}(\text{HFA})_4]^{26}$, where HFA represents the hexafluoroacetylacetonate ion, $\text{CF}_3\text{COCHOCF}_3$. $-\text{CF}_3$ groups as well as such groups as perchlorates are noted for undergoing large thermal motions. The assumption was made that the fluorine groups had D_{3h} symmetry and that the dominant motion was in the plane of the fluorine atoms. The results of the hindered rotor calculations are given in Table V. Although the R factor is not as good as that of the individual atom model, the hindered rotor model does lengthen the C-F and C- CF_3 bonds and brings them close to the corrected values obtained by Cruickshank, Jones and Walker in ammonium trifluoroacetate³⁰. Since $-\text{CF}_3$ groups are less rigid, the assumption of D_{3h} symmetry of the fluorine groups is probably not very good and is the probable cause of the poor structure factor amplitude agreement of the hindered rotor model. This view gains some support from the deformations observed in the individual atom refinement.

Table V

Results of hindered rotor refinement of Cs[Y(HFA)₄]

(a) Hindered rotor parameters for CF₃ groups

Group	x	y	z	B	Bd	R	D	E	F
1	-.4526(7)	.4162(6)	.3871(5)	5.6(2)	0.65(4)	1.249(7)	0.175(8)	1.832(6)	2.511(9)
2	.1730(8)	.5064(5)	.3983(4)	7.5(3)	1.67(5)	1.234(7)	5.534(9)	1.121(9)	4.279(8)
3	-.2499(11)	.1165(6)	.3442(6)	8.3(3)	0.44(4)	1.228(8)	5.453(13)	1.509(16)	3.545(17)
4	.3133(9)	.2188(6)	.4640(5)	7.7(3)	0.84(5)	1.233(8)	3.232(8)	0.870(8)	3.755(17)

(b) Bond lengths (Å)

	Cs[Y(HFA) ₄]		Other molecules	
	Hindered rotor	Individual atom (uncorrected)	Uncorrected	Corrected
Average C-F	1.319	1.26(1)	1.334	1.346(5) ³⁰
Average C-CF ₃	1.576	1.501(16)	-	1.54(2) ³¹
			-	{ 1.55-1.60 ³²

Reliability Factor

Another test of the validity of the hindered rotor model is on purely statistical grounds. An increased number of parameters being refined will normally result in a lower R factor. Going from the hindered rotor model to the individual atom model requires more parameters and should result in a drop in the R factor. Whether such a drop is significant can be decided by comparing the ratio of R factors with those given by Hamilton³³. If the ratio is greater than that given by Hamilton then the increase in the number of parameters is statistically justified at the significance level chosen (in this case 90%). Thus the hindered rotor model, which has fewer parameters, would be statistically justified only if $\frac{R_{HR}}{R_{IA}} < \text{Hamilton ratio}$. This appears to be true for only $\pi\text{-C}_5\text{H}_5\text{Mn(CO)}_3$ and $\text{Ru(C}_5\text{H}_5)_2$.

Time and Storage Requirements

One practical advantage of the hindered rotor model over that of the isotropic or anisotropic atom model is that it requires fewer parameters. Since the matrix size = $n(n+1)/2$ where n is the number of parameters refined, the storage requirements increase approximately in proportion to n^2 . Furthermore, the time required to build the matrix and later to invert it is also related

Table VI
R factors (%)

	Published R Factor	Individual Atom R _I # Parameters	Hindered Rotor R _I # Parameters	Hamilton Ratio	Ratio of R's
π -C ₅ H ₅ Mn(CO) ₃	9.0	8.9	8.9	1.033	1.00
benzene	9.9	8.6	10.1	1.078	1.17
π -C ₆ H ₆ Cr(CO) ₃	4.1	4.1	4.6	1.039	1.12
π -C ₇ H ₇ Mo(CO)-C ₆ F ₅	5.9	5.5	5.7	1.015	1.03
Ru(C ₅ H ₅) ₂	9.8	8.6	8.7	1.010	1.01

to the matrix size and also increases as n^2 . This effect is especially noticeable in large structures such as $(C_5H_5CoCO)_2(GeCl_2)_2Fe(CO)_4^{34}$. In this structure refinement of both molecules in the asymmetric unit, with the heavy atoms anisotropic and all others isotropic, requires 339 parameters, 350K bytes of storage and two hours of computing time per cycle. Use of the hindered rotor for the four cyclopentadienyl rings drops the number of parameters to 295 and the storage to 300K bytes. The time per cycle is cut in half to 63 minutes. Thus, although the addition of subroutines to handle the hindered rotor calculations lengthens the programme (~12K bytes), it is far outweighed by the core saved due to the large drop in the number of parameters to be refined, and concomitant reduction in storage assigned to the matrix.

For crystal structures such as benzene in which only a few parameters are refined, a large percentage of the computing time is spent in calculating structure factors. Because calculation of the M function is slower than the corresponding individual atom expressions, the cycle time will be slower for the hindered rotor model than for the individual atom model. However, for larger structures such as $\pi-C_7H_7Mo(CO)_2C_6F_5$, in the order of 75% of computing time² is spent in building and inverting the matrix. In

this case, the slow calculation of the M function is outweighed by the reduction in matrix building time and the hindered rotor model is faster. The break even point appears to be for structures of the size of $\pi\text{-C}_5\text{H}_5\text{Mn(CO)}_3$ for which approximately the same time was required in both methods.

Conclusions

The development of the hindered rotor model within the framework of a conventional structure factor and least squares programme was undertaken for specific application to $\pi\text{-C}_n\text{H}_n$ groups bonded to metal atoms. The model is attractive in that it emphasizes the physically real phenomenon of hindered rotation of these groups in the solid state. The fixed form of the potential function does not appear to be a serious problem as long as no physical significance is attached to the height of the barrier to rotation that is obtained. Bond lengths derived from the hindered rotor model agree well with those obtained from independent atom models after corrections for thermal motion have been made. The model has been found to be competitive with independent atom models as judged by a statistical comparison of conventional crystallographic R factors.

Chapter II

THE CRYSTAL STRUCTURE OF FERROCINIUM PICRATE

INTRODUCTION

Once the hindered rotor programme was written and tested it was decided to test the programme for use in a routine crystal structure. The compounds ferrocene and ferrocinium picrate both contain cyclopentadienyl rings. The crystal structure of ferrocene by Dunitz, Orgel and Rich³⁵ shows a libration of the cyclopentadienyl rings about the five fold axis, making it ideal for the hindered rotor treatment. Such libration could also occur in the ferrocinium ion of such salts as ferrocinium picrate. Since use of the hindered rotor model gives good average bond lengths, the structural results could be used as evidence in determining the bonding in such structures. For these reasons, the data for ferrocene were re-collected by Kathleen Simpson³⁶ and the structure of ferrocinium picrate undertaken.

EXPERIMENTAL

The compound ferrocinium picrate $[\text{Fe}(\text{C}_5\text{H}_5)_2^+]$ $[\text{H}_2\text{C}_6\text{N}_3\text{O}_7^-]$ was prepared by Ester Lorah³⁷ from ferrocene dissolved in a benzene/ethanol solution. This was oxidized with quinone in the presence of a stoichiometric amount of picric acid. Her analysis showed

	C	H	N
calculated	46.36%	2.92%	10.15%
found	46.00%	3.06%	10.40%

A thermogram run to $150^\circ - 160^\circ$ gave no weight loss, indicating no solvent and no excess ferrocene present in the crystals. The crystals were long black needles extinguishing along the needle axis. The symmetry of rotation, Weissenberg ($hk0 - hk3$) and precession photographs ($0k\ell$) showed the compound to be orthorhombic. Systematic absences of $hk\ell$: $h+k = 2n+1$ and $h0\ell$: $\ell = 2n+1$ allowed three possible space groups - $\text{Cmc}2_1$, Cmcm or $\text{C}2\text{cm}$, a non-standard setting of $\text{Ama}2$. Of these, only Cmcm is centrosymmetric. The unit cell $a = 12.513(7)\text{\AA}$, $b = 20.267(9)\text{\AA}$ and $c = 6.903(7)\text{\AA}$ ($V = 1750.6\text{\AA}^3$) was obtained from least squares refinement of 20 values of carefully centred, high angle reflections measured on the Picker four circle manual diffractometer. The

density was determined by flotation, using a mixture of benzene and dibromomethane, to be 1.60 gm/cc. This is consistent with a calculated density of 1.57 gm/cc for 4 molecules per unit cell. Since both space groups $Cmc2_1$ and $Ama2$ have eight general positions and $Cmcm$ has sixteen, this would require the ionic species to sit on special positions such as a mirror or two fold axis. Each space group has one or more special positions. $Cmcm$ has site symmetries of mm and $2/m$. Site symmetry of mm is possible for both the ferrocinium and picrate ions. It would require the ferrocinium ion to be in the eclipsed formation while the picrate would be required to be planar. Site symmetry $2/m$ is possible for the ferrocinium ion requiring a staggered arrangement of the cyclopentadienyl rings, but is impossible for the picrate ion unless highly disordered. The space groups $Cmc2_1$ and $C2cm$ both have site symmetry m . Both the ferrocinium and picrate ions can possess this symmetry as well as the two-fold symmetry found in $C2cm$. The possibilities of site symmetries in the three space groups, assuming no major disorder, are summarized in Table VII.

Data were collected using a crystal (0.23 mm x 0.057 mm x 0.062 mm) mounted along the needle axis (c)

Table VII
Possible site symmetry

	Space Group	Site Symmetry of FeCp_2^+	Site Symmetry of Picrate Ion
1	Cmcm	mm	mm
2	Cmcm	2/m	mm
3	Cmc2 ₁	m	m
4	C2cm	m	m
5	C2cm	m	2
6	C2cm	2	m
7	C2cm	2	2

and with the c^* axis co-incident with the diffractometer ϕ axis. A moving crystal, moving counter technique ($\theta/2\theta$) was used with a scan width of 2° at two degrees per minute. CuK_α X-radiation was used, monochromated by an oriented graphite crystal (002 reflection). The diffracted beam aperture used was 1.0 mm and the pulse height analyser was centred on the CuK_α peak to accept 90% of the radiation. Stationary background counts of 30 seconds were taken on each side of the scan. Six hundred and forty-two unique reflections (hkl ; h , k and l positive only) were collected, of which 276 were rejected using the criteria (1) $I \leq 0$ and (2) $I \leq 3\sigma$ where $I = [P - t(B_1 + B_2)]$ and $\sigma = [P + t^2(B_1 + B_2)]^{1/2}$

where P peak count

$\left. \begin{matrix} B_1 \\ B_2 \end{matrix} \right\}$ background counts

and t ratio of total peak count to total background count

Seven well distributed standard reflections measured periodically throughout the data collection showed no decomposition of the crystal. No absorption correction was made ($\mu(\text{Cu}) = 75.1 \text{ cm}^{-1}$).

SOLUTION AND REFINEMENT

A three dimensional Patterson map gave the iron position as $x = 0.0$ and $y = 0.185$, the z co-ordinate being undefined. Although simple interpretation of the Patterson does not determine the space group, site symmetry considerations given in Table VII reduced the possible arrangements from seven to two. The Patterson solution eliminates possibilities 2, 4, 5, 6 and 7, leaving only 1 and 3 for serious consideration and reduces the space group choice to either $Cmcm$ or $Cmc2_1$. The space group $Cmc2_1$ was chosen for initial refinement because the crystal showed $mm2$ symmetry. Since this space group is polar, the z co-ordinate was fixed at $z = 0.0$. A difference map ($F_o - F_c$, F_c phased on iron only) revealed the location of all other non-hydrogen atoms. With the space group chosen as $Cmc2_1$, the iron only phased Fourier map is centrosymmetric and has the symmetry of $Cmcm$. If the non-centrosymmetric space group is correct, the effect of the extra symmetry is to show the true electron density distribution, together with its superimposed mirror image. The picrate atoms, with the exception of the ortho nitro group, were found on the mirror plane ($z = 0.0$). The appearance of diffuse ortho nitro group peaks on either side of the plane could result from

either non-centric space group with its superimposed mirror image, or centrosymmetric space group with disorder of the nitro group. The cyclopentadienyl rings of the ferrocenium cation were revealed as smears of electron density and possessing mm symmetry. This could also be due to superposition of the mirror image in the non-centrosymmetry space group, or if centrosymmetric, due to a large oscillation of the rings. Calculation of the structure factors using all non-hydrogen atoms gave an R factor of $R_1 = 26.7\%$, $R_2 = 28.3\%$, where

$$R_1 = \frac{\sum ||F_O| - |F_C||}{\sum |F_O|}$$

$$\text{and } R_2 = \left[\frac{\sum w (|F_O| - |F_C|)^2}{\sum w |F_O|^2} \right]^{\frac{1}{2}}$$

The weights used were obtained from peak and background counts (weight $w = 1/\sigma^2(F)$, $\sigma(F) = (P + t^2(B_1 + B_2) + p^2I)^{\frac{1}{2}}$, $p = 0.03$).

Four cycles of refinement, minimizing the function $\sum w (|F_O| - |F_C|)^2$, with all atoms isotropic and the cyclopentadienyl ring as a hindered rotor brought the R factor down to $R_1 = 11.6\%$.

The final orientation of the cyclopentadienyl ring

was 18° from the orientation required by the additional mirror of Cmc_m. There remained the distinct possibility of refinement into a false minimum corresponding to the pseudo-mirror at $z = 0.0$. Also, in a non-polar space group such as Cmc2₁, anomalous dispersion³⁸ (in this case for the iron atom) can produce significant co-ordinate changes. Four models were then refined with the inclusion of anomalous scattering for iron:

- (1) Previous model; refinement terminated with Bd negative.
- (2) Mirror image of previous model; $R = 11.9\%$.
- (3) Ferrocinium ion as in (1), picrate as in (2);
 $R = 11.8\%$.
- (4) Mirror image of (3); refinement terminated with Bd negative.

Refinement of models (1) and (4) resulted in negative values for Bd indicating that the mirror image of the cyclopentadienyl ring was preferred. This suggests that the omission of anomalous scattering favoured the choice of a false minimum for the cyclopentadienyl ring in the initial model.

For the ferrocinium ion, it is reasonable to expect the centres of gravity of the rings and the iron atom to be approximately collinear. In this crystal structure, this would require the z co-ordinate of the ring centre to be zero. However, models (2) and (3) showed a diff-

erence in z corresponding to $\sim 0.03\text{\AA}$.

Cruickshank and MacDonald's³⁸ formula

$$\Delta z \text{ (in } \text{\AA}) = \frac{2}{\pi S_{\max}} \left[\frac{\Delta f''}{|f|} \right]_{\frac{1}{2} S_{\max}}$$

where $S = \frac{2 \sin \theta}{\lambda}$

predicts a difference in the z co-ordinate of 0.03\AA in the incorrect model which agrees with the observed shifts. This suggests that models (2) and (3) are also wrong and that none of the cases considered so far are correct.

The structure was then refined using the centrosymmetric space group $Cmcm$. This group requires a mirror plane through the plane of the picrate ion and the iron atom. Isotropic refinement of all atoms brought the R factor to $R_1 = 10.7\%$, $R_2 = 13.7\%$, with disorder allowed for the ortho nitro group. These atoms (N1, O2 and O3) were given half weight above and below the mirror plane for the following reasons:

(1) The initial difference map showed considerable smearing of the peaks for these atoms off the mirror plane.

(2) When confined to the mirror plane, the B_{33} values obtained were very large, suggesting that the atoms were in fact off the $z = \frac{1}{4}$ plane, and that the ortho nitro group was not planar with the rest of the

picrate group.

(3) If O1, O2 and O3 were all in the plane, there would be a very short O1-O2 contact of 2.64\AA , and an even shorter O2-O3 contact of 1.90\AA , which is much shorter than the Van der Waal's radius sum. These short contacts would be lengthened if O2 and O3 were off the mirror plane.

(4) Work by Maartmann-Moe³⁹ on the structure of potassium and ammonium picrates showed that the ortho nitro groups are twisted $25^\circ - 27^\circ$ out of the plane of the benzene ring, making a non-planar picrate ion a plausible structure.

(5) Study of the packing diagram (Figure 15) suggests that the nitro group could be accommodated in either the "up" ($z(\text{O3}) > 0.25$) or "down" position. In the "up" position, O3 is within Van der Waal's contact of H14 of the neighbouring ferrocinium ion, while in the "down" it is the same distance from H13. Similarly, O2 is equidistant from H11 and also O1 in either position. This would suggest that the picrate ion crystallizes with a random arrangement of "up" or "down" meta nitro groups. The disordered centric model was chosen rather than a non-centric model as:

(1) no clear distinction could be made among the four possible arrangements of the ferrocinium and

picrate ion;

(2) use of Cruickshank and MacDonald's³⁸ formula for the error in co-ordinates due to the wrong hand when using anomalous dispersion indicated that both hands were incorrect and that the mean - a disordered model - would be a better description; and

(3) the R factor was consistently higher for the non-centric model than for the centric model, despite the fact that the latter has fewer variable parameters.

Allowing the iron atom anisotropic motion and adding the hydrogens of the cyclopentadienyl rings as hindered rotors with C-H distance of $1.0\overset{\circ}{\text{\AA}}$, dropped the R factor to $R_1 = 8.4\%$, $R_2 = 12.3\%$. Further refinement with all atoms anisotropic and the picrate hydrogen added gave a final reliability factor of $R_1 = 6.0\%$, $R_2 = 8.5\%$. It was noticed during anisotropic refinement that some atoms, noticeably O1 as well as the disordered nitro group, had much larger thermal parameters in the z direction. For this reason, O1 was also moved off the mirror plane and given a half population above and below the mirror plane. Refinement of this disordered model gave a final z co-ordinate of $z = 0.1794 \pm .1842$. Because of this large error in the z co-ordinate, O1 was returned to the mirror plane and no other atoms were allowed to become disordered.

In the final cycle of refinement no shift was greater than 1/10 of an esd. A final difference map gave no peak greater than $0.42 \text{ e}^-/\text{\AA}^3$. The observed and calculated structure factor amplitudes (x10) are given in Table VIII. Table IX gives the final positional co-ordinates of the individual atoms and the anisotropic B values are listed in Table X. Hindered rotor parameters for the cyclopentadienyl rings are in Table XI.

Table VIII

Observed and calculated structure factor amplitudes
(x10) in electrons for $[\text{Fe}(\text{C}_5\text{H}_5)_2]^+ [\text{H}_2\text{C}_6\text{N}_3\text{O}_7]^-$

H	K	FO	FC	H	K	FO	FC	H	K	FO	FC	H	K	FO	FC	H	K	FO	FC
L = 0				10	0	597	641	8	6	206	229	9	1	196	187	L = 4			
0	2	252	334	10	2	200	155	8	10	188	192	9	3	257	281	0	0	960	930
0	4	83	78	10	4	132	80	9	1	409	424	9	5	341	363	0	4	123	116
0	6	948	959	10	6	175	180	9	9	264	273	9	11	215	197	0	6	180	135
0	8	721	693	10	8	394	363	9	11	146	140	9	13	146	152	0	8	521	455
0	10	240	248	10	14	249	221	9	15	263	247	10	0	664	702	0	10	142	88
0	12	142	129	11	1	355	349	10	4	383	382	10	6	181	208	0	14	267	299
0	14	654	604	11	3	176	203	10	6	181	197	10	8	319	309	0	16	218	237
0	16	503	451	11	11	235	258	10	10	366	366	11	3	263	270	1	1	428	439
1	1	833	985	12	0	159	210	10	12	177	143	11	5	207	181	1	3	237	196
1	3	1489	1572	12	6	204	192	11	7	268	274	12	0	260	289	1	5	209	178
1	5	1002	1005	12	8	286	251	12	2	205	204	12	2	146	97	1	11	264	250
1	7	80	66	L = 1				12	4	180	147	12	6	209	180	2	0	814	794
1	9	175	179	L = 2				13	1	157	142	L = 3				2	2	138	168
1	11	318	316	0	2	1281	1350	L = 2				0	2	237	221	2	6	413	417
1	13	155	127	0	4	2734	2660	0	0	3450	3525	0	4	730	649	2	8	326	331
1	19	291	258	0	6	498	537	0	2	407	298	0	6	375	335	2	10	142	110
1	21	144	113	0	8	91	89	0	4	303	310	0	8	229	237	2	16	152	170
2	0	852	695	0	10	1167	1120	0	6	290	284	0	10	568	553	3	1	207	210
2	2	1106	1195	0	12	122	117	0	8	604	608	0	12	238	228	3	3	315	326
2	4	543	576	0	14	223	212	0	14	510	516	0	18	191	195	3	5	455	466
2	6	707	737	0	18	226	215	0	16	372	406	1	1	100	120	3	11	206	212
2	8	993	994	0	20	430	434	0	16	372	406	1	1	100	120	3	13	199	194
2	10	130	161	1	1	1254	1341	1	1	1086	1072	1	7	119	107	4	0	509	512
2	12	172	164	1	3	233	211	1	3	454	450	1	11	264	266	4	2	405	401
2	14	177	176	1	5	161	143	1	5	364	358	1	13	148	167	4	4	321	333
2	16	435	437	1	7	130	120	1	11	273	296	1	15	234	256	4	8	253	299
2	20	143	158	1	9	160	114	1	13	174	146	1	17	169	204	5	3	264	287
3	1	217	359	1	11	455	460	1	19	218	216	2	2	736	748	5	5	471	514
3	3	513	467	1	15	356	339	2	0	1190	1204	2	4	600	608	5	11	236	209
3	5	346	348	1	17	299	264	2	2	554	599	2	6	305	315	6	0	234	232
3	7	484	474	2	2	1374	1433	2	4	133	126	2	10	531	526	6	8	157	175
3	9	483	491	2	4	478	475	2	6	761	798	2	12	267	269	7	5	268	254
3	11	623	645	2	6	792	826	2	8	677	664	3	1	548	566	7	11	236	216
3	13	553	589	2	8	200	217	2	10	202	209	3	3	330	334	7	13	203	227
3	15	133	122	2	10	590	609	2	12	141	140	3	5	301	337	8	0	201	199
3	17	224	222	2	12	622	626	2	14	181	155	3	7	518	537	8	8	196	152
4	0	188	186	2	20	218	238	2	16	337	327	3	9	409	436	9	3	222	192
4	2	894	804	3	1	466	440	2	20	126	140	3	11	142	121	9	5	179	169
4	4	443	438	3	3	355	329	3	1	319	300	3	15	309	275	10	0	430	417
4	6	198	197	3	5	315	281	3	3	358	396	3	17	149	145	L = 5			
4	8	462	454	3	7	803	824	3	5	753	775	4	2	300	334	0	2	200	244
4	10	486	488	3	9	593	561	3	7	393	377	4	4	716	764	0	4	339	332
4	12	136	119	3	11	235	224	3	9	321	310	4	6	353	377	0	6	234	244
4	14	229	270	3	13	212	184	3	11	512	502	4	8	126	142	0	10	302	309
5	1	106	137	3	15	397	392	3	13	492	461	4	10	218	228	0	12	188	168
5	3	231	203	3	17	212	265	4	0	825	793	4	12	320	316	1	7	178	167
5	5	1582	1559	3	19	151	156	4	2	396	396	4	16	226	221	1	11	141	139
5	7	131	113	4	2	667	632	4	4	575	555	5	1	728	718	2	2	297	259
5	9	408	404	4	4	777	724	4	6	180	208	5	7	489	518	2	4	281	299
5	11	669	685	4	6	227	250	4	8	607	578	5	9	398	388	2	10	141	207
5	13	267	313	4	10	607	615	4	10	400	370	5	11	138	118	2	12	138	130
5	19	298	301	4	12	376	398	4	14	200	159	6	2	132	130	3	1	236	202
6	0	751	775	4	16	190	201	4	16	154	138	6	4	330	323	3	7	154	173
6	2	829	827	4	20	149	152	5	3	343	343	6	10	169	173	4	2	153	151
6	4	632	612	5	1	826	815	5	5	1213	1189	6	12	203	223	4	4	226	243
6	6	109	94	5	3	370	375	5	9	320	318	7	1	177	182	5	1	216	235
6	8	342	333	5	5	157	122	5	11	513	510	7	3	210	225	5	7	182	191
6	10	114	66	5	7	619	589	5	13	300	280	7	5	207	207	6	4	186	184
6	16	209	199	5	9	745	703	6	0	541	484	7	7	210	222	6	12	168	153
7	1	609	555	5	11	123	108	6	2	379	344	8	2	205	188	7	7	184	179
7	3	1008	993	5	15	817	849	6	4	321	311	8	4	142	137	7	9	193	172
7	5	635	595	6	2	343	345	6	6	223	222	8	6	161	128	L = 6			
7	7	328	328	6	4	628	566	6	16	186	162	9	1	223	242	0	0	392	437
7	9	173	185	6	6	174	183	7	1	324	334	9	9	143	136	0	8	172	227
7	11	239	253	6	12	318	340	7	3	350	307	9	13	132	101	1	3	188	173
7	13	404	398	6	14	285	277	7	5	400	382	10	2	164	180	2	0	262	289
7	17	164	178	6	16	234	249	7	7	148	131	10	4	335	333	2	8	161	159
8	0	1203	1117	7	1	934	871	7	9	235	245	10	6	152	147	4	0	178	97
8	2	572	557	7	3	519	477	7	11	335	337	11	1	176	171				
8	6	335	362	7	5	248	257	7	13	151	154	11	7	206	198				
8	10	148	130	7	7	491	516	8	0	455	460	12	2	192	206				
8	16	242	190	7	9	158	187	8	2	281	282								
9	3	279	288	7	15	369	359	8	4	161	143								
9	5	535	565	7	17	312	279	8	6	128	171								
9	11	235	243	8	2	588	600	8	16	158	174								
9	13	198	178	8	4	428	432												

Table IX

Positional co-ordinates for ferrocinium picrate

Atom	x	y	z
Fe	0.0	0.1851(1)	0.2500
Cl	0.0	0.4408(8)	0.2500
C2	0.0932(9)	0.4829(6)	0.2500
C3	0.9464(9)	0.5500(5)	0.2500
C4	0.0	0.5827(7)	0.2500
N1	0.1996(11)	0.4511(5)	0.2010(29)
N2	0.0	0.6533(7)	0.2500
O1	0.0	0.3801(5)	0.2500
O2	0.2092(8)	0.3960(5)	0.266(10)
O3	0.2733(9)	0.4814(5)	0.1285(21)
O4	0.0862(7)	0.6822(4)	0.2500
H1	0.1680	0.5730	0.2500

Parameters without estimated
errors were not refined.

Table X

Anisotropic temperature (B) factors for ferrocenium picrate

Atom	B ₁₁	B ₂₂	B ₃₃	B ₁₂	B ₁₃	B ₂₃	Equivalent Isotropic B
Fe	3.146	3.709	6.356	0.0	0.0	0.0	4.404
Cl	3.805	3.516	4.807	0.0	0.0	0.0	4.042
C2	3.558	4.001	6.168	1.005	0.0	0.0	4.576
C3	4.350	3.046	7.152	-0.433	0.0	0.0	4.849
C4	3.465	1.972	6.399	0.0	0.0	0.0	3.945
N1	5.768	2.720	4.167	-0.721	0.140	0.201	4.218
N2	5.644	3.662	5.044	0.0	0.0	0.0	4.784
O1	5.621	2.430	12.982	0.0	0.0	0.0	7.011
O2	5.930	4.745	15.585	1.270	-4.983	-3.219	8.744
O3	5.480	5.728	14.339	-0.711	2.073	0.196	8.516
O4	8.351	3.864	9.461	-2.003	0.0	0.0	7.225
H1	-	-	-	-	-	-	5.00

Table XI

(a) Hindered rotor parameters

	x	y	z	B	Bd	R	D	E	ξ
C ring	0.2500	.1354(4)	.1856(3)	5.6(2)	.40(7)	1.186(5)	0.0	$\pi/2$.3141
H ring	0.2500	.1354	.1856	5.6	.40	2.186	0.0	$\pi/2$.3141

(b) Derived co-ordinates for hindered rotor atoms

	x	y	z
Cl1	0.1354	0.2441	0.2500
Cl2	0.1354	0.2037	0.4134
Cl3	0.1354	0.1382	0.3510
H11	0.1354	0.2935	0.2500
H12	0.1354	0.2189	0.5512
H13	0.1354	0.0983	0.4361

RESULTS

A drawing of the ferrocinium picrate molecule is shown in Figure 14. The packing of the ions projected onto the ab plane is given in Figure 15. The atoms in the upper picrate ion are dotted in, while those below are not. Bond angles and bond lengths are given in Table XII, while some interionic contacts are listed in Table XIII.

Figure 14

Drawing of ferrocenium picrate
showing numbering system used

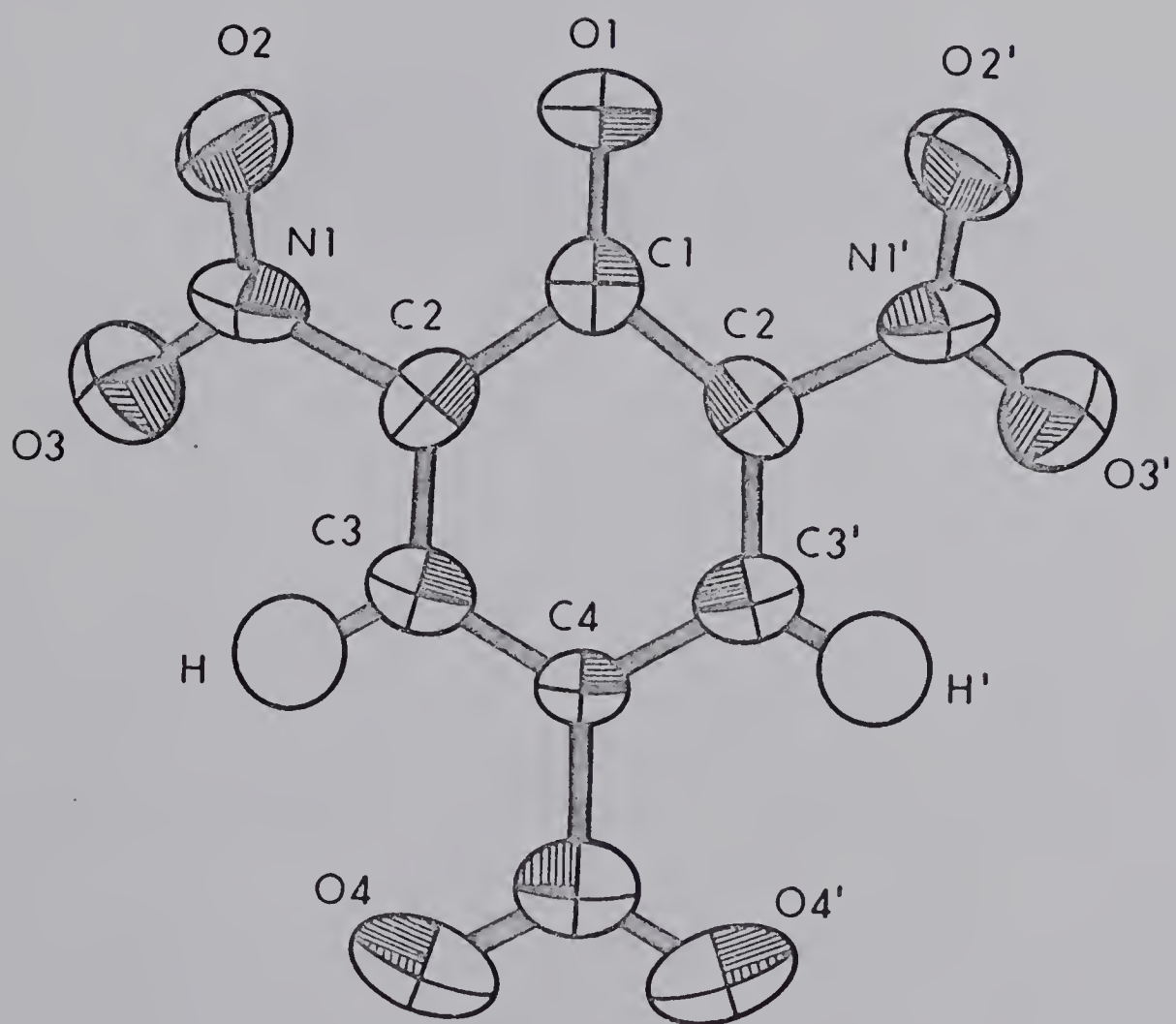
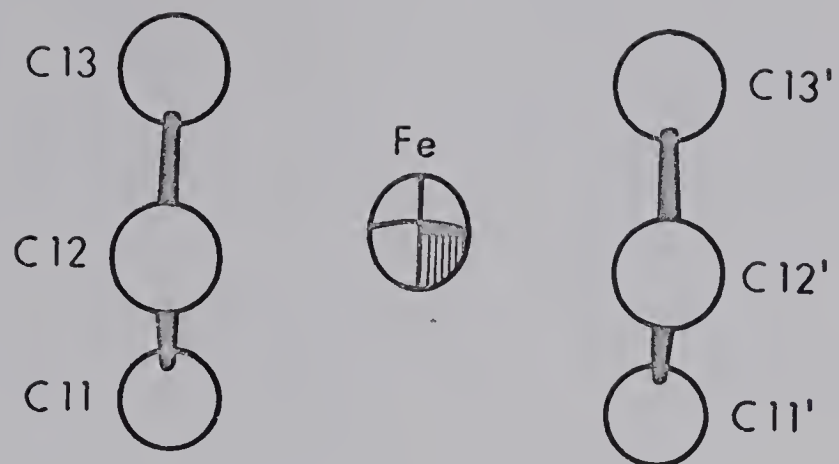


Figure 15

Packing of ferrocinium picrate
as projected onto the ab plane

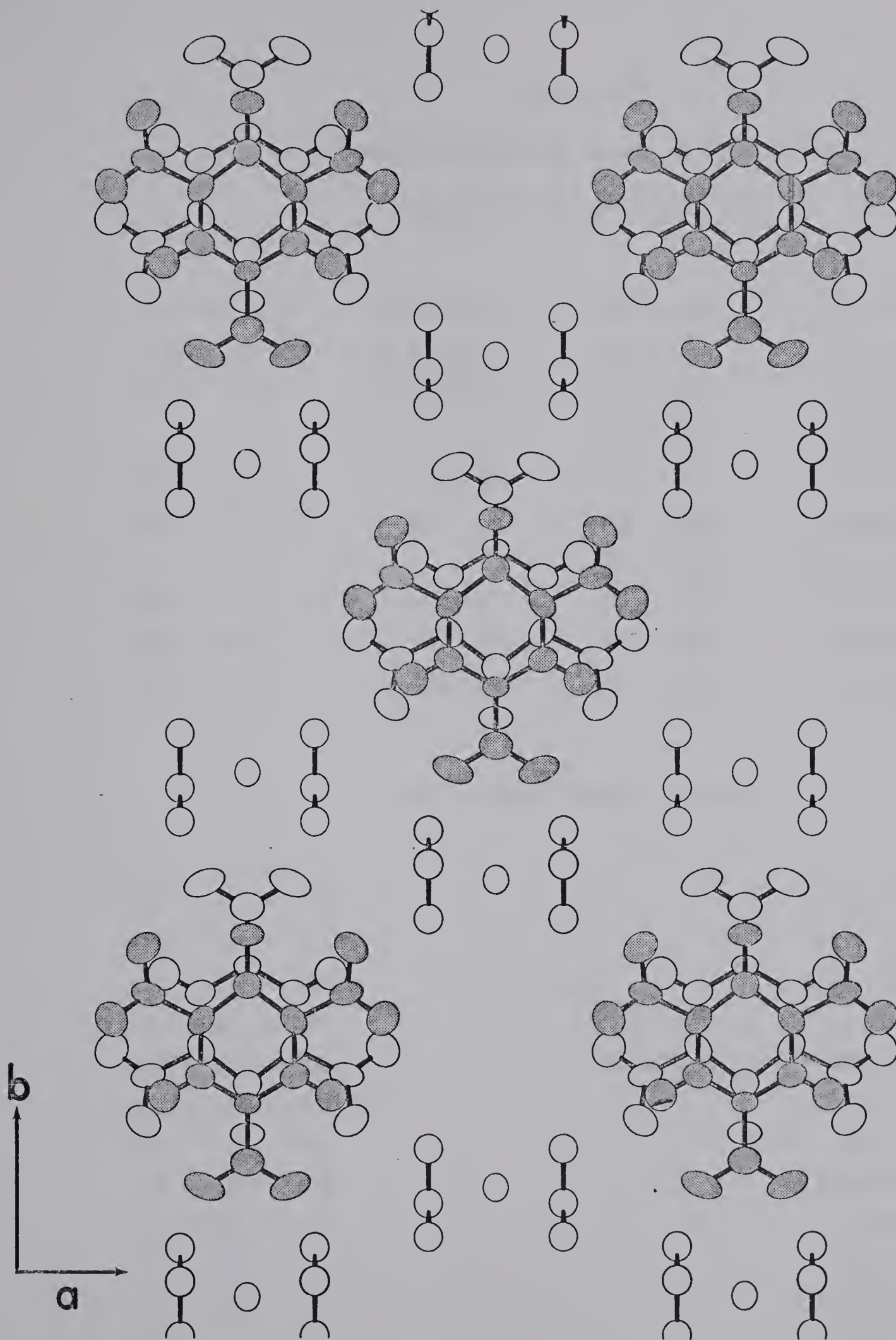


Table XII

(a) Intramolecular bond lengths in
ferrocinium picrate ($\overset{\text{O}}{\text{Å}}$)

Fe - C11	2.074(7)	C1 - O1	1.231(17)
Fe - C12	2.070(5)	C1 - C2	1.445(14)
Fe - C13	2.064(6)	C2 - C3	1.360(15)
		C3 - C4	1.357(12)
C11 - C12	1.395(6)	C2 - N1	1.517(16)
C11 - H11	1.00	C4 - N2	1.429(18)
		N1 - O2	1.210(32)
H12 - O1	2.439(8)	N1 - O3	1.216(15)
H12 - O2	2.277(10)	N2 - O4	1.228(9)
O1 - O2	2.641(10)	(C3 - H1)	1.029(11)

(b) Bond angles (deg)

C2 - C1 - C2'	107.6(14)
O1 - C1 - C2	126.2(7)
C1 - C2 - C3	126.9(12)
C1 - C2 - N1	117.2(9)
C2 - C3 - C4	118.5(12)
C3 - C4 - C3'	121.5(14)
C3 - C4 - N2	119.2(7)
C4 - N2 - O4	118.6(8)
O4 - N2 - O4	122.9(15)
C2 - N1 - O2	113.(2)
C2 - N1 - O3	123.(1)
O2 - N1 - O3	123.(2)

dihedral angle: (z = 0.25 plane) and

(N1, O2, O3 plane) 26.6°

Table XIII
Interionic contacts

Fe - O1	3.951(11)	O3 - H13	2.383(12)
O1 - H11	2.439(18)	O3 - H13	2.668(12)
O1 - O2	2.641(10)	O3 - O3	1.928(29)
O1 - C11	3.234(12)	O4 - H11	2.505(7)
O1 - O4	3.830(6)	O2 - H11	2.277(10)

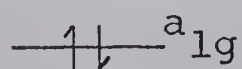
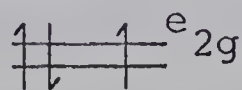
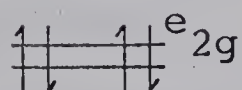
DISCUSSION

The compound ferrocinium picrate is an ionic species consisting of the ferrocinium ion $\text{Fe}(\text{Cp})_2^+$ in which the iron may be considered to be in the formal +3 oxidation state, and the picrate ion $[\text{H}_2\text{C}_6\text{N}_3\text{O}_7]^-$.

The ferrocinium ion has two cyclopentadienyl rings which are related by a mirror plane, and thus unlike ferrocene in the solid state^{35,36}, the rings are eclipsed. Ferrocene in the gaseous state⁴⁰, as well as ruthenocene²² in the solid state, are also eclipsed. The Fe-C distances were found to be longer (average $2.070(5)\text{\AA}$) than those of ferrocene³⁶ ($2.047(5)\text{\AA}$) while the average C-C distance of $1.395(6)\text{\AA}$ for ferrocinium was shorter ($1.416(4)\text{\AA}$) as obtained from the hindered rotor model.

Although distortions from D_{5h} symmetry may be present, because of the disorder in the crystal and large oscillations of the rings (RMS $\sim 18^\circ$), the results of any individual atom refinement would likely be unreliable. Evidence of this can be seen in the results of Pettersen⁴¹ in which the C-C distance ranges from 1.33\AA to 1.44\AA with an error of $\pm 0.02\text{\AA}$ or greater. Use of the hindered rotor model allows for the oscillation of the rings and should give a good average value for the C-C distance. The lengthening of the Fe-C distances and shortening of the

C-C distances in the ferrocenium ion is in agreement with the molecular orbital calculations^{42,43} and spectral results⁴⁴⁻⁴⁸ which indicated that the ground state of ferrocene in the eclipsed formation is $E_{2g}[(a_{1g})^2(e_{2g})^4]$ and that of the ferrocenium ion is $E_{2g}[(a_{1g})^2(e_{2g})^3]$. Since a_{1g} and e_{2g} are the highest bonding orbitals, removal of one electron from the e_{2g} level



ferrocene

ferrocenium ion

would weaken the Fe-C(ring) bonds and result in a longer Fe-C distance in ferrocenium ion than in ferrocene. It would also strengthen the C-C bond thus shortening it in

$\text{Fe}(\text{Cp})_2^+$ since an electron is being removed from a C-C antibonding orbital. These calculations were done⁴² assuming D_{5d} symmetry. However, the symmetry of the ferrocenium ion in this structure is D_{5h} rather than D_{5d} . Molecular orbital calculations using D_{5h} symmetry give the same pattern of energy levels. If the assumption is made that the order of the energy levels is the same in the eclipsed form as in the staggered form, then a ground state of the ferrocenium ion would be $E_2'(a_1'^2, e_2'^3)$ and the argument given above would still apply.

The picrate ion lies on a mirror plane at $z = 0.2500$ and across a second mirror at $x = 0.0$. However, three atoms, namely those of the ortho nitro group (N1, O2 and O3), were allowed to move off the mirror, and have "half atoms" above and below the $z = \frac{1}{4}$ mirror. The reasons for this were discussed in some detail earlier.

Although because of the disorder problem the results are somewhat suspect, the geometry calculated from the disordered picrate model agrees fairly well with those of potassium and ammonium picrate³⁹. The N-O distances within the ortho nitro group ($1.210(32)\text{\AA}$ and $1.216(15)\text{\AA}$) are the same within experimental error as those of the potassium and ammonium structures ($1.232(5)\text{\AA}$, $1.229(5)\text{\AA}$, $1.237(5)\text{\AA}$ and $1.206(5)\text{\AA}$) respectively. The ortho nitro group was found to be twisted 26.6 degrees from the

plane of the benzene ring, while in the potassium and ammonium compounds it was $25^{\circ} - 27^{\circ}$. The C-N bond is slightly shorter ($1.429(18)\text{\AA}$) than those found by Maartmann-Moe ($1.457(6)\text{\AA}$ and $1.461(5)\text{\AA}$). The rest of the molecule also has a geometry very similar to the potassium and ammonium picrates.

Chapter III

THE CRYSTAL STRUCTURE OF HYDRIDOTRIPHENYLSILYL (π -CYCLOPENTADIENYL) DICARBONYL MANGANESE

INTRODUCTION

The ultraviolet irradiation of metal carbonyls in the presence of substituted silanes has lead to the preparation of a new class of metal-silicon hydrides. These compounds seem to fall into two groups; for example, those prepared from triaryl silanes such as $(\pi\text{-C}_5\text{H}_5)(\text{OC})_2\text{Mn}(\text{H})\text{Si}(\text{C}_6\text{H}_5)_3$ (1), and those containing the more electronegatively substituted silyls such as $(\pi\text{-C}_5\text{H}_5)(\text{OC})_2\text{Mn}(\text{H})\text{SiCl}_3$ (2). The latter group of compounds shows well defined infrared bands in the region of "normal" terminal metal-hydrogen stretching frequencies (1887 cm^{-1} , deuteride 1360 cm^{-1})⁴⁹. The former group, however, showed no bands in either terminal or bridging M-H regions. The solid state Raman spectra⁴⁹ shows a weak broad band at 1900 cm^{-1} which shifts to 1350 cm^{-1} in the deuteride, these bands being in what is considered the terminal metal-hydrogen region. In order to understand the differences in the spectral results, the crystal structures of (1) and (2) were undertaken.

EXPERIMENTAL

The compound hydridotrichlorosilyl (π -cyclopentadienyl) dicarbonyl manganese, $(\pi\text{-C}_5\text{H}_5)(\text{OC})_2\text{Mn}(\text{H})\text{SiCl}_3$ (2), forms light brown needle crystals. Rotation, Weissenberg ($h0\ell$ - $h4\ell$ and $0k\ell$ - $9k\ell$) and precession ($h0\ell$ - $h3\ell$) photographs using several crystals showed the crystals to be monoclinic with unit cell $a = 14.71\text{\AA}$, $b = 11.46\text{\AA}$, $c = 14.85\text{\AA}$ and $\beta = 107.75^\circ$. Systematic absences of $h k \ell$; $h + \ell = 2n + 1$, $0 k 0$; $k = 2n + 1$ and $h 0 \ell$; $\ell = 2n + 1$ fixed the space group at $B2_1/c$ (a non-standard setting of $P2_1/c$). Eight molecules per unit cell gave a calculated density of 1.75 gm/cc. Since the molecule does not possess a centre of symmetry, it is most likely that the molecules lie in general positions with one molecule per asymmetric unit. The crystal was mounted on the PAILRED automatic diffractometer in order to collect data. However, measurement of standard reflections showed that decomposition of the crystal in the X-ray beam was so rapid (over 30% in 20 hours), an accurate crystal structure was impossible. The determination was terminated at this point.

The compound hydridotriphenylsilyl (π -cyclopentadienyl) dicarbonyl manganese (π -C₅H₅)(OC)₂Mn(H)Si(C₆H₅)₃ was prepared by J.K. Hoyano⁴⁹ by the ultraviolet radiation of (C₅H₅)Mn(CO)₃ in the presence of an excess of triphenylsilane and recrystallized from a mixed hexane-dichloromethane solvent system to obtain crystals suitable for single crystal X-ray diffraction study. The yellow crystals were mounted parallel to an extinction direction (a face diagonal). Rotation, Weissenberg (0k ℓ -2k ℓ) and precession (kh0-hk2 and h0 ℓ -h2 ℓ) photographs showed the crystal to be monoclinic. Systematic absences of 0k0, k = 2n+1 and h0 ℓ , h+ ℓ = 2n+1 fixed the space group as P2₁/n. The general positions in this space group are

$$\begin{array}{ccc} x & y & z \\ \bar{x} & \bar{y} & \bar{z} \\ \frac{1}{2}+x, \frac{1}{2}-y, \frac{1}{2}+z \\ \frac{1}{2}-x, \frac{1}{2}+y, \frac{1}{2}-z \end{array}$$

The unit cell dimensions were obtained from least squares refinement of 2 θ values of reflections that had been carefully centred in the counter window on the Picker four circle manual diffractometer, using CuK _{α 1} radiation (λ = 1.5405 $\overset{\circ}{\text{\AA}}$). The cell dimensions are a = 13.200(3) $\overset{\circ}{\text{\AA}}$, b = 17.328(4) $\overset{\circ}{\text{\AA}}$, c = 9.438(2) $\overset{\circ}{\text{\AA}}$ and β = 92.47(4) $^{\circ}$. The density was determined by

flotation to be 1.347(3) gm/cc which agrees well with the calculated density of 1.342 gm/cc ($Z = 4$, $V = 2156.8 \text{ \AA}^3$) and is consistent with one molecule per asymmetric unit.

The crystal used in data collection was a prism bound by the faces 010, $0\bar{1}0$, 203, $\bar{2}0\bar{3}$, 212 and $\bar{2}\bar{1}\bar{2}$ with dimensions 0.17 mm x .097 mm x 0.097 mm. It was mounted along a face diagonal, the a^* axis being coincident with the diffractometer ϕ axis. Data was collected manually on the Picker four circle diffractometer using CuK_α X-radiation and a coupled $\omega/2\theta$ scan technique. The angular settings (ϕ , χ , 2θ) were calculated using MIXG2 for unique sections of the reciprocal lattice limited by $\sin \theta/\lambda \leq 0.5$. A scan width of 2 degrees at a scan speed of 2 degrees per minute was employed with a stationary background count of 30 seconds on each side of the scan. The beam was monochromated by means of a graphite crystal (002 reflection). The diffracted beam collimator aperture was 0.5 mm and the pulse height analyser was centred on the CuK_α peak. Intensities were calculated using $I = P - t(B_1 + B_2)$ where P is the peak count, B_1 and B_2 are the background counts and t is the ratio of total time of peak count to total time of background count. Two thousand, two hundred and thirteen unique (hkl

and $h\bar{k}\bar{l}$, h , k and l positive) reflections were measured of which 662 were rejected using the criteria (1) $I \leq 0$, or (2) $I \leq 3\sigma$ where $\sigma = [P + t^2(B_1 + B_2)]^{\frac{1}{2}}$. Lorentz and polarization corrections were applied to the remaining 1551 reflections. Twelve well distributed standard reflections measured at various times throughout the data collection indicated no decomposition of the crystal. The maximum variation was ± 5 sigma. There was no apparent pattern in the peak height variation. Absorption corrections ($\mu_{\text{Cu}} = 58.8 \text{ cm}^{-1}$) were applied to the data. Transmission factors ranged from 0.51 to 0.63 with the majority lying close to 0.59.

SOLUTION AND REFINEMENT

The three dimensional Patterson map gave a set of three peaks agreeing with a manganese position of $x = -0.112$, $y = 0.250$ and $z = 0.060$. A further set of peaks gave the position of the silicon atom as $x = -.043$, $y = 0.199$ and $z = -.165$. A structure factor calculation using these co-ordinates returned

$$R_1 = \frac{\sum ||F_o| - |F_c||}{\sum |F_o|} \text{ of } 51.6\% \text{ and } R_2 = \left[\frac{\sum w ||F_o| - |F_c||^2}{\sum w |F_o|^2} \right]^{\frac{1}{2}}$$

of 64.0%. The weights used were calculated from

$\sigma(F): w = 1/\sigma(F)^2$ where $\sigma(F) = [P + t^2(B_1 + B_2) + p^2 I]^{\frac{1}{2}}$, $p = .03$. The scattering factors were calculated from Cromer's co-efficients⁵⁰ for all atoms except hydrogen. These were obtained from Mason and Robertson⁵¹.

A difference map phased using the manganese and silicon atom positions revealed the location of the two carbonyl groups. A further difference map, phased using the six atoms located gave the positions of all remaining non-hydrogen atoms. After one cycle in which the phenyl groups were treated as rigid bodies and the cyclopentadienyl ring as a hindered rotor, the R factor dropped to $R_1 = 28.6\%$ and $R_2 = 38.9\%$. The starting parameters for the four rings were calculated using MMR with a radius of 1.397\AA for the rigid bodies. At the end of isotropic refinement the R factor was $R_1 = 8.6\%$ and

$R_2 = 10.5\%$. A difference map at this time showed the hydrogen atoms attached to the aromatic rings, the peak heights ranging from $0.76\text{e}^{-\text{\AA}^3}$ to $0.38\text{e}^{-\text{\AA}^3}$. These were added as rigid bodies with the same orientation and centre of gravity as the carbon rings to which they were bonded and with a C-H distance of 1.0\AA . The hydrogens on the cyclopentadienyl ring were added as a hindered rotor rather than as a rigid body. The difference map also showed a peak in a bridging position between the manganese and silicon atoms. This peak was of the same magnitude as those of the twenty aromatic hydrogens ($\sim 0.5\text{e}^{-\text{\AA}^3}$) and was the twelfth largest peak on the map. Adding the hydrogen atoms, giving the manganese, silicon and carbonyls anisotropic temperature factors of the form: $\exp(-(\beta_{11}h^2 + \beta_{22}k^2 + \beta_{33}l^2 + \beta_{12}hk + \beta_{13}hl + \beta_{23}kl))$ and allowing for anomalous dispersion ($\Delta f'_{\text{Mn}} = -.050$, $\Delta f''_{\text{Mn}} = 2.90$, $\Delta f'_{\text{Si}} = 0.20$ and $\Delta f''_{\text{Si}} = 0.40$) dropped the R factor to $R_1 = 6.2\%$, $R_2 = 7.3\%$ after four cycles. A difference map showed that the peak previously found at the bridging position ($x = -.009$, $y = 0.248$ and $z = -.014$) was still present with a peak height of $0.53\text{e}^{-\text{\AA}^3}$ and was the second largest peak present. This hydrogen peak is of very similar density to those of the hydrogen peaks

found by Ibers et al.⁵² in the structure $\text{CoH}(\text{N}_2)(\text{P}(\text{C}_6\text{H}_5)_3)_3$. The largest peak was located near the para carbon of a phenyl ring (C44), and since it disappeared on the low angle ($\sin \theta/\lambda \leq 0.30$) difference map, was most likely due to the thermal motion of the phenyl ring. To further check the validity of the hydride peak, difference maps were calculated using various $\sin \theta/\lambda$ data cutoffs ($\sin \theta/\lambda \leq 0.35$, ≤ 0.30 , ≤ 0.25 , and ≤ 0.20). In all four maps, the hydrogen peak was the largest peak on the map and was approximately a factor of two greater than the next highest. The peak positions remained constant on all four maps while its height varied according to the $\sin \theta/\lambda$ cutoff. Table XIV gives the observed and calculated peak heights. The calculations are the same as those of La Placa and Ibers⁵.

The hydridic hydrogen was added as an isotropic atom and allowed to refine. At the same time the carbons in the phenyl rings were refined using the individual atom, anisotropic temperature factor model. Three cycles of refinement gave a final R factor of $R_1 = 4.2\%$ and $R_2 = 4.7\%$. In the final cycle no shift was greater than 1/10 of an e.s.d. A final difference map showed no peak greater than $0.19\text{e}/\text{\AA}^3$

Table XIV

Hydrogen peak as a function of $\sin \theta/\lambda$

$\sin \theta/\lambda$ cutoff(\AA^{-1})	Number of terms in the unique section	$\rho^E (\text{e}^-/\text{\AA}^3)$	$\rho^C (\text{e}^-/\text{\AA}^3)^b$
0.20	149	0.20	0.16
0.25	278	0.29	0.25
0.30	464	0.37	0.34
0.35	703	0.41	0.42
0.50 ^a	1551	0.53	0.53

a full data set

$$b \quad \rho^C = \frac{1}{2\pi^2} \int_0^{s_0} (1+a^2 s^2/4)^{-2} \exp(-Bs^2/16\pi^2) s^2 ds$$

$$s = 4\pi \sin \theta/\lambda$$

a = Bohr's radius

or less than $-0.20\text{e}^{-}/\text{\AA}^3$.

Table XV lists the observed and calculated structure factor amplitudes; $10|F_o|$ and $10|F_c|$, both in absolute units of electrons. The final positional parameters for the individual atoms are given in Table XVI and their anisotropic temperature factors in Table XVII. Parameters for the hindered rotors are found in Table XVIII. The estimated standard deviations were obtained from the inverse matrix of the final least squares cycle.

Table XV

Observed and calculated structure factor amplitudes
(x 10) in electrons for $(\pi\text{-C}_5\text{H}_5)(\text{OC})_2\text{Mn}(\text{H})\text{Si}(\text{C}_6\text{H}_5)_3$

	L	A	N	3	2	724	327	7	11	200	204	-9	3	176	125	-3	2	297	299	5	14	143	143	-3	10	217	228	-3	9	313	314	-3	13	217	218
2	0	179	104	-2	2	1123	1217	8	11	184	182	-10	5	266	273	-7	2	121	105	6	14	154	162	-4	10	130	139	-4	9	131	111	-2	13	122	105
4	0	576	562	-2	2	269	263	-1	11	171	172	1	0	152	146	-8	2	171	172	-1	14	278	274	-3	10	140	144	-5	9	217	236	-5	0	204	489
6	0	255	416	-7	2	204	203	-3	11	257	262	2	5	156	166	-10	2	102	48	0	15	366	356	0	11	124	125	-6	8	122	123	-5	0	504	489
12	0	409	406	-6	2	167	165	-3	11	218	214	3	6	184	181	-11	2	241	255	1	13	142	146	1	11	229	228	0	10	224	252	3	0	303	304
3	1	122	117	10	2	142	149	-5	11	352	360	6	6	202	203	1	3	317	327	2	12	276	213	2	11	174	175	0	10	277	295	5	0	120	103
5	1	1042	1105	-1	2	173	166	-6	11	336	331	7	4	244	234	2	4	146	133	-13	13	136	150	4	11	307	300	-2	10	176	205	-5	0	504	489
9	1	342	342	-3	2	222	217	-5	11	259	251	-1	6	236	242	3	3	223	228	-5	13	136	160	5	11	129	130	-3	10	113	224	-9	0	317	528
1	1	253	260	-4	2	70	70	0	12	224	223	-2	6	187	191	3	3	275	290	-1	14	240	220	-2	11	146	146	-5	10	153	256	1	1	95	130
10	1	112	110	-6	2	223	203	3	12	114	114	0	6	192	199	6	3	164	171	-2	16	214	214	-2	11	244	244	-7	10	171	296	6	1	149	167
11	1	121	107	-2	2	173	164	3	12	172	138	-5	6	227	277	7	3	403	397	0	17	190	190	-7	11	146	146	-5	10	153	256	1	1	95	130
1	2	358	371	-17	2	216	211	5	12	142	128	-6	6	254	263	-8	3	152	149	1	1	214	214	-2	11	146	146	-5	10	153	256	1	1	95	130
2	2	345	345	-12	2	117	117	0	12	168	167	-7	6	100	113	11	3	124	152	0	0	166	187	3	12	127	127	-5	11	146	146	-5	10	153	256
4	2	344	344	0	3	745	745	8	12	105	105	-8	0	207	201	-2	3	152	149	0	0	166	187	3	12	127	127	-5	11	146	146	-5	10	153	256
2	2	344	344	0	3	745	745	8	12	105	105	-8	0	207	201	-2	3	152	149	0	0	166	187	3	12	127	127	-5	11	146	146	-5	10	153	256
2	2	344	344	0	3	745	745	8	12	105	105	-8	0	207	201	-2	3	152	149	0	0	166	187	3	12	127	127	-5	11	146	146	-5	10	153	256
2	2	344	344	0	3	745	745	8	12	105	105	-8	0	207	201	-2	3	152	149	0	0	166	187	3	12	127	127	-5	11	146	146	-5	10	153	256
2	2	344	344	0	3	745	745	8	12	105	105	-8	0	207	201	-2	3	152	149	0	0	166	187	3	12	127	127	-5	11	146	146	-5	10	153	256
2	2	344	344	0	3	745	745	8	12	105	105	-8	0	207	201	-2	3	152	149	0	0	166	187	3	12	127	127	-5	11	146	146	-5	10	153	256
2	2	344	344	0	3	745	745	8	12	105	105	-8	0	207	201	-2	3	152	149	0	0	166	187	3	12	127	127	-5	11	146	146	-5	10	153	256
2	2	344	344	0	3	745	745	8	12	105	105	-8	0	207	201	-2	3	152	149	0	0	166	187	3	12	127	127	-5	11	146	146	-5	10	153	256
2	2	344	344	0	3	745	745	8	12	105	105	-8	0	207	201	-2	3	152	149	0	0	166	187	3	12	127	127	-5	11	146	146	-5	10	153	256
2	2	344	344	0	3	745	745	8	12	105	105	-8	0	207	201	-2	3	152	149	0	0	166	187	3	12	127	127	-5	11	146	146	-5	10	153	256
2	2	344	344	0	3	745	745	8	12	105	105	-8	0	207	201	-2	3	152	149	0	0	166	187	3	12	127	127	-5	11	146	146	-5	10	153	256
2	2	344	344	0	3	745	745	8	12	105	105	-8	0	207	201	-2	3	152	149	0	0	166	187	3	12	127	127	-5	11	146	146	-5	10	153	256
2	2	344	344	0	3	745	745	8	12	105	105	-8	0	207	201	-2	3	152	149	0	0	166	187	3	12	127	127	-5	11	146	146	-5	10	153	256
2	2	344	344	0	3	745	745	8	12	105	105	-8	0	207	201	-2	3	152	149	0	0	166	187	3	12	127	127	-5	11	146	146	-5	10	153	256
2	2	344	344	0	3	745	745	8	12	105	105	-8	0	207	201	-2	3	152	149	0	0	166	187	3	12	127	127	-5	11	146	146	-5	10	153	256
2	2	344	344	0	3	745	745	8	12	105	105	-8	0	207	201	-2	3	152	149	0	0	166	187	3	12	127	127	-5	11	146	146	-5	10	153	256
2	2	344	344	0	3	745	745	8	12	105	105	-8	0	207	201	-2	3	152	149	0	0	166	187	3	12	127	127	-5	11	146	146	-5	10	153	256
2	2	344	344	0	3	745	745	8	12	105	105	-8	0	207	201	-2	3	152	149	0	0	166	187	3	12	127	127	-5	11	146	146	-5	10	153	256
2	2	344	344	0	3	745	745	8	12	105	105	-8	0	207	201	-2	3	152	149	0	0	166	187	3	12	127	127	-5	11	146	146	-5	10	153	256
2	2	344	344	0	3	745	745	8	12	105	105	-8	0	207	201	-2	3	152	149	0	0	166	187	3	12	127	127	-5	11	146	146	-5	10	153	256
2	2	344	344	0	3	745	745	8	12	105	105	-8	0	207	201	-2	3	152	149	0	0	166	187	3	12	127	127	-5	11	146	146	-5	10	153	256
2	2	344	344	0	3	745	745	8	12	105	105	-8	0	207	201	-2	3	152	149	0	0	166	187	3	12	127	127	-5	11	146	146	-5	10	153	256
2	2	344	344	0	3	745	745	8	12	105	105	-8	0	207	201	-2	3	152	149	0	0	166	187	3	12	127	127	-5	11	146	146	-5	10	153	256
2	2	344	344	0	3	745	745	8	12	105	105	-8	0	207	201	-2	3	152	149	0	0	166	187	3	12	127	127	-5	11	146	146	-5	10	153	256
2	2	344	344	0	3	745	745	8	12	105	105	-8	0	207	201	-2	3	152	149	0	0	166	187	3	12	127	127	-5	11	146	146	-5	10	153	256
2	2	344	344	0	3	745	745	8	12	105	105	-8	0	207	201	-2	3	152	149	0	0	166	187	3	12	127	127	-5	11	146	146	-5	10	153	256
2	2	344	344	0	3	745	745	8	12	105	105	-8	0	207	201	-2	3	152	149	0	0	166	187	3	12	127	127	-5	11	146	146	-5	10	153	256
2	2	344	344	0	3	745	745	8	12	105	105	-8	0	207	201	-2	3	152	149	0	0	166	187	3	12	127	127	-5	11	146	146	-5	10	153	256
2	2	344	344	0	3	745	745	8	12	105	105	-8	0	207	201	-2	3	152	149	0	0	166	187	3	12	127	127	-5	11	146	146	-5	10	153	256
2	2	344	344	0	3	745	745	8	12	105	105	-8	0	207	201	-2	3	152	149	0	0	166	187	3	12	127	127	-5	11	146	146	-5	10	153	256
2	2	344	344	0	3	745	745	8	12	105	105	-8	0	207	201	-2	3	152	149	0	0	166	187	3	12	127	127	-5	11	146	146	-5	10	153	256
2	2	344	344	0	3	745	745	8	12	105	105	-8	0	207	201	-2	3	152	149	0	0	166	187	3	12	127	127	-5	11	146	146	-5	10	153	256
2	2	344	344	0	3	745	745	8	12	105	105	-8	0	207	201	-2	3	152	149	0	0	166	187	3	12	127	127	-5	11	146	146	-5	10	153	256
2	2	344	344	0	3	745	745	8	12	105	105	-8	0	207	201	-2	3	152	149	0	0	166	187	3	12	127	127	-5	11	146	146	-5	10	153	256
2	2	344	344	0	3	745	745	8	12	105	105	-8	0	207	201	-2	3	152	149	0	0	166													

Table XVI

Final atom co-ordinates for $(C_5H_5)(OC)_2Mn(H)Si(C_6H_5)_3$

Atom	x	y	z
Mn	-0.1187(1)	0.2382(1)	0.0658(1)
Si	-0.0492(1)	0.1991(1)	-0.1567(1)
Cl	-0.0827(4)	0.3331(4)	0.1165(6)
O1	-0.0578(3)	0.3956(3)	0.1474(5)
C2	-0.2109(4)	0.2782(3)	-0.0535(6)
O2	-0.2751(3)	0.3056(2)	-0.1262(5)
C21	-0.1459(4)	0.1368(3)	-0.2546(6)
C22	-0.2167(5)	0.1689(3)	-0.3513(6)
C23	-0.2910(5)	0.1250(5)	-0.4210(7)
C24	-0.2964(6)	0.0476(5)	-0.3957(8)
C25	-0.2278(7)	0.0136(4)	-0.3015(8)
C26	-0.1538(5)	0.0574(4)	-0.2314(6)
C31	-0.0206(4)	0.2819(3)	-0.2787(6)
C32	-0.0198(4)	0.3590(3)	-0.2374(6)
C33	0.0085(5)	0.4173(3)	-0.3274(7)
C34	0.0374(4)	0.4005(3)	-0.4615(7)
C35	0.0382(4)	0.3247(4)	-0.5066(6)
C36	0.0096(4)	0.2661(3)	-0.4157(6)
C41	0.0721(4)	0.1415(3)	-0.1424(6)
C42	0.0997(5)	0.9020(3)	-0.2512(6)
C43	0.1950(6)	0.0545(4)	-0.2480(8)
C44	0.2608(5)	0.0676(4)	-0.1333(10)
C45	0.2378(5)	0.1154(4)	-0.0260(8)
C46	0.1435(4)	0.1515(3)	-0.0306(7)
H1	-0.0132(32)	0.2467(23)	0.0020(45)

Table XVII

Anisotropic temperature factors ($\times 10^4$)

Atom	β_{11}	β_{22}	β_{33}	β_{12}	β_{13}	β_{23}	Equivalent isotropic B
Mn	45.1(6)	28.1(4)	89.2(1)	0.0(4)	2.6(6)	1.9(6)	3.23
Si	45(1)	19(1)	89(2)	0(6)	0(1)	3(1)	2.88
Cl	60(5)	37(3)	107(9)	3(3)	11(5)	-4(4)	4.18
O1	100(4)	36(2)	172(7)	-12(2)	31(4)	-24(3)	5.76
C2	43(4)	27(2)	132(9)	-3(3)	10(5)	-3(4)	3.66
O2	56(3)	46(2)	172(7)	16(2)	-13(4)	17(3)	5.22
C21	59(5)	24(3)	85(8)	-9(3)	12(5)	-5(3)	3.35
C22	61(4)	34(3)	113(9)	-11(3)	-17(6)	0(4)	4.17
C23	63(5)	56(4)	154(11)	-16(4)	-21(6)	-9(6)	5.56
C24	105(7)	53(4)	164(13)	-37(4)	-8(7)	-24(5)	6.52
C25	141(8)	29(3)	182(13)	-32(4)	10(8)	-8(5)	6.58
C26	85(6)	30(3)	126(10)	-9(3)	-8(6)	1(4)	4.67
C31	35(4)	20(2)	90(9)	-4(2)	-12(4)	8(3)	2.71
C32	63(5)	26(3)	79(8)	-2(3)	12(5)	9(4)	3.43
C33	78(5)	24(3)	133(10)	-2(3)	8(6)	-3(4)	4.34
C34	58(6)	22(3)	151(11)	-4(3)	-2(6)	30(4)	4.05
C35	59(5)	41(3)	79(8)	-1(3)	-5(5)	17(4)	3.99
C36	61(4)	23(2)	92(8)	0(2)	-11(5)	0(4)	3.47
C41	48(4)	21(2)	102(9)	1(2)	10(5)	14(4)	3.16
C42	88(6)	31(3)	118(10)	9(3)	12(6)	5(4)	4.65
C43	109(7)	33(3)	187(13)	24(4)	65(8)	10(5)	6.00
C44	71(6)	34(3)	257(16)	18(3)	44(8)	40(6)	6.01
C45	50(5)	42(3)	210(13)	6(3)	2(6)	23(6)	5.35
C46	46(5)	33(3)	164(11)	3(3)	10(6)	15(4)	4.35
H1							3.31(99)

Table XVIII

(a) Rigid body and hindered rotor parameters

Ring	x	y	z	B	Bd	R	D	E	ξ
1	-.2218	.0917	-.3253			2.392	3.313	2.469	4.785
2	.0088	.3414	-.3717			2.392	3.233	1.263	3.319
3	.1678	.1040	-.1389			2.392	2.257	2.480	6.524
4	-.1448(2)	.1654(1)	.1906(3)	3.91(8)	3.36(28)	1.203(2)	2.351(3)	2.849(3)	4.461(3)
5	-.1448	.1654	.1903	3.91	3.36	2.203	2.351	2.849	4.461

The parameters for which no estimated errors are given were not refined.

Table XVIII
(Contd.)

(b) Derived positional parameters for hindered rotors and rigid bodies

Ring	Atom	x	y	z
1	H22	-.2128	.2274	-.3701
	H23	-.3415	.1510	-.4899
	H24	-.3505	.0152	-.4451
	H25	-.2308	-.0440	-.2805
	H26	-.1021	.0324	-.1607
2	H32	-.0427	.3657	-.1358
	H33	.0051	.4708	-.2833
	H34	.0566	.4465	-.5192
	H35	.0603	.3171	-.6076
	H36	.0126	.2119	-.4600
3	H42	.0490	.0831	-.3347
	H43	.2109	.0199	-.3283
	H44	.3297	.0408	-.1325
	H45	.2867	.1249	.0569
	H46	.1248	.1881	.0505
4	C11	-.1412	.2126	.2839
	C12	-.2287	.1915	.2007
	C13	-.2002	.1343	.1036
	C14	-.0952	.1207	.1267
	C15	-.0587	.1685	.2381
5	H11	-.1383	.2519	.3614
	H12	-.2985	.2132	.2091
	H13	-.2463	.1084	.0312
	H14	-.0538	.0824	.0736
	H15	.0129	.1710	.2777

RESULTS

The molecular geometry and the numbering system used is shown in Figure 16, while the geometry of the manganese atom is shown in Figure 17. The molecular packing as viewed down the c axis is depicted in Figure 18. The drawings were made using the programme ORTEP. Table XIX gives the bond lengths and angles within the molecule. Some intermolecular contacts are listed in Table XX. The results and estimated errors associated with them were calculated using the programme ORFFE2.

Figure 16

Molecular structure of $(\pi\text{-C}_5\text{H}_5)(\text{OC})_2\text{Mn}(\text{H})\text{Si}(\text{C}_6\text{H}_5)_3$

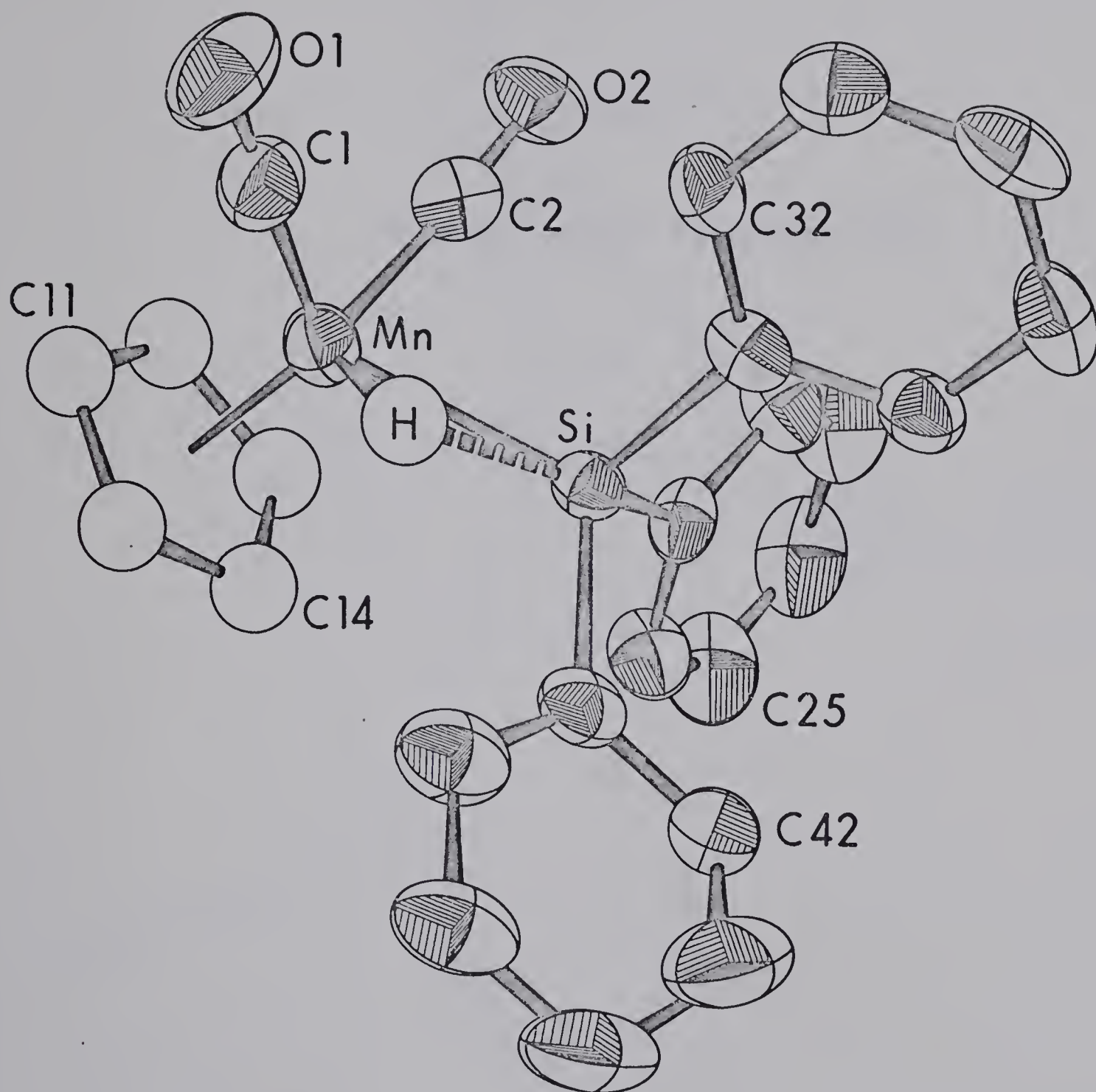


Figure 17

Geometry around manganese atom



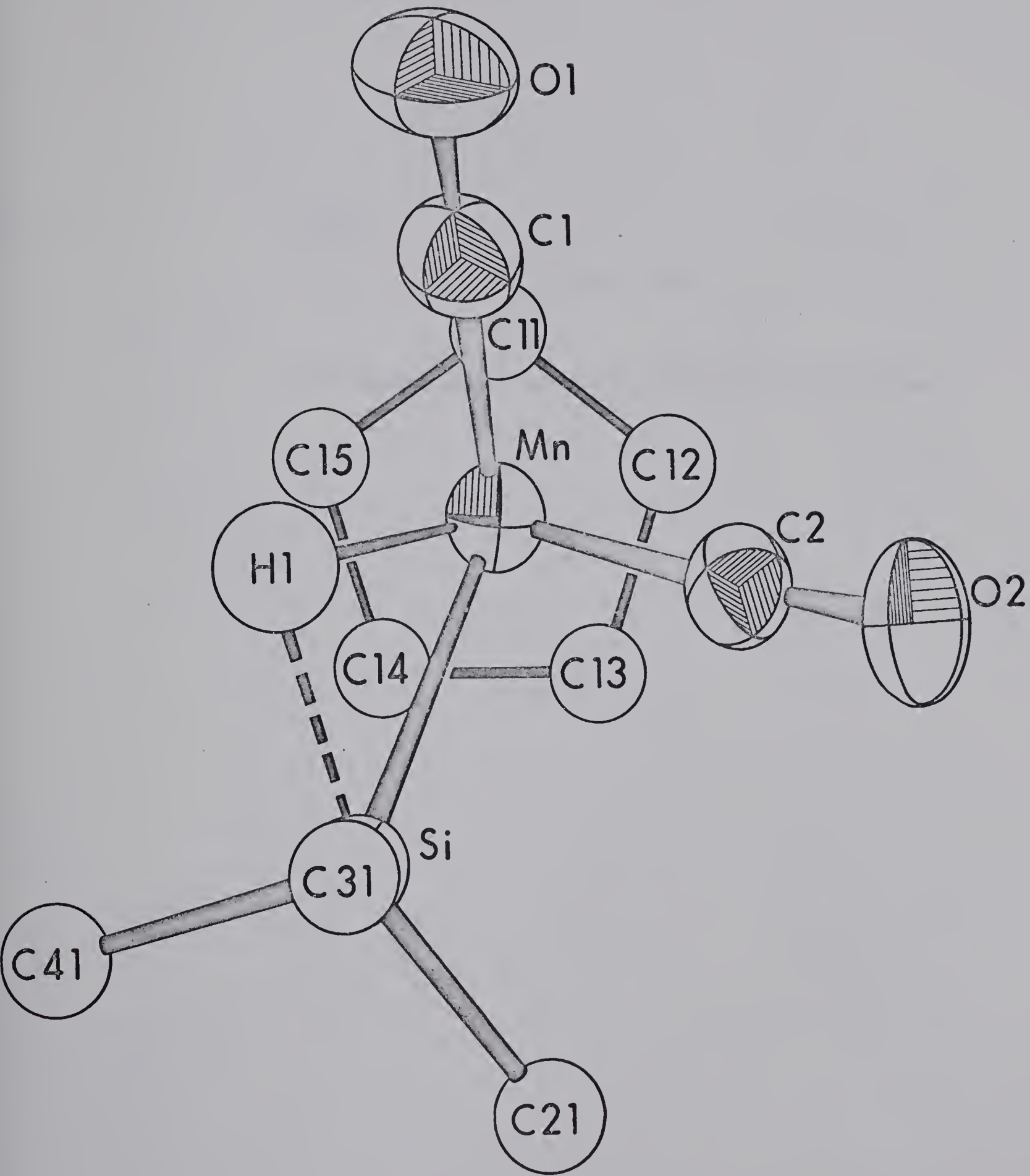


Figure 18

Molecular packing projected onto ab plane

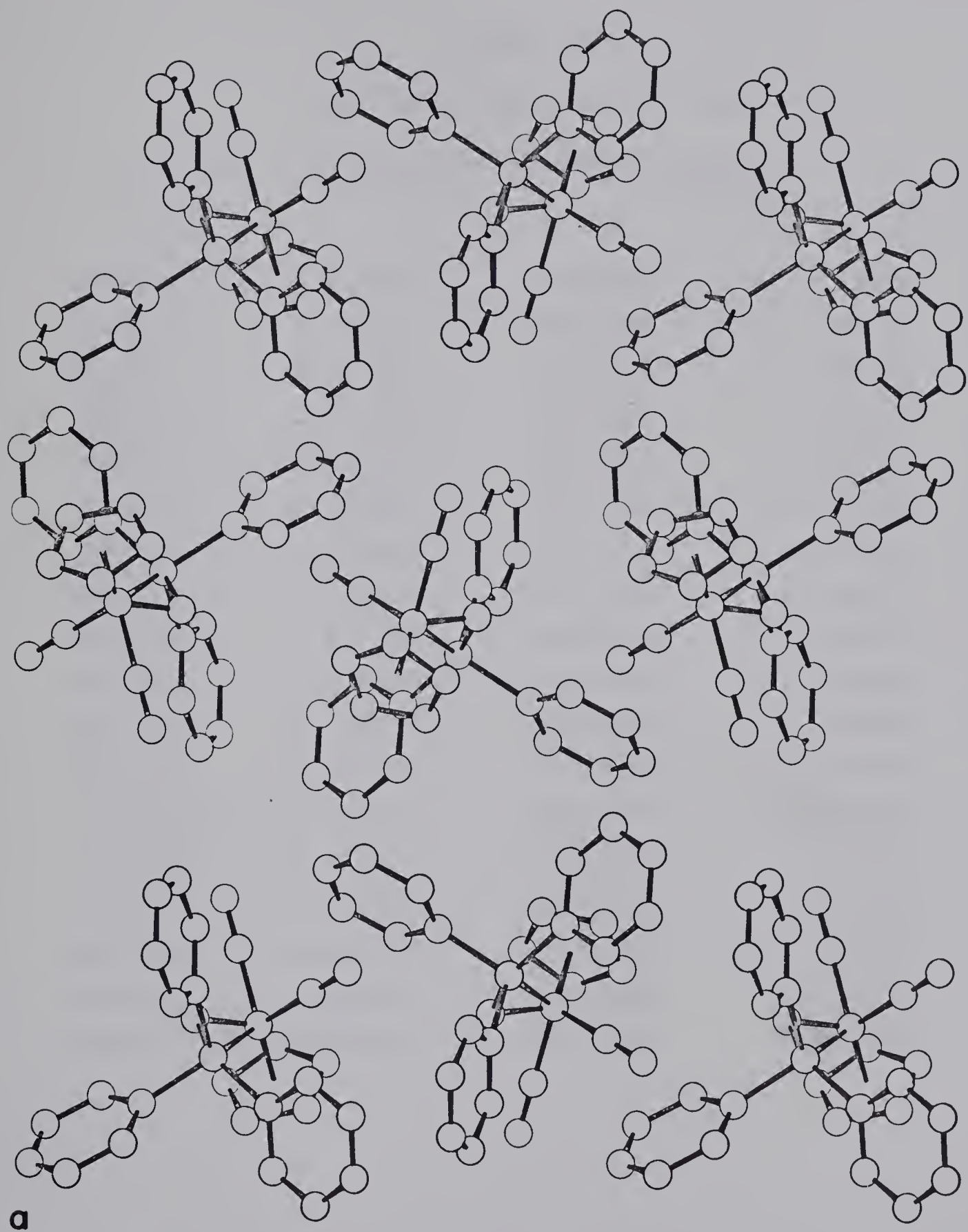


Table XIX

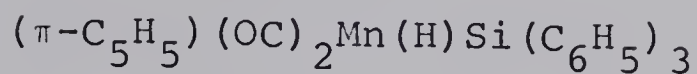
(a) Bond lengths ($\overset{\circ}{\text{\AA}}$) in



Mn-Si	2.424(2)	C21-C22	1.393(7)
Si-H	1.76(4)	C21-C26	1.398(7)
Mn-H	1.55(4)	C22-C23	1.385(8)
Mn-Cl	1.771(7)	C23-C24	1.365(9)
Mn-C2	1.764(7)	C24-C25	1.374(9)
Cl-O1	1.162(6)	C25-C26	1.384(8)
C2-O2	1.168(6)	C31-C32	1.391(7)
Mn-Cl1	2.140(4)	C31-C36	1.396(7)
Mn-Cl2	2.132(4)	C32-C33	1.382(7)
Mn-Cl3	2.136(4)	C33-C34	1.369(8)
Mn-Cl4	2.146(4)	C34-C35	1.382(8)
Mn-Cl5	2.148(4)	C35-C36	1.392(7)
ave Mn-C	2.140(4)	C41-C42	1.398(7)
Cl1-Cl2	1.414(3)	C41-C46	1.395(7)
Si-C21	1.883(5)	C42-C43	1.415(9)
Si-C31	1.888(5)	C43-C44	1.378(9)
Si-C41	1.886(5)	C44-C45	1.353(9)
Cl-H	2.08(4)	C45-C46	1.391(8)

Table XIX
(Contd.)

(b) Bond angles (degrees) in



Mn-H1-Si	94(2)	C26-C21-C22	116.3(5)
Mn-Si-H1	39(1)	C21-C22-C23	122.3(6)
Si-Mn-H1	46(2)	C22-C23-C24	119.9(6)
C1-Mn-C2	88.7(3)	C23-C24-C25	119.8(6)
Si-Mn-Cp	118.0(1)	C24-C25-C26	120.4(6)
C1-Mn-Cp	122.6(2)	C25-C26-C21	121.4(6)
C2-Mn-Cp	123.7(2)	C36-C31-C32	116.6(5)
H1-Mn-Cp	122.0(1)	C31-C32-C33	122.0(5)
C1-Mn-H	77.0(1)	C32-C33-C34	120.4(5)
C2-Mn-H	109.0(2)	C33-C34-C35	119.5(5)
Mn-Si-C21	108.1(2)	C34-C35-C36	119.9(5)
Mn-Si-C31	114.1(2)	C35-C36-C31	121.6(5)
Mn-Si-C41	115.9(2)	C46-C41-C42	116.2(5)
C21-Si-C31	106.6(2)	C41-C42-C43	121.8(6)
C21-Si-C41	106.7(2)	C42-C43-C44	118.1(6)
C31-Si-C41	104.7(2)	C43-C44-C45	122.3(7)
O1-C1-Mn	178.6(6)	C44-C45-C46	118.8(6)
O2-C2-Mn	176.3(5)	C45-C46-C41	122.9(6)

Table XX

Intermolecular contacts

O1-H33	2.723 (5)	O1-H24	3.038 (5)
O1-H44	2.829 (5)	O1-H35	3.048 (5)
O1-H23	3.288 (5)	O2-H25	2.749 (4)
O2-H15	2.937 (4)	O2-H35	3.049 (4)
O2-H46	3.270 (5)	O2-H36	3.284 (4)
H22-H45	2.646 (1)	H22-H46	2.671 (1)
H22-H11	2.787 (1)	H23-H46	2.849 (1)
H25-H45	2.662 (1)	H33-H34	2.457 (1)
H33-H44	2.647 (1)	H34-H34	2.415 (1)
H34-H44	2.610 (2)	H34-H13	2.796 (1)
H36-H15	2.576 (1)	H36-H11	2.645 (1)

DISCUSSION

A drawing of the compound $(\pi\text{-C}_5\text{H}_5)(\text{OC})_2\text{Mn}(\text{H})\text{Si}(\text{C}_6\text{H}_5)_3$ is found in Figure 16. The geometry about the manganese atom (Figure 17) can be described as a distorted square pyramid and that of the silicon atom as a distorted tetrahedron. The molecule has a hydrogen cis to the silicon and close enough ($1.76(4)\text{\AA}$) to interact with it.

The $\text{Si-C}(\text{C}_6\text{H}_5)$ distances of $1.883(5)\text{\AA}$, $1.888(5)\text{\AA}$ and $1.886(5)\text{\AA}$ are the same within experimental error. There is some distortion from tetrahedral geometry about the silicon atom, the $\text{Mn-Si-C}(\text{C}_6\text{H}_5)$ angles varying from $108.1(2)^\circ$ to $115.9(2)^\circ$ and the $\text{C}(\text{C}_6\text{H}_5)\text{-Si-C}(\text{C}_6\text{H}_5)$ (different rings) from $104.7(2)$ to $106.7(2)$. Angles within the phenyl rings range from $116.3(5)$ at the attached carbons to $122.3(6)$. It will be noticed that in the phenyl rings there is a range in C-C bond lengths from $1.365(9)$ to $1.415(9)\text{\AA}$. This bond length range and distribution is not that expected for a single population of formally identical bond lengths with an estimated standard deviation of 0.008\AA as determined from the least squares refinement. The variation of bond lengths follows a definite pattern, the longest being those close to the silicon-bonded carbon (average 1.395\AA). The shortest, with an average C-C distance of

1.370Å, are those furthest away. A similar pattern was observed by Beauchamp et al.⁵³ and results from the wagging motion of the rings. The apparent shortening of the C-C bond is due to the same effect as discussed in the hindered rotor section; that is, the atoms move along an arc, the centre of which lies near the silicon atom. Attempts to describe the distribution of electron density resulting from such motion by means of anisotropic temperature factors would give an error in the centre of electron density of each atom, and hence in the calculated bond lengths. The greatest apparent shortening would occur between those atoms which have the greatest motion, i.e., CR3-CR4 and CR4-CR5 where R is the number of the ring (R = 2,3,4).

The Mn-C(π -C₅H₅) and Mn-C(CO) distances were not unusual and agree with those of such compounds as (C₆H₇)Mn(CO)₃⁵⁴, (C₁₀H₈)Mn₂(CO)₆⁵⁵, (π -C₅H₅)Mn(CO)₃²⁴ and [(π -C₅H₄CH₃)Mn(CO)₂]₂C₆H₄(As(CH₃)₂)₂⁵⁶, as can be seen in Table XXI. The variations, where significant, are consistent with the trends that would be predicted on the basis of the relative π acceptor properties of the other ligands present. In all these compounds the carbonyl groups lie trans to a delocalized π system. One would expect that the

Table XXI
Comparison of bond lengths

Molecule	Mn-C (π -C ₅ H ₅)	Mn-C (CO)	C-O
this work			
(π -C ₅ H ₅)Mn(CO) ₃	2.140 (4)	1.767 (7)	1.165 (6)
(C ₁₀ H ₈ Mn ₂ (CO) ₆	2.16 (2)	1.80 (2)	1.15 (3)
(C ₆ H ₇)Mn(CO) ₃	2.169 (4)	1.803 (5)	1.151 (6)
[(π -C ₅ H ₄ CH ₃)Mn(CO) ₂] ₂ diars	2.141 (9) to 2.219 (7)	1.788 (6)	1.157 (8)
	2.15 (3)	1.77 (3)	

Mn-C(O) distances trans to other carbonyls would be longer than those trans to π -ring systems. This more dramatic demonstration of the trans effect can be seen in such compounds as $\text{HMn}(\text{CO})_5$ ⁵⁷, $(\text{CO})_4\text{Mn}(\text{H})(\text{P}(\text{C}_6\text{H}_5)_2)_2\text{Mn}(\text{CO})_4$ ⁵⁸, $\text{Mn}_2(\text{CO})_{10}$ ⁵⁹, and $\text{HRe}_2\text{Mn}(\text{CO})_{14}$ ⁶⁰ in which Mn-C distances trans to carbonyl groups are longer (Table XXII).

The Mn-Si bond length of $2.424(2)\text{\AA}$ lies between that found in $(\text{OC})_4\text{Mn}(\text{Si}(\text{C}_6\text{H}_5)_2)_2\text{Mn}(\text{CO})_4$ ⁶¹ of 2.40\AA and that of $2.497(5)\text{\AA}$ in $\text{Mn}(\text{Si}(\text{CH}_3)_3)(\text{CO})_5$ ⁶². It is significantly shorter than the calculated single bond distance of 2.56\AA (using 1.17\AA ⁶³ for the radius of silicon and 1.39\AA ⁶⁴ for manganese), and would suggest some double bond character between the manganese and silicon.

The most exciting part of the work is the location of the bridging hydrogen. It lies in an otherwise empty space between the manganese and silicon atoms, and as described earlier was found to be the highest peak on all low angle difference maps. It is bonded to the manganese with a distance of $1.55(4)\text{\AA}$ or ca 1.62\AA when corrected for the use of a spherical electron distribution⁶⁵. This is close to the terminal Mn-H distance of 1.60\AA found by neutron diffraction techniques in $\text{HMn}(\text{CO})_5$ ⁵⁷, $1.60(12)\text{\AA}$ in $\text{RhH}(\text{CO})(\text{P}(\text{C}_6\text{H}_5)_3)_3$ ⁴

Table XXII
Comparison of bond lengths

Molecule	Mn-C trans to carbonyl	Mn-C not trans to carbonyl
$\text{HMn}(\text{CO})_5$	1.836(5)	1.821(9)
$\text{Mn}_2(\text{CO})_{10}$	1.830(8)	1.792(14)
$(\text{CO})_4\text{Mn}(\text{H})(\text{P}(\text{C}_6\text{H}_5)_2\text{Mn}(\text{CO})_4$	1.840(9)	1.788(9)
$\text{HRe}_2\text{Mn}(\text{CO})_{14}$	1.840(15)	1.80(3)

and 1.6(1) in $\text{CoH}(\text{N}_2)(\text{P}(\text{C}_6\text{H}_5)_3)_3$ ⁵², the latter two being done using X-ray methods. The Si-H distance of 1.76(4) Å, although considerably longer than the Si-H bond distance of 1.48 Å in SiH_4 ⁶⁶, is much shorter than would be expected for a normal Si-H intramolecular contact. Because silicon has a larger radius than carbon, one would expect the C...H contact to be less than that of Si...H. This is not the case, Cl-H being 2.08(4) Å and Si-H being 1.76(4) Å. This would suggest an interaction between the Si-H and formation of a weak bond. A similar type of interaction was found in $\text{bipy}(\text{OC})_3\text{ClMoSnCH}_3\text{Cl}_2$ ⁶⁷ in which the Sn-Cl distance of 2.805(4) Å was longer than would be expected for a bridging chlorine distance, but much shorter than Van der Waals distance (4.0 Å). The bond length evidence would suggest a bent 3 centre, 2 electron bond $\left[\text{Mn} \begin{array}{c} \text{H} \\ \diagup \quad \diagdown \end{array} \text{Si} \right]$ with the predominant atomic orbitals lying on the manganese and hydrogen atoms.

The presence of such a bridge would help to explain the different spectral and kinetic properties observed between this compound and the trihalogenated silyl derivatives. In the structure of $(\pi\text{-C}_5\text{H}_5)(\text{OC})\text{Fe}(\text{H})(\text{SiCl}_3)_2$ ⁶⁸, the hydrogen atom was not located. However, if an Fe-H distance of ca 1.6 Å is assumed, the hydrogen atom lying midway between the Si's and ca 120° from

Fe-Centre (ring), then one would obtain a Si-H distance of $2.3\overset{\circ}{\text{\AA}}$. This is greater than the $\text{Cl}\cdots\text{H}$ distance of $2.08(4)\overset{\circ}{\text{\AA}}$ and implies that there is no $\text{Fe}-\overset{\text{H}}{\text{Si}}$ bridge. The more electronegative chlorines pull sufficient electron density from the silicon to destroy any Si-H interaction. Alternatively, taking the electrostatic view of the hydrogen bond $\overset{\delta-}{\text{X}}-\overset{\delta+}{\text{H}}\cdots\overset{\delta-}{\text{Y}}$, the requirement is that the Y group (in this case silicon) have a partial negative charge. In SiCl_3 , the highly electronegative chlorine would be expected to produce a slight positive charge on the silicon, which while improving the electron acceptor properties as far as transition metal interactions were concerned, would reduce the ability to hydrogen bond. It would thus appear that the difference in physical properties of the two types of compound is due, in part at least, to the presence or lack of a bridging hydrogen. The chlorinated silyl compounds have a terminal M-H bond and clearly show this in the infrared. The phenyl substituted silyl class have a Mn-H bond that is close to, but not exactly the same as, a terminal bond. This difference, due to the $\text{Si}\cdots\text{H}$ interaction, could help to explain the lack of a peak in the infrared M-H terminal region, but the appearance of a weaker peak in the Raman.

The preliminary results of a kinetic study⁶⁹ of the displacement of a silane by a phosphine, which may be represented by the following general chemical equation:



may be rationalized in terms of the structural results reported here. In the case of the triphenyl silyl derivative, the dominant reaction proceeds via an initial dissociation in which triphenyl silane is eliminated and the second step involves very rapid phosphine attack on the electron deficient species $\pi\text{C}_5\text{H}_5\text{Mn}(\text{CO})_2$. The reaction is the first order in transition metal complex and zero order with respect to phosphine. However, for the trichlorosilyl derivative the main reaction is found to be bimolecular. The formation of a weak silicon hydrogen bond corresponds to a partial activation of the leaving group and significant formation of the silicon hydrogen bond in the trichlorosilyl derivative occurs only with the involvement of the entering phosphine ligand.

Chapter IV

THE CRYSTAL STRUCTURE OF DIHYDROOCTACARBONYL DIRHENIUM

INTRODUCTION

There has been great interest lately in the bonding and geometry of M-H-M bridges. The presence of the hydrogen should result in a lengthening of the M-M distance, since a 2 electron/3 centre bond is postulated. This was observed in the structure of $\text{HRe}_2\text{Mn}(\text{CO})_{14}$ ⁶⁰ in which a Re-Re distance of 3.39Å was found. This is longer than the Re-Re distance of 2.96Å to 3.02Å found in such structures as $\text{Re}_4(\text{CO})_{16}$ ⁷⁰. A similar result was found by Dahl in the structure of $(\text{HCr}_2(\text{CO})_{10})$ ⁷¹. In this molecule a Cr-Cr distance of 3.41Å was obtained. If a linear, symmetrical M-H-M bond is assumed, both compounds have a M-H distance of 1.70Å which is similar to the neutron diffraction result of 1.68(1)Å for Re-H in ReH_9 ⁷². It should also be possible to have a bent M-H-M bridge. One would then expect a M-M bond shorter than that found in the linear bridge, but longer than that in the unbridged M-M bond. This has been found in such structures as $(\text{H}_2\text{Re}_3(\text{CO})_{12})^-$ ⁷³ in which bridged Re-Re distances of 3.181(7)Å and 3.173(7)Å were found, and in $(\text{H}_6\text{Re}_4(\text{CO})_{12})$ ⁷⁴ having an average Re-Re distance of 3.160(7)Å. Having established a pattern for single bridged metal-metal bonds, one wonders what

would occur in doubly hydrogen bridged bonds such as occur in $\text{H}_2\text{Re}_2(\text{CO})_8$ and $\text{H}_2\text{Os}_3(\text{CO})_9$ ⁷⁵. These bridges are now 4 centre rather than 3 centre systems. Bonding schemes such as protonated metal-metal double bonds ($\text{M}=\text{M}$) have been suggested as well as a 4 centre delocalization of electrons. In an attempt to understand the bonding in such systems, the crystal structure of $\text{H}_2\text{Re}_2(\text{CO})_8$ was undertaken.

EXPERIMENTAL

The compound $\text{H}_2\text{Re}(\text{CO})_8$, molecular weight 598.4 gm/mole, was prepared by J.K. Hoyano and was recrystallized from a mixture of dichloromethane and hexane to obtain crystals suitable for single crystal X-ray diffraction studies. The yellow crystals were mounted parallel to the elongated edge of the crystal. Rotation, Weissenberg ($0k\ell$, $h0\ell$ - $h7\ell$, $hk0$ - hkl) and precession ($h0\ell$, $0k\ell$) photographs showed the crystal to be monoclinic with systematic absences $h0\ell$: $\ell = 2n+1$ and $0k0$: $k = 2n+1$, fixing the space group at $P2_1/c$. The unit cell was found to be $a = 8.96(2)\text{\AA}$, $b = 11.62(2)\text{\AA}$, and $c = 12.85(2)\text{\AA}$, $\beta = 109.2(1)^\circ$ from precession and Weissenberg photographs⁷⁶. The density was determined experimentally by means of flotation in a solution of Clerici's Reagent and water to be 3.16 gm/cc which agreed with the calculated density of 3.18(2) gm/cc for four molecules per unit cell. This would require that the molecules lie in general positions in the unit cell. Data within the 1.0\AA sphere were collected on the PAILRED automatic diffractometer using Weissenberg geometry and a moving crystal/stationary counter technique (ω scan). The hexagonal crystal was mounted along the b axis, and 13 layers ($h0\ell$ - $h12\ell$) were collected

using MoK_α -radiation. The X-ray beam was monochromated using a graphite single crystal (002 reflection).

A scintillation counter, aperture 2.5° , was used, with the pulse height analyser centred on the MoK_α peak and set to accept 90% of the peak. A half scan of 1.5° for the inner reflections ($\sin \theta/\lambda \leq 0.20$) and 1.0° for the outer reflections was used at a scan speed of $2.5^\circ/\text{min}$. The scan widths were increased to 2.0° and 1.5° for inner and outer reflections for the higher levels. The peak was scanned three times or until the peak count passed 4,000 counts. A stationary background was counted for twenty seconds on each side of the scan. Because the crystal was found to decompose in light, and to a lesser extent in radiation, the data were collected in a darkened room. Twelve standard reflections were measured at the end of each layer and a decomposition correction applied as a function of time and $\sin \theta/\lambda$. The intensities were calculated using $I = P - t(B_1 + B_2)$ where P is the peak count, B_1 , B_2 the background count, and t is the ratio of peak count time to that of total background. Four hundred and twenty-two reflections were rejected using the criteria (1) $I = 0$ and (2) $I \leq 3\Delta I$ where $\Delta I = (P + t^2(B_1 + B_2))^{1/2}$. Absorption corrections were applied ($\mu(\text{Mo}) = 217.\text{cm}^{-1}$, max crystal dimensions: .14 mm x .11 mm x .14 mm, mean $\mu_r = 1.4$).

SOLUTION AND REFINEMENT

The positions of the rhenium atoms were obtained from a three dimensional Patterson map as Re1: $x = 0.295$, $y = 0.130$, $z = 0.273$ and Re2: $x = -0.046$, $y = 0.130$, $z = 0.180$. A structure factor calculation using the form factors calculated from Cromer's coefficients⁵⁰ gave reliability factors of $R_1 = 27.9\%$, $R_2 = 31.3\%$, where $R_1 = \frac{\sum ||F_O| - |F_C||}{\sum |F_O|}$, $R_2 = \left[\frac{\sum w (|F_O| - |F_C|)^2}{\sum w |F_O|^2} \right]^{\frac{1}{2}}$, $w = 1/\sigma(F)^2$ and $\sigma(F) = (P + t^2(B_1 + B_2) + p^2 I)^{\frac{1}{2}}$, $p = 0.06$. A difference map phased by the two rhenium atoms gave the positions of all other non-hydrogen atoms. Four cycles of full matrix least squares refinement, minimizing the function $\sum w (|F_O| - |F_C|)^2$ with all atoms isotropic, reduced the R factors to $R_1 = 16.0\%$ and $R_2 = 20.4\%$. It was noticed that four layers (h1l, h2l, h3l and h9l) had much higher R factors (of the order of 20%) than the others (~10%). For this reason, data for these layers were recollected using a second crystal of similar dimensions (.12 mm x .13 mm x .10 mm). The data were treated as for the first crystal, corrected for absorption and combined with the previous

data using a second scale factor. Inclusion of the anomalous dispersion corrections for rhenium ($\Delta f' = -1.40$, $\Delta f'' = 7.70$) in the calculated structure factors and further refinement dropped the reliability factors to $R_1 = 10.3\%$, $R_2 = 12.0\%$. A difference map at this point showed peaks of 3 - 4 electrons near each of the rheniums, suggesting an error in the absorption correction. Since any anisotropic refinement would be meaningless for this reason, and also because of the large decomposition correction (~30%) applied, refinement was terminated at this point. In the final cycle, all atoms isotropic, no shift was greater than 1/10 the corresponding e.s.d. The observed and calculated structure factor amplitudes ($\times 10$) are listed in Table XXIII. The positional parameters and temperature factors are given in Table XXIV.

Table XXIII

Observed and calculated structure factor
amplitudes (x10) in units of electrons

K	L	FO	FC	K	L	FO	FC	K	L	FO	FC	K	L	FO	FC	K	L	FO	FC	K	L	FO	FC	K	L	FO	FC	K	L	FO	FC
0	4	1755	1930	3	8	1177	1082	3	7	793	736	2	9	1313	1337	1	7	1771	1818	0	6	1828	1847	3	-3	2564	1814	0	4	1416	1314
0	8	735	511	3	10	956	994	3	12	653	387	2	11	1541	1602	1	8	1500	1553	0	4	627	572	3	-4	1342	1158	0	6	663	683
0	12	651	620	3	11	716	602	3	11	716	602	2	11	1493	3770	1	9	1232	1338	0	-6	525	510	3	-5	797	649	0	-4	762	677
0	14	875	762	3	-3	3622	3575	3	-4	1372	1200	2	-2	1113	657	1	10	857	920	0	-8	1540	1545	3	-12	414	587	0	-6	1383	1244
1	2	774	1215	3	-5	2418	2955	3	-5	1490	1133	2	-9	1275	1220	1	-4	1691	1481	0	-10	1674	1672	3	-14	798	645	0	-8	937	764
1	4	955	1212	3	-6	2151	2325	3	-7	1469	1251	2	-11	1403	1423	1	-5	2429	2225	1	0	1454	1207	4	4	566	501	1	2	759	571
1	6	578	665	3	-7	1719	1635	3	-8	813	802	3	-8	1033	1053	1	-6	2173	2073	1	1	1682	1534	4	6	515	461	1	4	953	817
1	7	475	683	3	-8	1558	1458	3	-9	1558	1458	3	-9	1558	1458	1	-7	1740	1788	1	2	613	645	4	8	1321	1230	1	5	671	758
1	8	992	1300	4	1	557	514	4	4	2373	2533	3	-11	1558	1458	1	-8	926	900	1	3	775	721	4	15	0	1437	1	-5	850	640
1	9	1400	1809	4	4	2042	2266	4	6	2167	2176	4	6	2167	2176	1	-11	736	623	1	-1	1076	827	4	0	1335	1159	1	-6	1356	1099
1	10	1434	1474	4	6	2602	2840	4	8	876	845	3	4	1103	1078	1	-12	712	796	1	-4	1436	1007	4	-1	1141	1132	1	-7	1754	1406
1	11	1240	1733	4	8	2232	2310	4	12	1295	1229	3	7	940	889	1	-13	893	1014	1	-5	1926	1643	4	-2	1351	839	1	-8	1639	1464
1	12	1039	1421	4	10	1106	1127	4	15	0	1025	3	8	1346	1218	1	-14	971	1084	1	-7	2410	2405	5	0	2006	1727	1	-9	1761	1510
1	13	710	948	4	-1	1067	1293	4	0	755	664	3	9	966	977	1	-15	621	878	2	1	2642	2403	5	3	936	812	2	1	2044	1941
1	14	1822	1215	4	-2	1125	1474	4	-3	702	885	3	10	1032	1152	2	1	682	697	2	3	3166	3169	5	4	561	578	2	3	1824	1725
1	-3	882	832	4	-3	1224	1492	4	-5	674	794	3	15	0	1206	3	12	1257	1290	2	-5	2330	2191	5	-2	601	564	2	5	1115	958
1	-4	1251	1212	4	-4	1022	1285	4	-6	627	561	3	-11	1688	1362	2	5	1640	1711	2	7	773	1146	5	-8	871	1115	2	-1	1293	1045
1	-5	675	665	4	-5	1004	1218	5	0	1648	1376	3	-2	2234	1937	2	7	946	1085	2	-1	952	726	6	1	496	422	2	-5	1316	1158
1	-6	644	663	4	-6	692	857	5	1	2117	1950	3	-3	1019	777	2	-1	3632	2746	2	-7	1260	1066	6	3	709	637	2	-7	1375	1100
1	-8	1209	1300	5	2	493	595	5	2	2543	2482	3	-8	591	597	2	-3	4750	3755	2	-9	1734	1484	6	7	742	778	3	3	893	793
1	-9	1716	1509	5	3	689	782	5	3	2671	2815	3	-9	1120	1079	2	-5	2854	2593	2	-11	1394	1310	7	0	1995	1424	3	4	886	787
1	-10	1822	1509	5	4	1891	1966	5	4	1891	1966	3	-10	1220	1133	2	-7	1249	1210	3	0	944	902	7	1	1564	1378	3	5	957	738
1	-11	1658	1733	5	8	591	738	5	5	1458	1545	5	5	1458	1545	3	-12	724	724	3	-8	899	974	8	2	1541	1595	4	0	1439	1175
1	-12	1264	1421	5	9	938	984	5	6	1198	1195	3	-12	1697	1564	3	-15	1160	1254	3	2	1250	1213	7	3	808	830	3	-8	1249	1000
1	-13	872	958	5	10	723	723	5	15	0	614	3	-13	1400	1320	3	3	983	909	3	3	647	574	7	-1	616	532	3	-7	1555	1406
1	-14	666	696	5	11	814	801	5	-1	732	826	3	-14	1040	1031	3	4	858	838	3	-1	625	499	7	-2	529	445	3	-8	1694	1625
2	3	2866	3240	5	-2	661	721	5	-4	933	1105	3	-15	727	566	3	5	1150	1163	3	-4	1415	1294	7	-3	677	615	3	-9	1597	1475
2	7	552	577	5	-5	440	463	5	-6	1557	1890	4	0	3006	2674	3	6	1729	1591	3	-5	2033	1777	7	-7	1015	1015	3	-10	1337	1257
2	13	710	829	5	-6	1332	1616	4	2	1395	1242	3	7	1607	1547	3	7	1607	1547	3	-6	2526	2174	7	-9	1490	1675	3	-11	1006	943
2	-3	2942	3240	6	2	390	373	5	-10	897	878	4	8	857	868	3	9	1369	1280	3	-7	609	974	8	2	1541	1595	4	0	1439	1175
2	-7	542	577	6	3	474	948	6	0	633	586	4	8	857	868	3	9	1369	1280	3	-9	1426	1566	8	15	0	1957	4	4	1516	1367
2	-13	765	823	6	5	2731	3141	6	1	563	447	4	10	1502	1504	3	10	993	949	3	-10	1205	1168	8	0	1786	1542	5	2	665	577
2	3	1113	1082	6	7	2047	2370	6	3	1629	1625	4	15	0	1160	3	15	0	749	4	0	1498	1206	8	-1	1674	1445	5	3	921	494
3	3	971	1056	6	8	1458	1576	6	5	1969	2118	4	0	1008	881	3	-3	1873	1670	4	2	3041	2750	8	-2	1517	1297	5	4	691	692
3	4	722	748	6	11	676	694	6	7	1245	1310	4	-1	533	454	3	-4	1875	1668	4	4	2787	2696	8	-3	616	579	6	1	1856	1604
3	7	1081	1166	6	11	676	694	6	11	676	694	4	-2	674	642	3	-5	1718	1429	4	4	2787	2696	8	-3	616	579	6	1	1856	1604
3	8	1594	1515	7	2	311	469	7	0	1493	1340	4	-3	1426	1264	3	-6	1889	1721	4	5	1691	1764	8	-4	603	654	6	3	1654	1531
3	9	1811	1740	7	4	561	639	7	1	1532	1504	4	-6	583	599	3	-7	1549	1330	4	15	0	894	8	-7	598	452	7	15	0	1366
3	10	1781	1822	7	5	379	500	7	2	1891	1873	4	-8	712	740	3	-8	1404	1296	4	-3	731	594	9	-10	508	273	7	-3	732	532
3	11	1704	1644	7	7	495	593	7	3	1786	1888	5	0	2203	1879	3	-11	733	649	4	-4	875	796	9	0	1354	1365	7	-5	976	852
3	12	1482	1412	7	8	690	827	7	4	1601	1648	5	1	2575	2367	3	-12	1131	1087	4	-5	1146	1129	9	1	1243	1348	8	0	1368	1189
3	13	1305	1212	7	9	695	698	7	5	1341	1380	5	2	1968	1915	3	-13	1008	1041	4	-6	1063	1742	9	2	971	1020	8	-4	947	821
3	-11	1083	1083	7	10	927	936	7	6	644	609	5	3	1591	1515	3	-14	972	983	4	-7	1179	1276	9	4	559	469	8	-5	901	805
3	-3	1643	1256	7	-1	1265	1176	7	7	5124	5124	5	4	686	762	3	-15	1728	1756	4	-8	1092	702	9	-2	1303	1403	9	-6	1471	798
3	-4	755	748	7	-3	428	468	7	11	503	402	5	7	715	693	4	0	1752	1336	5	1	1093	922	9	-2	1335	1314	7	-7	1337	1186
3	-7	1082	1116	7	-11	944	1048	7	0	685	631	5	8	761	709	4	2	720	669	5	2	954	833	9	-3	949	909	8	-8	1272	1081
3	-8	1542	1515	8	1	408	468	7	-2	579	593	5	9	1087	1158	4	3	463	459	5	3	494	457	9	-4	943	941	8	-9	1177	1129
3	-9	1850	1780	8	4	1492	1642	7	-3	896	902	5	10	869	856	4	4	1760	1662	5	7	463	440	9	-6	865	624	9	0	978	861
3	-10	1950	1822	8	6	1979	2323	7	-5	1774	2130	5	11	710	725	4	6	1412	1338	5	-1	755	635	9	-8	1663	1635	9	-2	459	229
3	-11	1653	1646	8	8	1823	2090	7	-7	1969	2375	5	0	2219	1893	4	8	662	662	5	-3	676	645	9	-10	1966	2042	9	-4	946	850
3	-12	1516	1612	8	10	963	963	7	-8	1389	1656	5	-2	1692	1561	4	0	568	520	5	-4	692	702	10	-6	491	462	10	-8	1469	1373
3	-1																														

Table XXIV

Positional parameters for $\text{H}_2\text{Re}_2(\text{CO})_8$

Atom	x	y	z	Isotropic B
Re1	.2948(2)	.1289(2)	.2691(1)	1.69(5)
Re2	-.0467(2)	.1190(2)	.1800(2)	1.80(5)
C1	.2948(78)	.2262(53)	.1432(50)	2.6(11)
O1	.3114(65)	.2936(42)	.0798(41)	4.0(10)
C2	.4701(67)	.0340(48)	.2412(43)	1.8(9)
O2	.5621(65)	-.0183(40)	.2354(40)	3.8(9)
C3	.4502(65)	.2249(42)	.3710(43)	1.7(9)
O3	.5460(64)	.2812(43)	.4340(41)	4.0(10)
C4	.2994(71)	.0220(47)	.3921(46)	2.1(10)
O4	.3135(69)	-.0325(48)	.4675(46)	2.0(10)
C5	-.0512(85)	.0178(56)	.3111(55)	3.0(13)
O5	-.0721(61)	-.0365(40)	.3680(41)	3.7(9)
C6	-.2190(88)	.2035(56)	.2005(54)	3.0(12)
O6	-.3072(59)	.2611(40)	.2199(38)	3.4(9)
C7	-.1934(95)	.0163(61)	.0827(58)	3.6(13)
O7	-.2868(76)	-.0499(51)	.0298(48)	5.3(12)
C8	-.0451(73)	.2144(51)	.0457(48)	2.3(10)
O8	-.0743(73)	.2576(50)	-.0362(49)	5.1(12)

RESULTS

A drawing of the molecule is shown in Figure 19, giving the numbering scheme used. The diagram was drawn using the Calcomp plotter and Carroll K. Johnson's programme ORTEP. The bond lengths and angles within the molecule are found in Table XXV, while Table XXVI gives some intermolecular contacts. These distances and their associated errors were calculated using ORFFE2.

Figure 19

View of $\text{H}_2\text{Re}_2(\text{CO})_8$ down b axis

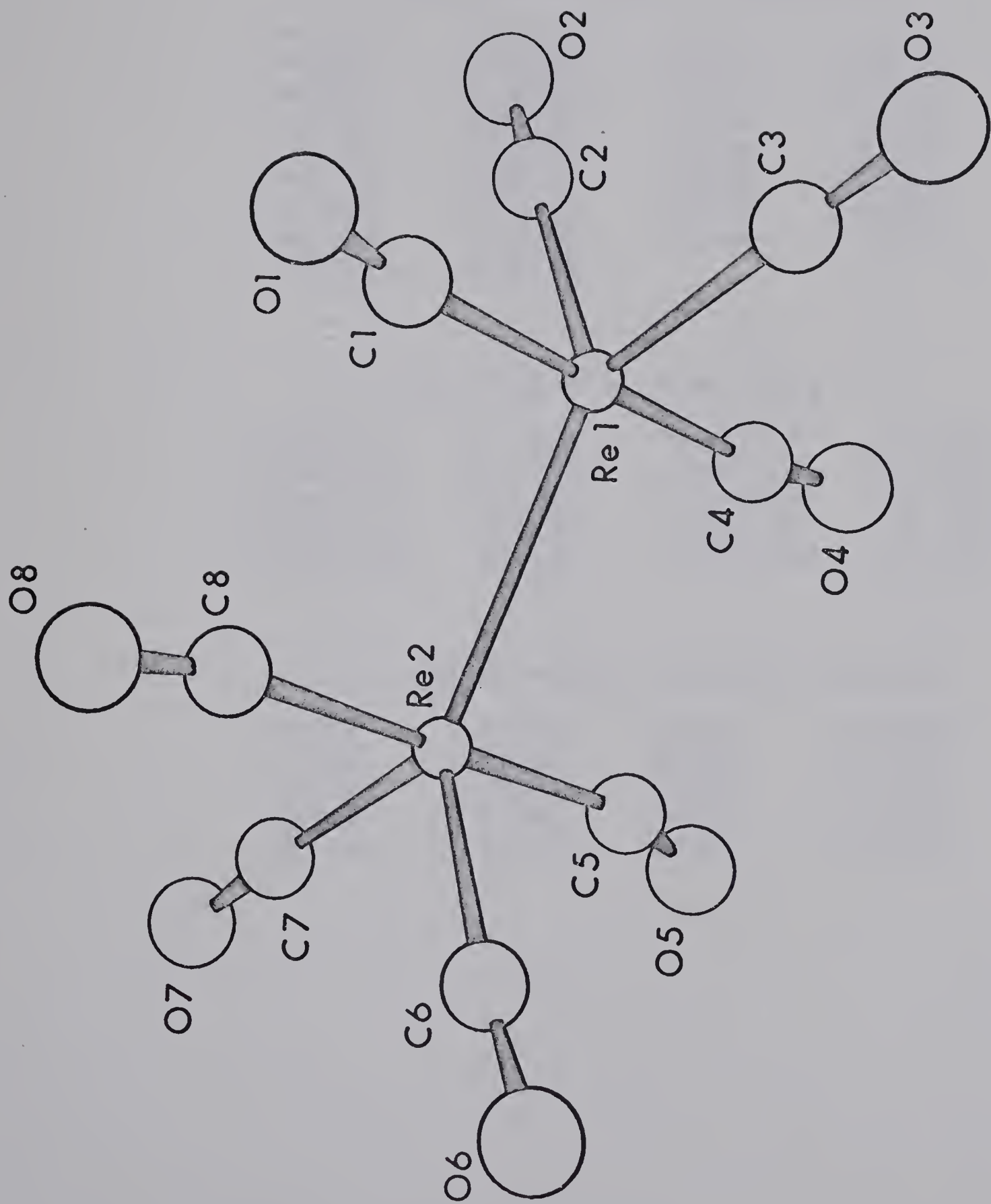


Table XXV

(a) Bond lengths ($\overset{\circ}{\text{\AA}}$) in $\text{H}_2\text{Re}_2(\text{CO})_8$

Re1-Re2	2.896(3)	C1-O1	1.17(7)
Re1-C1	1.97(6)	C2-O2	1.05(7)
Re1-C2	2.05(6)	C3-O3	1.17(7)
Re1-C3	1.92(5)	C4-O4	1.13(8)
Re1-C4	2.00(6)	C5-O5	1.03(7)
Re2-C5	2.06(7)	C6-O6	1.13(8)
Re2-C6	1.92(7)	C7-O7	1.17(9)
Re2-C7	1.91(7)	C8-O8	1.12(7)
Re2-C8	2.06(6)		

(b) Bond angles in $\text{H}_2\text{Re}_2(\text{CO})_8$

C1-Re1-C2	87.(2)	C5-Re2-C6	87.(3)
C1-Re1-C4	176.(2)	C5-Re2-C8	178.(2)
C2-Re1-C3	90.(2)	C6-Re2-C7	90.(3)
C3-Re1-C4	90.(2)	C7-Re2-C8	89.(3)
C1-Re1-Re2	89.(2)	C5-Re2-Re1	90.(2)

(c) Intramolecular non-bond contacts

C1-C8	2.89(9)	C4-C5	2.97(10)
O1-O8	3.31(8)	O4-O5	3.27(8)
C1-C2	2.78(8)	C5-C6	2.74(9)
C2-C3	2.82(8)	C6-C7	2.70(10)
C3-C4	2.77(7)	C7-C8	2.78(9)

Table XXVI

Intermolecular contacts (\AA)

C1-O7	3.40 (8)	C2-O7	3.33 (8)
C1-O2	3.40 (8)	C2-O6	3.37 (7)
C1-C7	3.93 (9)	C2-C6	3.59 (9)
O1-O4	3.13 (8)	C2-O4	3.60 (8)
O1-O7	3.14 (8)	O2-O3	3.12 (7)
O1-O5	3.15 (7)	O2-O5	3.16 (8)
O1-O2	3.15 (7)	O2-C3	3.27 (7)
O1-O8	3.31 (8)	O2-C5	3.30 (9)
O1-C8	3.21 (8)	O2-C6	3.36 (8)
O1-C3	3.32 (8)	O2-O7	3.36 (8)
O1-O6	3.32 (7)	O2-C7	3.41 (9)
C3-O4	3.30 (7)	O3-O7	3.19 (8)
C3-O6	3.38 (7)	O3-O4	3.24 (8)
C3-O7	3.44 (8)	O3-O8	3.33 (8)
C4-O5	3.31 (8)	O3-O6	3.43 (7)
C4-O4	3.34 (8)	O4-O4	3.26 (12)
C4-O6	3.37 (7)	O4-O6	3.38 (7)
C5-O8	3.31 (9)	O5-C8	3.15 (7)
C6-O8	3.23 (9)	O6-O8	3.15 (7)

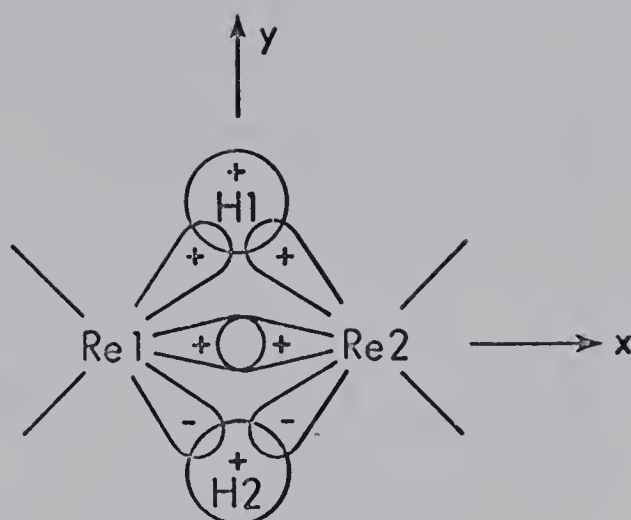
DISCUSSION

The compound $\text{Re}_2\text{H}_2(\text{CO})_8$ shows octahedral geometry about the rhenium atoms if one assumes that the hydrogen atoms occupy the empty region between the Re atoms and complete the octahedron. The molecule as a whole has D_{2h} symmetry, although none is required crystallographically.

There is little deviation from 90° of the C-Re-C angles, the largest being -3.2 degrees. The Re-C bonds range in length from 1.91\AA to 2.06\AA . Those being cis to the hydrogens are longer than those trans (average $2.02(3)\text{\AA}$ as compared to $1.95(4)$). Since the hydrogen-rhenium bond is weaker than rhenium-carbon, the bonds trans to the hydrogen would be stronger than those trans to carbonyl and thus have shorter bond lengths. The hydrogen atoms are assumed to occupy the two remaining octahedral sites. If there is undistorted octahedral symmetry (i.e., H-Re-H of 90°), then a Re-Re distance of $2.896(3)\text{\AA}$ would require a Re-H distance of 2.05\AA . This is rather longer than the average Re-H distance (terminal) in ReH_9^{72} of $1.68(1)\text{\AA}$ from neutron diffraction work, and would suggest that there is a slight distortion from octahedral symmetry by the reduction of the H-Re-H angle

in order to shorten the Re-H distance. The Re-H bond length lies between two extremes - $2.05\overset{\circ}{\text{\AA}}$ for undistorted octahedral symmetry, and $1.45\overset{\circ}{\text{\AA}}$ if the M-H-M unit is linear. Most likely it is close to those found in other Re-H structures of $\sim 1.70\overset{\circ}{\text{\AA}}$. This would require a H-Re-H angle of 63° . The Re-Re distance of $2.896(3)\overset{\circ}{\text{\AA}}$ is shorter than would be expected for a Re-Re unbridged bond. A Re-Re bond of $3.035\overset{\circ}{\text{\AA}}$ was found in $[\text{H}_2\text{Re}_3(\text{CO})_{12}]^=^{73}$, 2.96 , 2.96 and $3.02\overset{\circ}{\text{\AA}}$ in $\text{Re}_4(\text{CO})_{16}^=^{70}$ and $3.000\overset{\circ}{\text{\AA}}$ in $(\text{Si}(\text{C}_6\text{H}_5)_2)_2\text{Re}_2(\text{CO})_8^{77}$. A similar shortening of the metal-metal bond was observed in $\text{H}_2\text{Os}_3(\text{CO})_9^{75}$, also containing a doubly hydrogen bridged Os-Os bond. This shortening of M-M bonds is the opposite to the effect of bond lengthening found in singly bridged bonds, and suggests that the bonding of the $\begin{array}{c} \text{H} \\ \diagup \quad \diagdown \\ \text{M} \quad \text{M} \\ \diagdown \quad \diagup \\ \text{H} \end{array}$ unit is very different from that of the $\begin{array}{c} \text{H} \\ \diagup \quad \diagdown \\ \text{M} \quad \text{M} \end{array}$ group.

In order to explain the bonding of the rhenium and hydrogen atoms, a molecular orbital calculation was done. If one considers only the rhenium-hydrogen unit, then the following orbitals can be drawn⁷⁷, assuming no interaction with the carbonyl groups:



It should be borne in mind that such an assumption, as well as the choice of orbitals, can prejudice the results and may invalidate them. Using D_{2h} symmetry, the following orbitals are permitted⁴²:

$$\Gamma_{x^2 - y^2} = A_g + B_{3u}$$

$$\Gamma_H = A_g + B_{2u}$$

$$\Gamma_{xy} = B_{1g} + B_{2u}$$

The orbitals ϕ_1 to ϕ_6 are defined as follows:

ϕ_1	$5d_{x^2 - y^2}$	orbital on Re1
ϕ_2	$5d_{x^2 - y^2}$	orbital on Re2
ϕ_3	1s	orbital on H1
ϕ_4	1s	orbital on H2
ϕ_5	$5d_{xy}$	orbital on Re1
ϕ_6	$5d_{xy}$	orbital on Re2

Then the A_g orbitals can be expressed as

$$\psi'_{A_g} = \frac{1}{\sqrt{2}} \begin{pmatrix} \phi_1 + \phi_2 \end{pmatrix}$$

$$\text{and } \psi''_{A_g} = \frac{1}{\sqrt{2}} \begin{pmatrix} \phi_3 + \phi_4 \end{pmatrix}$$

Symmetry considerations permit that these may combine. Using Hückel molecular orbital theory⁷⁸, the Slater equations may be set up and the determinant obtained:

$$\begin{vmatrix} H_{A'A'} - E & H_{A'A''} \\ H_{A'A''} & H_{A''A''} - E \end{vmatrix} = 0$$

Defining

$$\alpha_R = \int \phi_1^H \phi_1 d\tau = \int \phi_2^H \phi_2 d\tau$$

$$\alpha_{R'} = \int \phi_5^H \phi_5$$

$$\alpha_H = \int \phi_3^H \phi_3 d\tau = \int \phi_4^H \phi_4 d\tau$$

$$= \int \phi_6^H \phi_6$$

$$\beta_R = \int \phi_1^H \phi_2 d\tau = \int \phi_2^H \phi_1 d\tau$$

$$\beta_{R'} = \int \phi_5^H \phi_6$$

$$\beta_{R-H} = \int \phi_1^H \phi_3 d\tau = \int \phi_1^H \phi_4 d\tau$$

$$\beta_{R'H} = \int \phi_6^H \phi_3$$

$$= \int \phi_2^H \phi_3 d\tau = \int \phi_2^H \phi_4 d\tau$$

$$\int \phi_5^H \phi_3$$

$$\beta_{H-H} = \int \phi_3^H \phi_4 d\tau = \int \phi_4^H \phi_3 d\tau$$

the determinant can then be rewritten as

$$\begin{vmatrix} \alpha_R + \beta_{RR}^{-E} & 2\beta_{R-H} \\ 2\beta_{R-H} & \alpha_H + \beta_{HH}^{-E} \end{vmatrix} = 0$$

If the assumption is made that the overlap between the rhenium and hydrogen atoms is very small, then $\beta_{R-H} \approx 0$ and the expression simplifies to

$$\begin{vmatrix} \alpha_R + \beta_{RR}^{-E} & 0 \\ 0 & \alpha_H + \beta_{HH}^{-E} \end{vmatrix} = 0$$

which can be very easily solved to give

$$\begin{aligned} E_1 &= \alpha_R + \beta_{RR} \\ \text{and } E_2 &= \alpha_H + \beta_{HH} \end{aligned}$$

Similar calculations can be carried out for the other set of orbitals, B_{2u} :

$$\psi'^{B_{2u}} = \frac{1}{\sqrt{2}} \begin{pmatrix} \phi_5 + \phi_6 \end{pmatrix} \quad \text{and} \quad \psi''^{B_{2u}} = \frac{1}{\sqrt{2}} \begin{pmatrix} \phi_3 - \phi_4 \end{pmatrix}$$

The energies were calculated to be

$$E_3 = \frac{-b + \sqrt{b^2 - 4c}}{2c}$$

$$\text{and } E_4 = \frac{-b - \sqrt{b^2 - 4c}}{2c}$$

where $b = \alpha_{R'} + \beta_{R'} + \alpha_H - \beta_H$

and $c = 4\beta_{R'H}^2 + \alpha_{R'}\alpha_H - \alpha_{R'}\beta_H + \alpha_H\beta_{R'} - \beta_H\beta_{R'}$

If the assumption is made that the β^2 and cross β terms are much smaller than other terms, then E_3 and E_4 can be approximated to

$$E_3 = \alpha_H - \beta_H + \beta_{R'H}$$

and $E_4 = \alpha_{R'} + \beta_{R'} - \beta_{R'H}$

There is only one orbital combination with B_{3u} symmetry:

$$\psi_{B_{3u}} = \frac{1}{\sqrt{2}} \begin{pmatrix} \phi_1 - \phi_2 \end{pmatrix}$$

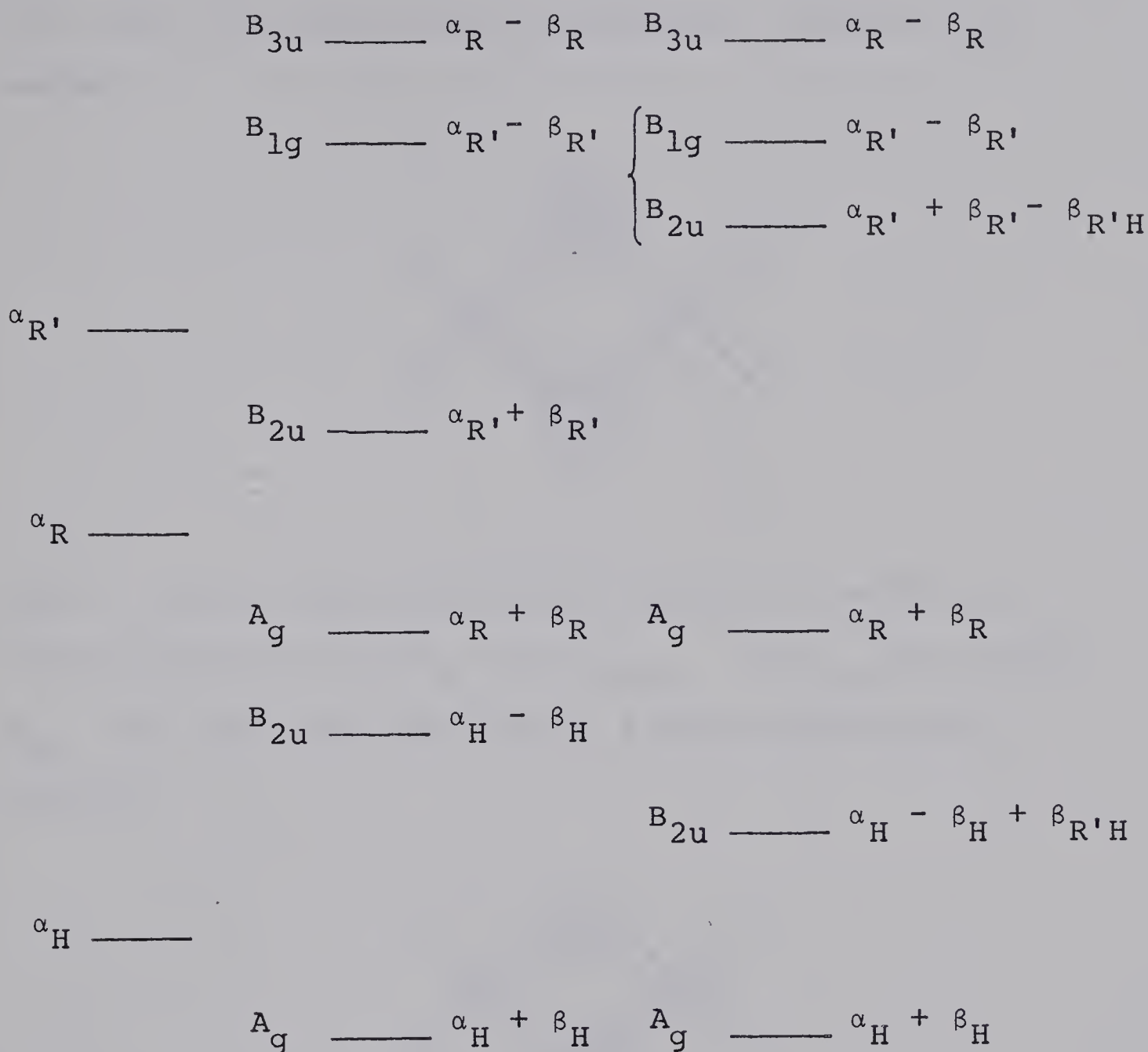
$$E_5 = \alpha_R - \beta_{RR}$$

and one with B_{1g} symmetry:

$$\psi_{B_{1g}} = \frac{1}{\sqrt{2}} \begin{pmatrix} \phi_5 - \phi_6 \end{pmatrix}$$

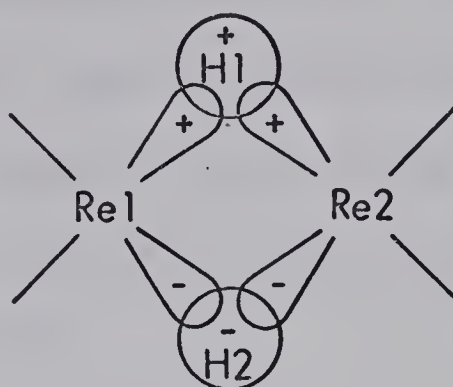
$$E_6 = \alpha_R - \beta_{R'R'}$$

The following energy level diagram will be obtained:

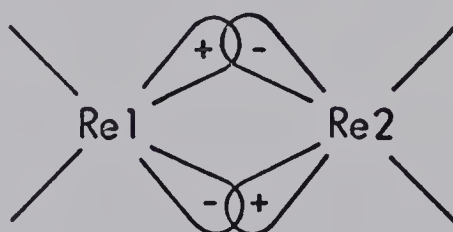


where the order of the levels B_{1g} and the higher B_{2u} is unknown and depends on the relative sizes of $\beta_{R'}$ and $\beta_{R'H}$. If $\beta_{R'H}$ is greater than $2\beta_{R'}$, then B_{1g} will lie below B_{2u} . There are eight electrons which occupy these orbitals so that the bottom four will be filled. The bonding of the Re-H unit will be (1) both A_g orbitals filled (that is, (a) the electrons localized on the

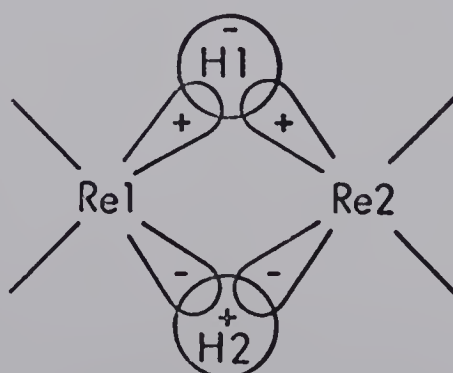
hydrogen atoms, and (b) a Re-Re σ bond), (2) the bonding B_{2u} orbital filled in which the electrons are localized over both rheniums and both hydrogens, forming Re-H bonds:



and (3) one of the antibonding orbitals depending on the relative size of β_R , and $\beta_{R'H}$. If B_{1g} lies below B_{2u} , then the electrons occupy a Re-Re antibonding orbital:



Otherwise, the two remaining electrons will occupy a Re-H antibonding orbital in which the electrons are delocalized over all four atoms:



The bonding can thus be considered as a Re-Re single bond with delocalization of the remaining electrons over both the rheniums and hydrogens, rather than a protonation of a Re=Re double bond. Examination of the energy level diagram shows that the hydrogens act as a stabilizing unit even if there were no Re-H bond. Removal of the hydrogens to give Re-Re^{++} would result in no bond at all, since all four of the remaining levels would be filled.

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Appendix A

SUBROUTINES ROTOR AND BESSEL
OF SFLS5HR


```

SUBROUTINE ROTOR
COMMON TITLE(18),MODE,INV,NAD,ISAN,NPU,NEW(5),NFSQ,NFOUR,
INCOR,KARD,ID,LV,NLV,NS,NU,NU,MLSC,KRIT,RPAR,NATE,ICENT
2,NRIG,NONR,NHIN
COMMON A,B,C,AL,BE,GA,AQ,BQ,CQ,ALQ,BEQ,GAQ,NSL,SUMDL
COMMON NPAR,NVAR,NCNT,KSEL(921)
COMMON S(50),BO,GP(70),GQ(60),GR(60),MF(60),MG(60),MI(60),
2X(140),Y(140),Z(140),B11(140),B22(60),B33(60),B12(60),
2B13(60),B23(60),NA(70),NAME(140),XC(10),YC(10),ZC(10),
3P(10),R(10),U(10),BRB(10,8),NSTEP(10),NORD(10),RID(10)
COMMON MLSW,AW,BW,CW,DW,EW,FW,GW,HW,SIZE1,SIZE2
COMMON SAER(10,10),NSAER(10,10),SAES(10),NSAES(10),SAEF(10),
1NSAEF(10)
COMMON NTRAN(4,12,12),NVTN(4,36)
COMMON NF,FM(32,20),F(20),FP(20),FPP(20),RHO,RHOSQ
COMMON FOBS,FCAL,SFC,AOBS,ACA_,BOBS,BCAL,DFL,SIGMA,EXT1,EXT2
1,MS1,MREJ,Q,MH,MK,ML,TH,TK,TL,DA(180),DB(180),DF(180)
COMMON CHG(921),STD(921),EPS(180),DIAG(180),V(180)
COMMON A11,A12,A13,A22,A23,A33,BA1,BA2,BA3,BA4,BA5,BA6
COMMON SN(50),BON,GPB(70),GQB(60),GRB(60),XB(140),YB(140),
1ZN(140),B11N(140),B22N(60),B33N(60),B12N(60),B13N(60),
2B23N(60),XCN(10),YCN(10),ZCN(10),PN(10),RN(10),UN(10),
3BRB(10,8),RIDN(10)
COMMON/SPSY/I,J,N,RH,RK,RL,T,RHH,RKK,RLL,RHK,RHL,RKL,NR
COMMON /PDC/SPK,SPL,SPP,SPR,SPU,CPH,CPK,CPL,CPP,CPR,CPU,WL,
1 SACAL, WB,SBCAL,K,L,TPI,JF,LC ,SPH,INV1
COMMON/BESS/DBJB(10,20),LM(6),BR,BEA,TE,TD,BDD,BD,BIA(10,20)
1,DBIB(10),BIO(10),JPA
DIMENSION WL(20),WB(20)
EQUIVALENCE(BI,B11),(BIN,B11N),(AR,SN)
DIMENSION AR(20100)
DIMENSION BI(140),BIN(140)
DIMENSION SUM(6),DELMS(2,6),DELMC(2,6),DELC(2,3),DELS(2,3)
REAL MTERM
REAL*8 CC,DD,AP,AP2
COMMON/RIGID/CLIMIT
TE=0.
TD=0.
QW=0.
BPOS=1.0
BNEG=-1.0
BD=BRB(I,2)
NK=MOD(NORD(I),2)+1
CD=COS(P(I))
CF=COS(R(I))
SD=SIN(P(I))
SE=SIN(R(I))
TPI=6.283185
XK=TPI*RID(I)
TF=EXP(-BRB(I,1)*RHOSQ)*GP(I+NU)*ICENT*NSTEP(I)
IF(NAD.EQ.2) TF=TF*2.
TF=TF*F(JF)
DO 14 IWA=1,2
DO 10 IW=1,6

```



```

        DELMC(IWA,IW)=0.
        DELMS(IWA,IW)=0.0
        IF(IW.GT.3) GO TO 10
        DELC(IWA,IW)=0.
        DELS(IWA,IW)=0.
10  CONTINUE
14  CONTINUE
        CODE=0.
        DO 23 J=1,MU
            ICCODE=0
                CALL SYM
                PHI=TPI*(RH*XC(I)+RK*YC(I)+RL*ZC(I)+T)
                CP=COS(PHI)*TF
                SP=SIN(PHI)*TF
            XK1=RH*BA1+RK*BA2+RL*BA3
            XK2=RL*BA6
            XK3=RK*BA4+RL*BA5
            CC=XK*(XK1*CE-XK2*SE)
            DD=XK*(XK1*SD*SE+XK3*CD+XK2*CE*SD)
            AP=DSORT(CC*CC+DD*DD)
            XNU=DARCOS(CC/AP)
            IF(DD.LT.0.0)XNU=-XNU
            ANGLE=NORD(I)*(U(I)-XNU)
            IF(MODE.EQ.1) GO TO 2
            IF(KSEL(L+7).EQ.0.AND.KSEL(L+8).FQ.0) GO TO 2
            DELCE=-XK*(XK1*SE+XK2*CE)
            DELDE=XK*(XK1*SD*CE-XK2*SD*SE)
            DELDD=XK*(CD*SE*XK1-SD*XK3+CE*CD*XK2)
            AP2=AP*AP
            DELAPD=DD*DELDD/AP
            DELAPE=(CC*DELCE+DD*DELDE)/AP
            TE=DELAPE/AP
            TD=DELAPD/AP

```

C
C ***** THIS NEXT SECTION WILL CAUSE DIVIDE BY ZERO ERRORS
C FOR CERTAIN REFLECTIONS IF D OR E ARE 0,90,180,OR 270DEGS
C BY PASS IS TO OMIT THESE REFLECTIONS FROM LEAST SQUARES
C

```

        CW=AP2*DSORT(1.-CC*CC/AP2)
        IF (ABS(CW).LT.0.00001) GO TO 1
        DXNU=1./CW
        DXNUE=-DXNU*(AP*DELCE-CC*DELAPE)
        DXNUD=DXNU*CC*DELAPD
        GO TO 2
1  ICODE=1
2  IF(NK.EQ.2) GO TO 3

```

C
C ***** CALC'N OF REAL TERMS *****
C

```

15  NPRIME=NK*NORD(I)/2
        NAA=1
        GO TO 21
22  F=SNGL(AP)
        CALL BPSSEL (BJ,OW,F ,BNEG)

```



```
SUM(1)=DBIB(I)*BJ
SUM(2)=BR
SUM(3)=BDD
SUM(4)=BEA
SUM(5)=0.
SUM(6)=BJ
GO TO 5
```

C
C
C

***** CALC'N OF IMAG TERMS *****

```
3 NPRIME=(NORD(I)-1)/2
  DO 4 IA=1,6
    SUM(IA)=0.
4 CONTINUE
  NAA=2
21 DO 24 IA=1,5
    LM(IA)=KSEL(IA+L+4)
24 CONTINUE
  LM(6)=1
  IF(NAA.EQ.1) GO TO 22
5 PSIGN=1.
  IF (ICODE.FO.0) GO TO 16
  SUM(3)=0.
  SUM(4)=0.
  LM(3)=0
  LM(4)=0
16 DO 6 IA=1,NPRIME
    PSIGN=PSIGN*(-1)
6 CONTINUE
  SIGN=PSIGN
  IF(NK.EQ.2) PSIGN=-1.
  NL=1
  IF(NAA.NE.NK)NL=2
```

C
C
C

***** CALC'N OF M AND DERIVATIVES *****

```
DO 11 JP=NL,20,NK
  JPA=JP
  PA=FLOAT(JP)
  PNA=PA*NORD(I)
  CU=COS(PA*ANGLE)
  SU=SIN(PA*ANGLE)
  TCDM=2.*SIGN*BIA(I,JP)/BID(I)
  E=SNGL(AP)
  CALL BESSEL (BJ,PNA,E ,BNES)
  IF(MODE.EQ.1) GO TO 34
  IF(LM(1).EQ.0) GO TO 30
```

C
C
C

***** CALC'N OF DM/D(BD) *****

```
  BDTERM=2.*SIGN*CU*DBJB(I,JP)*BJ
  SUM(1)=SUM(1)+BDTERM
  IF(ABS(BDTERM/SUM(1)).LE.CLIMIT) LM(1)=0
30 IF(LM(2).EQ.0) GO TO 31
```



```

C
C
C      ***** CALC'N OF DM/DR *****
C
C      RTERM=TCOM*BR*CU
C      SUM(2)=SUM(2)+RTERM
C      IF (ABS(RTERM/SUM(2)),LE,CLIMIT) LM(2)=0
31    IF (LM(3).EQ.0 ) GO TO 32
C
C
C      ***** CALC'N OF DELMD *****
C
C      DHD=SU*PNA*DXNUD
C      DTERM=TCOM*(BDD*CU+BJ*DMO)
C      SUM(3)=SUM(3)+DTERM
C      IF (ABS(DTERM/SUM(3)),LE,CLIMIT) LM(3)=0
32    IF (LM(4).EQ.0) GO TO 33
C
C
C      ***** CALC'N OF DM/DE *****
C
C      DNE=SU*PNA*DXNUE
C      ETERM=TCOM*(BEA*CU+BJ*DME)
C      SUM(4)=SUM(4)+ETERM
C      IF (ABS(ETERM/SUM(4)),LE,CLIMIT) LM(4)=0
33    IF (LM(5).EQ.0) GO TO 34
C
C
C      ***** CALC'N OF DM/DG *****
C
C      GTERM=-PNA*TCOM*SU*BJ
C      SUM(5)=SUM(5)+GTERM
C      IF (ABS(GTERM/SUM(5)),LE,CLIMIT) LM(5)=0
34    IF (LM(6).EQ.0) GO TO 35
C
C
C      ***** CALC'N OF M *****
C
C      MTERM=TCOM*CU*BJ
C      SUM(6)=SUM(6)+MTERM
C      IF (ABS(MTERM/SUM(6)),LE,CLIMIT) LM(6)=0
35    LSUM=LM(1)+LM(2)+LM(3)+LM(4)+LM(5)+LM(6)
C      IF (MODE.EQ.1) LSUM=LM(6)
C      IF (LSUM.EQ.0) GO TO 12
C      SIGN=SIGN*PSIGN
11  CONTINUE
C      WRITE(6,201)RH,RK,RL
12  IF (MODE.EQ.1) GO TO 7
C      SUM(1)=(SUM(1)-SUM(6)*DBIB(I))/BIO(I)
C      DO 36 IA=1,5
C      DELMC(NAA,IA)=SUM(IA)*CP+DELMC(NAA,IA)
C      DELMS(NAA,IA)=SUM(IA)*SP+DELMS(NAA,IA)
36  CONTINUE
C      IF ((KSEL(L+1)+KSEL(L+2)+KSEL(L+3)).EQ.0) GO TO 7
C      DELC(NAA,1)=DELC(NAA,1)+CP*SUM(6)*RH
C      DELC(NAA,2)=DELC(NAA,2)+CP*SUM(6)*RK
C      DELC(NAA,3)=DELC(NAA,3)+CP*SUM(6)*RL
C      DELS(NAA,1)=DELS(NAA,1)+SP*SUM(6)*RH
C      DELS(NAA,2)=DELS(NAA,2)+SP*SUM(6)*RK

```



```

      DELS(NAA,3)=DELS(NAA,3)+SP*SUN(6)*RL
7  DELMC(NAA,6)=SUN(6)*CP+DELMC(NAA,6)
      DELMS(NAA,6)=SUN(6)*SP+DELMS(NAA,6)
      IF(NAA.EQ.2) GO TO 15
23  CONTINUE
      AT=DELMC(1,6)-DELMS(2,6)
      ACAL=ACAL+AT
      IF(NAD.EQ.2) GO TO 40
      BT=DELMC(2,6)+DELMS(1,6)
      BCAL=BCAL+BT
40  IF(MODE.EQ.1) GO TO 9

C
C      ***** CALC'N OF DELF(POPULATION FACTOR) *****
C
      IF(KSEL(L).EQ.0) GO TO 58
      DA(K)=AT/GP(I+NU)
      IF(INV1.EQ.1) GO TO 59
      DB(K)=BT/GP(I+NU)
59  K=K+1
58  L=L+1

C
C      ***** DERIVATIVES W.R.T. CENTRE OF RING *****
C
      DO 56 IA=1,3
        IF(KSEL(L).EQ.0) GO TO 55
        DA(K)=-TPI*(DELS(1,IA)+DELC(2,IA))
        IF(INV1.EQ.1) GO TO 57
        DB(K)=TPI*(-DELS(2,IA)+DELC(1,IA))
57  K=K+1
55  L=L+1
56  CONTINUE
      IF(KSEL(L).EQ.0) GO TO 53

C
C      ***** CALC'N OF DELF(TEMP) *****
C
      DA(K)=-RHOSQ*(DELMC(1,6)-DELMS(2,6))
      IF(INV1.EQ.1) GO TO 54
      DB(K)=-RHOSQ*(DELMC(2,6)+DELMS(2,6))
54  K=K+1
53  L=L+1

C
C      ***** CALC'N OF DELFD TO DELF3 *****
C
      DO 50 IA=1,5
        IF(KSEL(L).EQ.0) GO TO 51
        DA(K)=DELMC(1,IA)-DELMS(2,IA)
        IF(INV1.EQ.1) GO TO 52
        DB(K)=DELMC(2,IA)+DELMS(1,IA)
52  K=K+1
51  L=L+1
50  CONTINUE
      9  RETURN
201  FORMAT(' REFLECTION ',3F3.0,' HAS NOT CONVERGED')
      END

```



```

SUBROUTINE BESSEL (BSUM, PP, XZ, BSIGN )
COMMON TITLE(18),MODE,INV,NAD,ISAN,NPJ,NEW(5),NFSQ,NFOUR,
1NCDP,KARD,IO,LV,NLV,NS,NU,MU,MLSC,KRIT,RPAR,NATE,ICENT
2,NRIG,NONR,NHIN
COMMON A,B,C,AL,BF,GA,AQ,BQ,CQ,ALQ,BEQ,GAQ,NSL,SUMDL
COMMON NPAR,NVAR,NCNT,KSEL(921)
COMMON S(50),BC,GP(70),GQ(60),GR(60),MF(60),MG(60),MI(60),
2X(140),Y(140),Z(140),B11(140),B22(60),B33(60),B12(60),
2B13(60),B23(60),NA(70),NAME(140),XC(10),YC(10),ZC(10),
3P(10),R(10),U(10),BRB(10,8),NSTEP(10),NORD(10),RID(10)
COMMON MLSW,AW,BW,CW,DW,EW,FW,GW,HW,SIZE1,SIZE2
COMMON SAER(10,10),NSAER(10,10),SAES(10),NSAES(10),SAEF(10),
1NSAEF(10)
COMMON NTRAN(4,12,12),NVTN(4,36)
COMMON NF,FM(32,20),F(20),FP(20),FPP(20),RHO,RHOSQ
COMMON FOBS,FCAL,SFC,AOBS,ACAL,BOBS,BCAL,DEL,SIGMA,EXT1,EXT2
1,MS1,PREFJ,Q,MH,MK,ML,TH,TK,TL,DA(180),DB(180),DF(180)
COMMON CHG(921),STD(921),EPS(180),DIAG(180),V(180)
COMMON A11,A12,A13,A22,A23,A33,BA1,BA2,BA3,BA4,BA5,BA6
COMMON SN(50),BON,GPN(70),GQN(60),GRN(60),XN(140),YN(140),
1ZN(140),B11N(140),B22N(60),B33N(60),B12N(60),B13N(60),
2B23N(60),XCN(10),YCN(10),ZCN(10),PN(10),RN(10),UN(10),
3BRB(10,8),RIDN(10)
COMMON/SPSY/I,J,N,RH,RK,RL,T,RHH,RKK,RLH,RHK,RHL,RKL,NR
COMMON /PDC/SPK,SPL,SPP,SPR,SPJ,CPH,CPK,CPL,CPP,CPR,CPU,WL,
1SACAL,WB,SBCAL,K,L,TPI,JF,LC,SPH,INV1
COMMON/BESS/DBJB(10,20),LM(6),BR,BEA,TE,TD,BDD,BD,BIA(10,20)
1,DSIB(10),BIO(10),JPA
COMMON/RIGID/CLIMIT
DIMENSION WL(20),WR(20)
EQUIVALENCE(BI,B11),(BIN,B11N),(AR,SN)
DIMENSION BI(140),BIN(140)
DIMENSION AR(20100)
DIMENSION LB(5)

```

C

```

BTERM=1.
BRTERM=0.
FRTERM=0.
DRTERM=0.
BJTERM=0.
LB(5)=1
IF(PP.EQ.0.) GO TO 17
FACP=1.
JP=IFIX(PP)
DO 1 M=1,JP
    XM=FLOAT(M)
    BTERM=BTERM!*XZ/(2.*XM)
1 CONTINUE
17 IF(BSIGN.LE.0.) GO TO 22

```

C

```

***** CALC'N OF I TERMS *****
LB(1)=1
DO 13 M=2,4
    LB(M)=0
13 CONTINUE

```



```

      IF (PP.NE.0.) BJTERM=PP*BTERM/BD
      DBJB(I,JPA)=BJTERM
      GO TO 5
C     ***** CALC'N OF M TERMS *****
22    LB(1)=0
      DO 14 M=2,4
          LB(M)=LM(M)
14    CONTINUE
      IF (PP.EQ.0.) GO TO 5
      BRTERM=BTERM*PP/RID(I)
      ERTERM=PP*BTERM*TE
      DRTERM=PP*BTERM*TD
5     BSUM=BTERM
      BR=BRTERM
      BEA=ERTERM
      BDD=DRTERM
      IF (BSIGN.GT.0) DBJB(I,JPA)=BJTERM
      DO 2 M=1,30
          BTERM=(BSIGN*BTERM*XZ*XZ)/(4.*M*(M+PP))
      IF (MODE.EQ.1) GO TO 11
      IF (LB(1).EQ.0) GO TO 6
C
C     ***** CALC'N OF DM/DBD TERM *****
C
      BJTERM=(2.*M+PP)*BTERM/BD
      DBJB(I,JPA)=DBJB(I,JPA)+BJTERM
      IF (ABS(BJTERM/DBJB(I,JPA)).LE.CLIMIT) LB(1)=0
6     IF (LB(2).EQ.0) GO TO 7
C
C     ***** CALC'N OF DM/DR TERM *****
C
      BRTERM=BTERM*(2.*M+PP)/RID(I)
      BR=BR+BRTERM
      IF (ABS(BRTERM/BR).LE.CLIMIT) LB(2)=0
7     IF (LB(3).EQ.0) GO TO 8
C
C     ***** CALC'N OF DM/DD *****
C
      DRTERM=BTERM*(2.*M+PP)*TD
      BDD=BDD+DRTERM
      IF (ABS(DRTERM/BDD).LE.CLIMIT) LB(3)=0
8     IF (LB(4).EQ.0) GO TO 11
C
C     ***** CALC'N OF DM/DE TERM *****
C
      ERTERM=BTERM*(2.*M+PP)*TE
      BEA=BEA+ERTERM
      IF (ABS(ERTERM/BEA).LE.CLIMIT) LB(4)=0

```



```
11 IF (LB(5).EQ.0) GO TO 12
C
C      ***** CALC'N OF BESSEL SUMMATION *****
C
      BSUM=BSUM+BTERM
      IF( ABS(BTERM/BSUM).LE.CLIMIT) LB(5)=0
12 LSUM=LB(1)+LB(2)+LB(3)+LB(4)+LB(5)
      IF(MODE.EQ.1) LSUM=LB(5)
      IF(LSUM.EQ.0) GO TO 3
2 CONTINUE
3 RETURN
END
```


Appendix B

PROGRAMMES USED

<u>Author</u>	<u>Title</u>	<u>Description</u>
D.P. Shoemaker	MIXG2	calculates Picker diff-ractometer settings modified by M. Elder to sort data
M.J. Bennett	PMMO	calculates intensities, makes Lp corrections for Picker data
M. Elder	Datap	as for PMMO, except for PAILRED data
M. Elder and K. Simpson	D-refine	refines axial lengths
A. Zalkin	FORDAP	Fourier summation for Patterson or Fourier maps
M.J. Bennett	MMMR	calculates starting para- meters for rigid bodies and hindered rotors
W.C. Hamilton	GON09	absorption corrections - Picker data
D.P. Shoemaker	GNABS	absorption corrections - PAILRED data
C.T. Prewitt	SFLS5HR	structure factor calcu- lation and least squares refinement of parameters modified by B.M. Foxman, M.J. Bennett and W.L. Brooks for rigid body and hindered rotor routines
G.J. Williams	CROMERS	calculates form factor curves using Cromer coefficients
J.S. Wood	MGEOM	calculates bond lengths, angles and best planes

Programmes Used (Contd.)

<u>Author</u>	<u>Title</u>	<u>Description</u>
W. Busing and H.A. Levy	ORFFE2	calculates bond lengths, angles and associated errors modified by B. Penfold for IBM 360, and W.L. Brooks and M. Elder for hindered rotors and rigid bodies
C. Johnson	ORTEP	writes plot command tape for Calcomp plotter
M. Elder and W.L. Brooks	SFLIST	lists structure factors for publication

Appendix C

CONVENTIONAL CRYSTALLOGRAPHIC SYMBOLS
AS DEFINED IN VOL. I, PAGE xi, OF THE
INTERNATIONAL TABLES FOR X-RAY CRYSTALLO-
GRAPHY

h, k, ℓ	indices of the reflection from a set of parallel planes; co-ordinates of a reciprocal lattice point
F_{hkl}	structure factor for the unit cell, corresponding to the Bragg reflection hkl
$a, b, c,$	lengths of unit cell edges
α, β, γ	interaxial angles
a^*, b^*, c^*	lengths of reciprocal lattice unit cell edges
$\alpha^*, \beta^*, \gamma^*$	interaxial angles in reciprocal space
x_i, y_i, z_i	fractional co-ordinates of an atom i (co-ordinates of atom i in units of a, b, c)
$\rho_{x,y,z}$	electron density at the point x, y, z
μ	linear absorption coefficient
\bar{u}^2	mean square amplitude of atomic vibration
B	Debye isotropic temperature parameter $B = 8\pi^2 \bar{u}^2$
$\beta_{11}, \beta_{22}, \beta_{33},$ $\beta_{12}, \beta_{13}, \beta_{23}$	anisotropic temperature parameters used to describe ellipsoidal electron distribution of the anisotropically vibrating atom; the temperature factor expression is then: $\exp[-(\beta_{11}h^2 + \beta_{22}k^2 + \beta_{33}\ell^2 + 2\beta_{12}hk + 2\beta_{13}h\ell + 2\beta_{23}k\ell)]$

F_o	observed structure factor
F_c	calculated structure factor

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